Electronegativity

Played on	**12 Nov 2019**
Hosted by	**JenKrug**
Played with	**23 players**
Played	**10 of 10**

Overall Performance

Total correct answers (%)	64,35%
Total incorrect answers (%)	35,65%
Average score (points)	5928,7

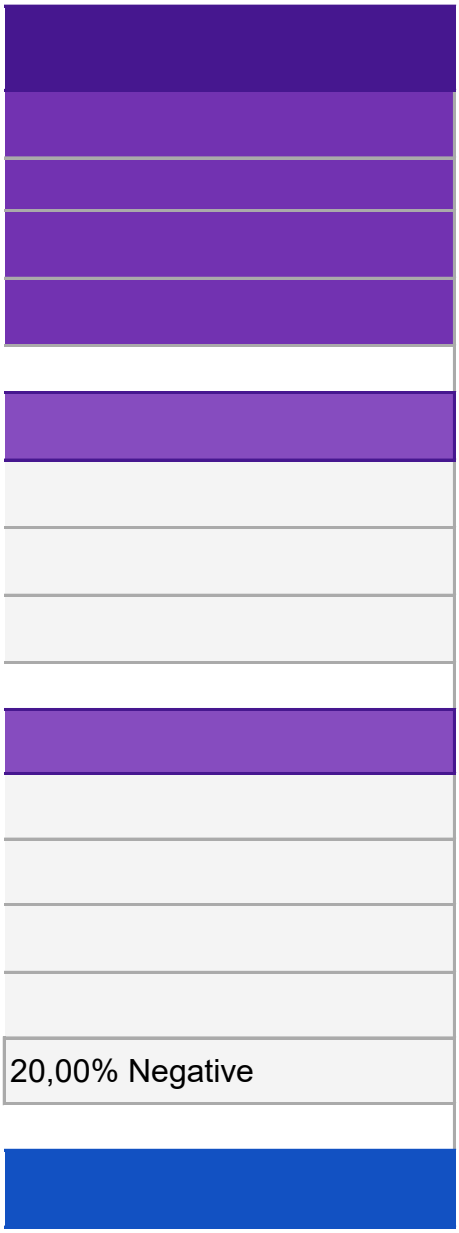
Feedback

Number of responses	5
How fun was it? (out of 5)	3,40 o
Did you learn something?	66,67%
Do you recommend it?	66,67%
How do you feel?	◉

Switch tabs/pages to view other result breakdown

%
%
74 points

ut of 5			
% Yes		33,33% No	
% Yes		33,33% No	
40,00% Positive	◉	40,00% Neutral	◉


20,00% Negative

Electronegativity

Final Scores

Rank	Players
1	caroline
2	aidan
3	Hudson
4	aj
5	braxton
6	reagan
7	Maddie
8	Cole
9	netta
10	Kyle
11	Karlee
12	kaitlyn
13	Shepard
14	bri
15	Parker
16	Kyle Daniels
17	Noah
18	Tyler the Great
19	Chris
20	Sam Sweetser
21	Mikayla :-)
22	Mithil

Final Scores

23	adelle

Total Score (points)	Correct Answers	Incorrect Answers
10232	9	1
10227	9	1
10224	9	1
8709	9	1
7956	9	1
7592	8	2
7517	8	2
7220	8	2
6652	7	3
6117	7	3
5726	7	3
5705	7	3
5695	7	3
5199	5	5
4889	6	4
4164	5	5
4057	5	5
3826	4	6
3789	5	5
3460	5	5
3066	4	6
2611	3	7

Final Scores

1728	2	8

Electronegativity

Kahoot! Summary

Rank	Players
1	caroline
2	aidan
3	Hudson
4	aj
5	braxton
6	reagan
7	Maddie
8	Cole
9	netta
10	Kyle
11	Karlee
12	kaitlyn
13	Shepard
14	bri
15	Parker

16	Kyle Daniels
17	Noah
18	Tyler the Great
19	Chris
20	Sam Sweetser
21	Mikayla :-)
22	Mithil
23	adelle

Total Score (points)	Q1
10232	823
10227	850
10224	893
8709	905
7956	603
7592	823
7517	803
7220	788
6652	908
6117	845
5726	705
5705	520
5695	735
5199	828
4889	803

