Ionic Bonding Tutorial
Tue, Dec 03 06:36 PM
Assignment Code: car15962

Name	Q1: Ions with the same charge will
AK	repel one another.
Caroline	repel one another.
Sam Carlo	repel one another.
Max	repel one another.
Shepard Munson	repel one another.
Sam Sweetser	repel one another.
Kyle Daniels	repel one another.
Ashley Hargrave	repel one another.
Sebastian Fox	
Sebastian Fox	repel one another.
Gina Edward	repel one another.
Will Cohen	repel one another.
Willwo Batty	repel one another.
Maya	repel one another.
Will Cohen	repel one another.
Fernanda More	repel one another.
wes	repel one another.
Willow Batty	repel one another.
kevy	repel one another.
Ryan Patel	repel one another.
Will Cohen	repel one another.
Brandon	repel one another.

Q2: In order for ions to attract each other, they must

have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.

have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.
have opposite charges.

Q3: Positive ions occur when atoms LOSE electrons. Positive ions come from

metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the right side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.

metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
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metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.

Q4: Negative ions occur when atoms

gain electrons.
lose electrons.
lose electrons.
gain electrons.
gain electrons.
lose electrons.
gain electrons.
gain electrons.

gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.

Q5: When Na+1 and Cl-1 ions bond together, the compound is called

sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride

sodium chlorate

sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride

Q6: Which of the following is NOT TRUE about a chloride ion?

decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
has 8 valence electrons

decreases in size
decreases in size
has 8 valence electrons
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size

Q7: If many ion pairs are nearby, they will spontaneously

self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.

line up with all the positive charges on one side.

self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.

Q8: Ion crystal are made of individual molecules.

False

False
False
False
False
False
False

False
False
False
False
False
False
False
False
False
False
False
False
True

Q9: In ionic compounds, the formula tells us the _____ of ions.

ratio
ratio
ratio
ratio
ratio
ratio
ratio
total number

ratio
ratio
ratio
ratio
size
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio

Q10: When Ca+2 ions and F-1 ions bond together, the calcium to fluoride ratio is 4:8, which reduces to a 1:2 ratio.

True
True
True
False
True
True
True
False

True
True
True
False
True
True
True
True
True
True
True
True
True

Q11: Magnesium forms a +2 ion. Chlorine forms a +1 ion. What is the chemical formula for magnesium chloride?
MgCl2
Mg2Cl
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2

MgCl
MgCl2
Mg2Cl
MgCl
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2

Q12: Sodium forms a +1 ion. Phosphorus forms a -3 ion. What is the chemical formula for sodium phosphide?
Na3P
S3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
NaP3
NaP3
NaP3
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P

Q13: Which of the following pairs is ionically bonded?

Na and F
P and Cl
Na and F
Na and F
Na and F
H and Cl
Na and F
Na and F

Na and F
Na and F
Na and F
Na and F
H and Cl
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F

Q14: The chemical formula for potassium chloride is
KCl
KCl3
KCl
KCl
KCl
KCl2
KCl
KCl

KCl
KCl
KCl
KCl
KCl2
KCl
KCl
KCl
KCl
KCl
KCl
KCl
KCl

Q15: The chemical formula for calcium bromide is	Number of Correct Answers
CaBr2	15
CaBr2	9
CaBr2	14
CaBr2	13
CaBr2	15
Ca2Br	11
CaBr2	15
CaBr2	12
	0
CaBr2	12
CaBr2	15
CaBr2	13
CaBr2	12
Ca4Br8	10
CaBr2	14
CaBr2	15
CaBr2	15
CaBr2	15
CaBr2	15
CaBr2	15
CaBr2	15
CaBr2	14

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