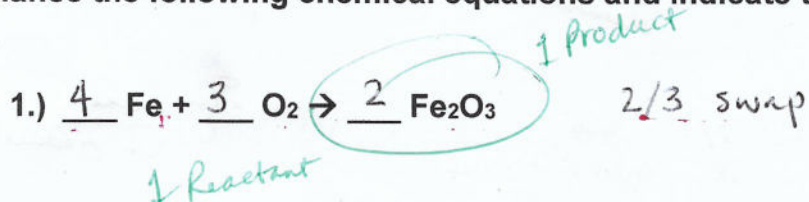


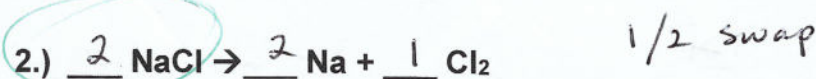
Unit 7: Chemical Formulas and Reactions
Intro to Balancing Chemical Equations

Name Key
Block _____

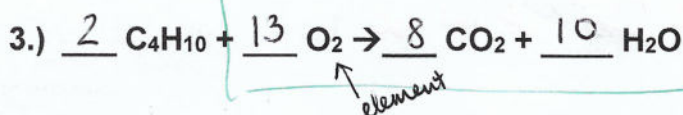
Balance the following chemical equations and indicate the type of reaction.



Type: Synthesis *(1 Product)*

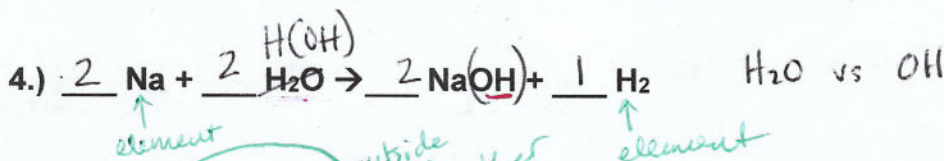


Type: Decomp *(1 Reactant)*

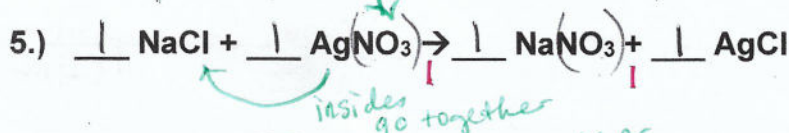


*swap H's
balance C's
adjust O's
simplify*

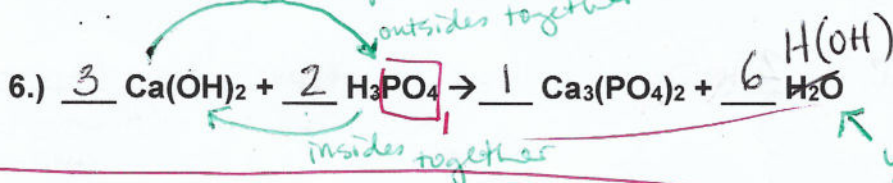
Type: Combustion *(O₂, CO₂, H₂O)*



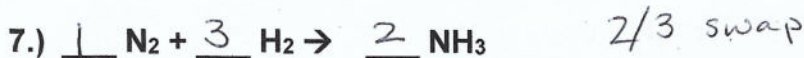
Type: Single Replacement *(element + compound)*



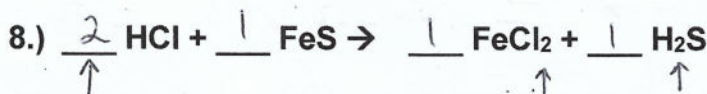
Type: Double Replacement *(4 compounds No water)*



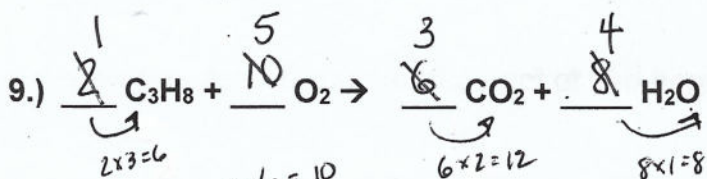
Type: Neutralization *(4 compounds makes water)*