Kahoot! Summary

4164	0
4057	808
3826	0
3789	730
3460	0
3066	763
2611	0
1728	815

Define the word "electronegativity"	Q2
a measure of the tendency to attract a bonding pair of electrons	1045
a measure of the tendency to attract a bonding pair of electrons	1048
a measure of the tendency to attract a bonding pair of electrons	1015
a measure of the tendency to attract a bonding pair of electrons	805
a measure of the tendency to attract a bonding pair of electrons	615
a measure of the tendency to attract a bonding pair of electrons	1003
a measure of the tendency to attract a bonding pair of electrons	950
a measure of the tendency to attract a bonding pair of electrons	783
a measure of the tendency to attract a bonding pair of electrons	1005
a measure of the tendency to attract a bonding pair of electrons	0
a measure of the tendency to attract a bonding pair of electrons	0
a measure of the tendency to attract a bonding pair of electrons	0
a measure of the tendency to attract a bonding pair of electrons	688
a measure of the tendency to attract a bonding pair of electrons	0
a measure of the tendency to attract a bonding pair of electrons	0

Kahoot! Summary

The electric charge of an atom	0
a measure of the tendency to attract a bonding pair of electrons	863
The electric charge of an atom	0
a measure of the tendency to attract a bonding pair of electrons	840
the least amount of energy needed to remove the most loosely bound electron	560
a measure of the tendency to attract a bonding pair of electrons	835
The electric charge of an atom	0
a measure of the tendency to attract a bonding pair of electrons	0

How is electronegativity measured?	Q3
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in Joules (J)	0
by the number of shells an atom has	0
by the number of shells an atom has	0
in the Pauling Scale	0
by the number of shells an atom has	0
in positive (+) or negative (-) charge	0

by the number of shells an atom has	0
in the Pauling Scale	0
in positive (+) or negative (-) charge	0
in the Pauling Scale	0
in the Pauling Scale	0
in the Pauling Scale	0
in Joules (J)	0
in positive (+) or negative (-) charge	0

Which element has the most electricity?	Q4
Fluorine	938
Fluorine	860
Fluorine	940
Fluorine	973
Fluorine	885
Fluorine	920
Fluorine	948
Fluorine	938
Fluorine	968
Fluorine	953
Fluorine	828
Fluorine	813
Fluorine	858
Fluorine	940
Fluorine	700

Kahoot! Summary

Fluorine	933
Fluorine	770
Radium	0
	853
Fluorine	765
Fluorine	553
Helium	938
Fluorine	913

Fluorine has the least electronegativity	Q5
False	1050
False	1055
False	1063
False	660
False	880
False	0
False	0
False	1055
False	1068
False	0
False	0
False	765
False	0
False	1073
False	1055

Kahoot! Summary

False	0
False	0
True	0
False	0
False	0
False	0
False	0
False	0

What is the range of the Pauling Scale?	Q6
0 to 4	1160
0 to 4	1153
0 to 4	1135
0 to 4	763
0 to 4	705
1 to 4	888
1 to 4	863
0 to 4	750
0 to 4	0
	530
1 to 4	555
0 to 4	758
1 to 4	615
0 to 4	1095
0 to 4	0

Kahoot! Summary

-4 to 4	923
1 to 4	0
0 to 9	0
1 to 4	0
1 to 4	510
1 to 4	0
1 to 4	0
	0

The _____ of two bonded atoms' electronegativity determines which type of bond they will form	Q7
difference	1268
difference	1273
difference	1285
difference	1195
difference	1095
difference	1023
difference	990
difference	1230
	980
difference	960
difference	910
difference	1158
difference	0
difference	1263
	833

Kahoot! Summary

difference	0
average	863
product	760
average	758
difference	0
	915
product	840
	0

Ionization energy is directly proportional to electronegativity	Q8
True	1268
True	1365
True	1348
True	1275
True	1143
True	865
True	820
True	1163
True	1053
True	1048
True	950
True	1158
False	908
True	0
True	633

Kahoot! Summary

False	740
True	753
True	923
True	608
False	865
True	0
True	0
	0

How does electronegativity increase on the periodic table?	Q9
up and right	1405
up and right	1423
up and right	1440
up and right	1053
up and right	1015
up and right	1095
up and right	1195
up and right	0
up and right	0
up and right	838
up and right	823
up and right	0
up and right	983
down and right	0
up and right	865

