Ionic Bonding Tutorial
Tue, Dec 03 06:35 PM
Assignment Code: car15962
Class: 2019-2020 Krug Chemistry

Name	Q1: Ions with the same charge will
Parker Tran	repel one another.
Mithil Kulkarni	repel one another.
Cole Blassic	repel one another.
Braxton Fagan	repel one another.
Sam Sweetser	repel one another.
Tyler P. Smith The Great & Powerful	repel one another.
adelle topp	repel one another.
AJ	repel one another.
Kyle Daniels	repel one another.
Kyle David	repel one another.
Reagan Vale	repel one another.
Shepard Munson	repel one another.
Noah Chapman	repel one another.
Keiona	repel one another.
aidan sankowsky	repel one another.
karlee angel	repel one another.
Hudson Kennedy	repel one another.
Kaitlyn Correll	repel one another.
Bri Straight	repel one another.
Reece	repel one another.
Sydney Matthews	repel one another.
Mason Earle	repel one another.
Alec Strahan	repel one another.
Chelsea Witz	repel one another.
Camden Kirker	repel one another.
Max Raderer	repel one another.
troy	repel one another.
Michael Logan	repel one another.
Ashley Hargrave	repel one another.
Sebastian Fox	repel one another.
Jason Seitz	repel one another.
Dahlia	repel one another.
Owen Lindsay	repel one another.
Gina Edward-Shalabi	repel one another.
Liam McLaughlin	repel one another.
maddie	repel one another.
Shane Smith	repel one another.
William Loudermilk	repel one another.
balin	repel one another.
Orion	repel one another.
Wes	repel one another.
Kent Turner	repel one another.
Victoria Ell	repel one another.
Dana Peace	repel one another.
Shane Brown	repel one another.
Niklas Hatchett	repel one another.
kevy	repel one another.
Maya	repel one another.
Fernanda More	repel one another.
Emily Dodge	cancel each other out.
Bianca DeCarli	repel one another.
Ryan Patel	repel one another.
Savannah	repel one another.
Willow Batty	repel one another.
kyle gensone	repel one another.

Q2: In order for ions to attract each other, they must

have opposite charges.
have opposite charges.
have opposite charges.
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Q3: Positive ions occur when atoms LOSE electrons. Positive ions come from

metals on the right side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.

nonmetals on the right side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
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nonmetals on the right side of the periodic table.
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nonmetals on the left side of the periodic table.
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metals on the left side of the periodic table.

Q4: Negative ions occur when atoms
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.

Q5: When Na+1 and Cl-1 ions bond together, the compound is called

sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride

sodium chlorite

sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride

sodium chlorate

sodium chloride
sodium chloride

sodium chlorine

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sodium chloride
sodium chloride

Q6: Which of the following is NOT TRUE about a chloride ion?

is negatively charged
is attracted to positive ions
decreases in size
decreases in size
decreases in size
is negatively charged

is negatively charged
is attracted to positive ions
decreases in size
has 8 valence electrons
is negatively charged
has 8 valence electrons
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
is negatively charged
is negatively charged
decreases in size
has 8 valence electrons
decreases in size
decreases in size
decreases in size
decreases in size
is negatively charged
decreases in size
decreases in size
has 8 valence electrons
decreases in size
has 8 valence electrons
has 8 valence electrons
decreases in size
decreases in size
has 8 valence electrons

has 8 valence electrons
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
decreases in size
has 8 valence electrons
decreases in size
decreases in size
has 8 valence electrons
decreases in size
has 8 valence electrons
has 8 valence electrons
decreases in size

Q7: If many ion pairs are nearby, they will spontaneously
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
combine to create individual molecules.
self-assemble into a crystal lattice.
combine to create individual molecules.
combine to create individual molecules.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
line up with all the positive charges on one side.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
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self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
combine to create individual molecules.
self-assemble into a crystal lattice.
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combine to create individual molecules.
self-assemble into a crystal lattice.
line up with all the positive charges on one side.
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self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.

Q8: Ion crystal are made of individual molecules.