Type: Synthesis



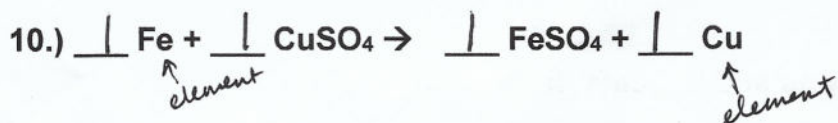
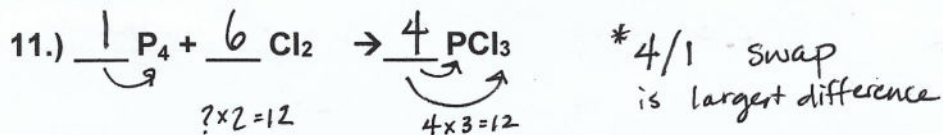
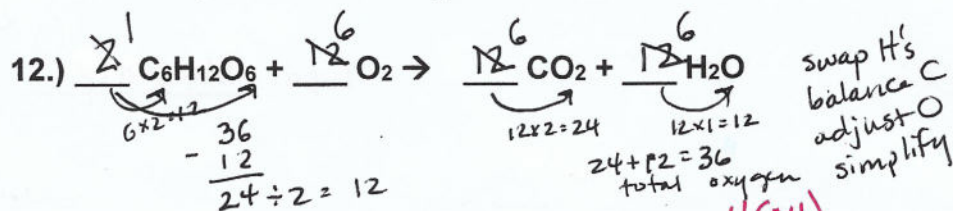
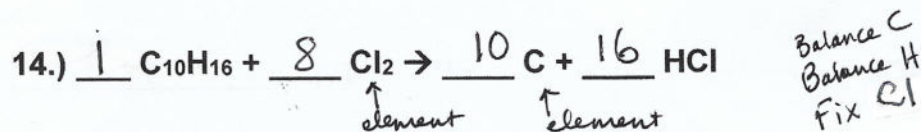
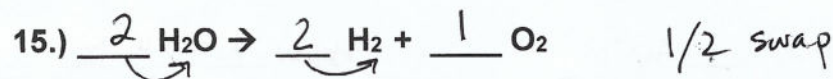
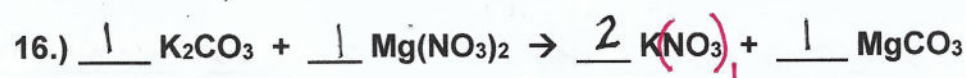
Type: Double Replacement



*swap H's
balance C's
adjust O
simplify*

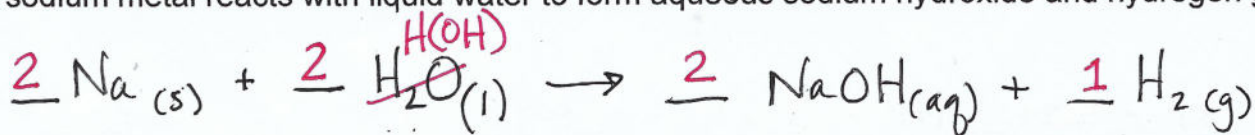
Type: Combustion

$12 + 8 = 20$

Type: Single Replacement
(element + comp)Type: Synthesis
(1 product)Type: Combustion
(O₂, CO₂, H₂O)Type: Neutralization
(acid + base → salt + water
4 compounds/makes water)Type: Single Replacement
(element + compound)Type: Decomposition
(1 reactant)Type: Double Replacement
(4 compounds)
no water

17.) Write the balanced chemical equation for the following:

(s) Solid sodium metal reacts with (l) liquid water to form (aq) aqueous sodium hydroxide and (g) hydrogen gas.



18.) Write the balanced chemical equation for the following:

(g) Methane gas burns in the presence of (g) oxygen gas to form (g) carbon dioxide gas and (g) water vapor.