Kahoot! Summary

up and right	653
up and right	0
up and right	1013
up and right	0
up and right	760
down and left	0
up and left	833
	0

Why don't noble gasses have electronegativity	Q10
they have full valence orbitals	1275
they have full valence orbitals	1200
they are stable	1105
they are stable	1080
they have full valence orbitals	1015
they have full valence orbitals	975
they have full valence orbitals	948
	513
their electronegativity cannot be measured	670
they are stable	943
they are stable	955
	533
they are stable	908
their electronegativity cannot be measured	0
they are stable	0

Kahoot! Summary

they have full valence orbitals	915
their electronegativity cannot be measured	0
they are stable	1130
their electronegativity cannot be measured	0
they are stable	0
their electronegativity cannot be measured	0
they have full valence orbitals	0
	0

Why do Cesium and Francium have high electronegativity?
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals and a nearly full valence orbital
They have many orbitals and a nearly full valence orbital

They have many orbitals but few valence electrons
They have many orbitals but few valence electrons
They have many orbitals and a nearly full valence orbital
They have the most orbitals

Electrone
1 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

Define the word "electronegativity"

s	a measu
(%)	82,61%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✓
	✓
	✓
	✓
	✓
	✗
	✓
	✓
	✗
	✓

1 Quiz

	☑
	✗
	☑
	✗
	☑
	☑
	☑
	☑
	☑
	☑
	☑
	☑

re of the tendency to attract a bonding pair of electrons
nds

A measure of the size of an atom	♦
✗	
0	
0,00	

	Score (p
a measure of the tendency to attract a bonding pair of electrons	730
a measure of the tendency to attract a bonding pair of electrons	788
a measure of the tendency to attract a bonding pair of electrons	893
a measure of the tendency to attract a bonding pair of electrons	705
a measure of the tendency to attract a bonding pair of electrons	845
The electric charge of an atom	0
a measure of the tendency to attract a bonding pair of electrons	803
a measure of the tendency to attract a bonding pair of electrons	763
The electric charge of an atom	0
a measure of the tendency to attract a bonding pair of electrons	808

a measure of the tendency to attract a bonding pair of electrons	803
the least amount of energy needed to remove the most loosely bound electron	0
a measure of the tendency to attract a bonding pair of electrons	735
The electric charge of an atom	0
a measure of the tendency to attract a bonding pair of electrons	815
a measure of the tendency to attract a bonding pair of electrons	850
a measure of the tendency to attract a bonding pair of electrons	905
a measure of the tendency to attract a bonding pair of electrons	603
a measure of the tendency to attract a bonding pair of electrons	828
a measure of the tendency to attract a bonding pair of electrons	823
a measure of the tendency to attract a bonding pair of electrons	520
a measure of the tendency to attract a bonding pair of electrons	908
a measure of the tendency to attract a bonding pair of electrons	823

The electric charge of an atom	●
✗	
3	
10,27	

oints)	Current
	730
	788
	893
	705
	845
	0
	803
	763
	0
	808

	803
	0
	735
	0
	815
	850
	905
	603
	828
	823
	520
	908
	823

the least amount of energy needed to remove the most loosely bound electron	■
✗	
1	
10,50	

Total Score (points)	Answer ti
	10,8
	8,5
	4,3
	11,8
	6,2
	19,4
	7,9
	9,5
	5,9
	7,7

	7,9
	10,5
	10,6
	5,5
	7,4
	6
	3,8
	15,9
	6,9
	7,1
	19,2
	3,7
	7,1

a measure of the tendency to attract a bonding pair of electrons

19

8,54

me (seconds)

Electrone
2 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

How is electronegativity measured?

s	in the Pa
(%)	60,87%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✓
	✓
	✓
	✗
	✗
	✗
	✓
	✓
	✗
	✓

