

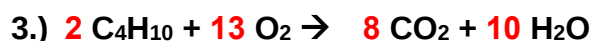
Balance the following chemical equations and indicate the type of reaction.



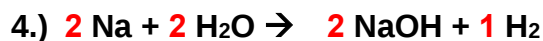
Type: **Synthesis**
(only 1 product)



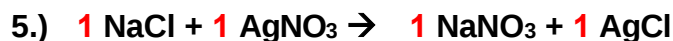
Type: **Decomposition**
(only 1 reactant)



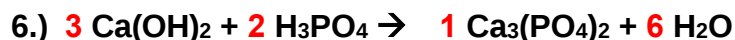
Type: **Combustion**
(look for O_2 , CO_2 , and H_2O)



Type: **Single Replacement**
(single element on both sides)



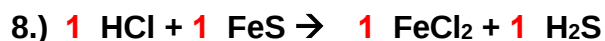
Type: **Double Replacement**
(all 4 are ionic compounds)
(2 compounds on both sides)
DOES NOT MAKE WATER



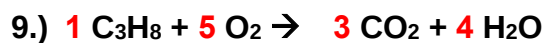
Type: **Neutralization**
(Double Replacement that
DOES MAKE WATER)



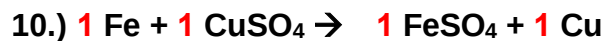
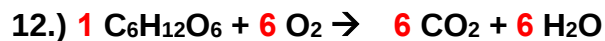
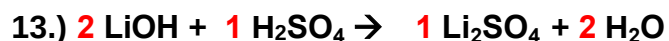
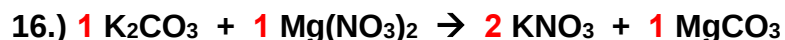
Type: **Synthesis**



Type: **Double Replacement**



Type: **Combustion**

Type: **Single Replacement**Type: **Synthesis**Type: **Combustion**Type: **Neutralization**Type: **Single Replacement**Type: **Decomposition**Type: **Double Replacement**

17.) Write the balanced chemical equation for the following:

Solid sodium metal reacts with liquid water to form aqueous sodium hydroxide and hydrogen gas.



18.) Write the balanced chemical equation for the following:

Methane burns in the presence of oxygen gas to form carbon dioxide gas and water vapor.