False
True
False
False
False
True
False
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True
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False
False
False
False
False
False
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False
False
False

Q9: In ionic compounds, the formula tells us the _____ of ions.

ratio
ratio
ratio
ratio
ratio
total number
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
total number
ratio
ratio
ratio
ratio
total number
ratio
ratio
charge
ratio

ratio
charge
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
ratio
size
ratio
ratio
total number
ratio
ratio
ratio
ratio

Q10: When Ca+2 ions and F-1 ions bond together, the calcium to fluoride ratio is 4:8, which reduces to a 1:2 ratio.
True
False
True
True
True
True

True
True
True
True
True
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True
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True
True
True
True
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False
True
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True
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True

True
True
True
True
True
True
True
True
True
True
True
False
False
True
True
True
True
True

Q11: Magnesium forms a +2 ion. Chlorine forms a +1 ion. What is the chemical formula for magnesium chloride?
MgCl2
MgCl2
Mg2Cl
MgCl2
Mg2Cl
Mg2Cl
MgCl2
MgCl2
MgCl2
Mg2Cl
MgCl2
MgCl2
Mg2Cl
MgCl2
MgCl
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2
MgCl2
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MgCl2
MgCl2
Mg2Cl
MgCl2
MgCl2
MgCl2
Mg2Cl
Mg2Cl
Mg2Cl
Mg2Cl
Mg2Cl
MgCl
Mg2Cl
MgCl2

Q12: Sodium forms a +1 ion. Phosphorus forms a -3 ion. What is the chemical formula for sodium phosphide?
Na3P
S3P
Na3P
Na3P
NaP3
Na3P
Na3P
Na3P
NaP3
NaP3
NaP3
Na3P
NaP3
Na3P
S3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
S3P
Na3P
Na3P
Na3P
NaP3
SP3
Na3P
Na3P
Na3P
Na3P
NaP3
Na3P
NaP3
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
Na3P
SP3
NaP3
Na3P
Na3P
Na3P
Na3P
NaP3
Na3P

Q13: Which of the following pairs is ionically bonded?

Na and F
Na and F
Na and F
Na and F

H and Cl
N and O

Na and F
Na and F

H and Cl

Na and F
Na and F
Na and F

P and Cl

Na and F

P and Cl

Na and F
Na and F
Na and F
Na and F

Na and F
Na and F

Na and F
Na and F

Na and F

H and Cl

Na and F

N and O
N and O
H and Cl

Na and F
Na and F

Na and F

P and Cl

Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F
Na and F

Q14: The chemical formula for potassium chloride is

KCl
KCl2
KCl
KCl
KCl
KCl

KCl
KCl
KCl
K2Cl
KCl
KCl
KCl
KCl
KCl
KCl
KCl
KCl
KCl
KCl

KCl
KCl

KCl
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KCl
K2Cl
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KCl
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K2Cl
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KCl
KCl2
KCl
KCl
KCl
K2Cl
KCl
KCl
KCl
KCl
KCl2
KCl
KCl

Q15: The chemical formula for calcium bromide is	Number of Correct Answers
CaBr2	13
CaBr	9
CaBr2	14
CaBr2	15
CaBr	11
CaBr2	8
CaBr2	13
CaBr	12
CaBr2	11
CaBr2	10
CaBr	12
CaBr2	13
CaBr2	12
CaBr2	15
CaBr	9
CaBr2	15
CaBr2	13
CaBr2	15
Ca4Br8	13
CaBr2	12
CaBr2	14
CaBr2	13
CaBr2	14
	13
CaBr2	13
	8
CaBr2	14
CaBr2	13
CaBr2	15
CaBr	7
CaBr2	10
CaBr2	13
CaBr2	13
CaBr2	12
	12
CaBr	8
CaBr2	7
CaBr	9
CaBr2	15
Ca2Br	14
CaBr2	14
CaBr2	14
CaBr2	15
CaBr2	13
CaBr2	15
CaBr2	15
CaBr2	14
CaBr	8
CaBr2	12
CaBr2	12
CaBr2	12
CaBr2	14
Ca2Br	11
CaBr2	12
CaBr2	15

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Daniel McCourt	repel one another.
Garrett Keeney	repel one another.
brandon	repel one another.
lauren bear	repel one another.
henry jacob	repel one another.
Carly	repel one another.

have opposite charges.
have opposite charges.

have opposite charges.
have opposite charges.
have opposite charges.

metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.
metals on the left side of the periodic table.

gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.
gain electrons.

sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride
sodium chloride

decreases in size
decreases in size
has 8 valence electrons
decreases in size
decreases in size
decreases in size

self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.
self-assemble into a crystal lattice.

False
False
True
False
False
False

charge
ratio
ratio
ratio
ratio
ratio

True
True
False
True
True
True

Mg2Cl
MgCl2
Mg2Cl
MgCl2
MgCl2
MgCl2

Na3P
Na3P
NaP3
Na3P
Na3P
Na3P

Na and F
Na and F
Na and F
Na and F
Na and F
Na and F

KCl
KCl
KCl
KCl
KCl
KCl

CaBr2	13
CaBr2	15
CaBr2	9
CaBr2	15
CaBr2	15
CaBr2	15

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