	✗
	☑
	☑
	✗
	✗
	☑
	☑
	☑
	✗
	☑
	✗
	☑
	☑

uling Scale
nds

by the number of shells an atom has	♦
✗	
4	
17,78	

	Score (p
in the Pauling Scale	840
in the Pauling Scale	783
in the Pauling Scale	1015
by the number of shells an atom has	0
in Joules (J)	0
by the number of shells an atom has	0
in the Pauling Scale	950
in the Pauling Scale	835
in Joules (J)	0
in the Pauling Scale	863

in positive (+) or negative (-) charge	0
in the Pauling Scale	560
in the Pauling Scale	688
in positive (+) or negative (-) charge	0
in positive (+) or negative (-) charge	0
in the Pauling Scale	1048
in the Pauling Scale	805
in the Pauling Scale	615
by the number of shells an atom has	0
in the Pauling Scale	1045
by the number of shells an atom has	0
in the Pauling Scale	1005
in the Pauling Scale	1003

in Joules (J)
✗
2
10,60

oints)	Current
	1570
	1571
	1908
	705
	845
	0
	1753
	1598
	0
	1671

	803
	560
	1423
	0
	815
	1898
	1710
	1218
	828
	1868
	520
	1913
	1826

in positive (+) or negative (-) charge	■
✗	
3	
13,63	

Total Score (points)	Answer ti
	10,4
	12,7
	3,4
	19,1
	15,4
	15,7
	6
	10,6
	5,8
	9,5

	18,9
	17,6
	16,5
	11,3
	10,7
	2,1
	11,8
	19,4
	16,6
	2,2
	19,7
	3,8
	3,9

in the Pauling Scale
√
14
9,28
me (seconds)

Electrone
3 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

Which element has the most electricity?

s	Cesium
(%)	0,00%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗

	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗
	✗

nds

Helium	♦
✗	
1	
15,70	

	Score (p
	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Helium	0
Fluorine	0

Fluorine	0
Fluorine	0
Fluorine	0
Radium	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0
Fluorine	0

Fluorine	●
✗	
20	
6,06	

oints)	Current
	1570
	1571
	1908
	705
	845
	0
	1753
	1598
	0
	1671

	803
	560
	1423
	0
	815
	1898
	1710
	1218
	828
	1868
	520
	1913
	1826

Radium	■
✗	
1	
7,50	

Total Score (points)	Answer ti
	20
	9
	1,3
	6
	12,1
	1,6
	5,4
	6,4
	15,7
	3,3

	7,3
	10,8
	2
	7,5
	2,9
	4,9
	11,3
	12,3
	3,7
	2
	13,1
	1,9
	3,9

Cesium
✓
0
0,00
ime (seconds)

Electrone
4 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

Fluorine has the least electronegativity

s	False
(%)	95,65%
on	20 secor

mmary

	▲
ct?	
wers received	
ken to answer (seconds)	

ails

	Answer
	☑
	☑
	☑
	☑
	☑
	☑
	☑
	☑
	☑
	☑

	☑
	☑
	☑
	✗
	☑
	☑
	☑
	☑
	☑
	☑
	☑
	☑

nds

False	♦
✓	
22	
5,13	

	Score (p
False	853
False	938
False	940
False	828
False	953
False	933
False	948
False	553
False	938
False	770

False	700
False	765
False	858
True	0
False	913
False	860
False	973
False	885
False	940
False	938
False	813
False	968
False	920

True	●
✗	
1	
0,80	

oints)	Current
	2423
	2509
	2848
	1533
	1798
	933
	2701
	2151
	938
	2441

	1503
	1325
	2281
	0
	1728
	2758
	2683
	2103
	1768
	2806
	1333
	2881
	2746

Total Score (points)	Answer ti
	5,9
	2,5
	2,4
	6,9
	1,9
	2,7
	2,1
	17,9
	2,5
	9,2

	12
	9,4
	5,7
	0,8
	3,5
	5,6
	1,1
	4,6
	2,4
	2,5
	7,5
	1,3
	3,2

ime (seconds)

Electrone
5 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

What is the range of the Pauling Scale?

s	0 to 4
(%)	43,48%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✗
	☑
	☑
	✗
	✗
	✗
	✗
	✗
	✗
	✗

	☑
	✗
	✗
	✗
	✗
	☑
	☑
	☑
	☑
	☑
	☑
	☑
	✗

nds

0 to 4	♦
✓	
10	
5,11	

	Score (p
1 to 4	0
0 to 4	1055
0 to 4	1063
1 to 4	0
	0
-4 to 4	0
1 to 4	0
1 to 4	0
1 to 4	0
1 to 4	0

0 to 4	1055
1 to 4	0
1 to 4	0
0 to 9	0
	0
0 to 4	1055
0 to 4	660
0 to 4	880
0 to 4	1073
0 to 4	1050
0 to 4	765
0 to 4	1068
1 to 4	0

-4 to 4	●
✗	
1	
17,40	

oints)	Current
	2423
	3564
	3911
	1533
	1798
	933
	2701
	2151
	938
	2441

	2558
	1325
	2281
	0
	1728
	3813
	3343
	2983
	2841
	3856
	2098
	3949
	2746

1 to 4	■
✗	
9	
4,41	

Total Score (points)	Answer ti
	7
	1,8
	1,5
	5,8
	20
	17,4
	2,7
	5
	4,4
	3

	1,8
	6,9
	3,1
	6,2
	20
	1,8
	17,6
	8,8
	1,1
	2
	13,4
	1,3
	1,8

0 to 9
✗
1
6,20
ime (seconds)

Electrone
6 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

The _____ of two bonded atoms' electronegativity determines wh	
s	differenc
(%)	65,22%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✗
	☑
	☑
	☑
	☑
	☑
	☑
	✗
	✗
	✗

	✗
	☑
	☑
	✗
	✗
	☑
	☑
	☑
	☑
	☑
	☑
	✗
	☑

hich type of bond they will form
ce
nds

sum	♦
✗	
0	
0,00	

	Score (p
average	0
difference	750
difference	1135
difference	555
difference	530
difference	923
difference	863
	0
product	0
average	0

	0
difference	510
difference	615
product	0
	0
difference	1153
difference	763
difference	705
difference	1095
difference	1160
difference	758
	0
difference	888

difference	
✓	
15	
11,20	

oints)	Current
	2423
	4314
	5046
	2088
	2328
	1856
	3564
	2151
	938
	2441

	2558
	1835
	2896
	0
	1728
	4966
	4106
	3688
	3936
	5016
	2856
	3949
	3634

average	■
✗	
2	
9,60	

Total Score (points)	Answer ti
	11,5
	18
	2,6
	17,8
	18,8
	3,1
	5,5
	20
	10,2
	7,7

	20
	19,6
	15,4
	12,4
	20
	1,9
	17,5
	19,8
	4,2
	1,6
	17,7
	20
	4,5

product

✗

2

11,30

me (seconds)

Electrone
7 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

lonization energy is directly proportional to electronegativity

s	True
(%)	82,61%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✓
	✓
	✓
	✓
	✓
	✗
	✓
	✓
	✓
	✓

	☑
	✗
	✗
	☑
	✗
	☑
	☑
	☑
	☑
	☑
	☑
	☑
	☑

nds

False	♦
✗	
3	
7,17	

	Score (p
True	758
True	1230
True	1285
True	910
True	960
False	0
True	990
True	915
True	840
True	863

True	833
False	0
False	0
True	760
	0
True	1273
True	1195
True	1095
True	1263
True	1268
True	1158
True	980
True	1023

True
✓
19
4,64

oints)	Current
	3181
	5544
	6331
	2998
	3288
	1856
	4554
	3066
	1778
	3304

	3391
	1835
	2896
	760
	1728
	6239
	5301
	4783
	5199
	6284
	4014
	4929
	4657

Total Score (points)	Answer ti
	9,7
	2,8
	0,6
	7,6
	5,6
	5,2
	4,4
	3,4
	6,4
	5,5

	6,7
	12,3
	4
	9,6
	20
	1,1
	4,2
	8,2
	1,5
	1,3
	5,7
	0,8
	3,1

ime (seconds)

Electrone
8 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

egativity

How does electronegativity increase on the periodic table?

s	up and ri
(%)	82,61%
on	20 secor

mmary

	▲
ct?	
wers received	
ken to answer (seconds)	

ails

	Answer
	✓
	✓
	✓
	✓
	✓
	✓
	✓
	✗
	✗
	✓

	☑
	☑
	☑
	☑
	✗
	☑
	☑
	☑
	✗
	☑
	☑
	☑
	☑

ight
nds

down and left	♦
✗	
1	
7,30	

	Score (p
up and right	608
up and right	1163
up and right	1348
up and right	950
up and right	1048
up and right	740
up and right	820
down and left	0
up and left	0
up and right	753

up and right	633
up and right	865
up and right	908
up and right	923
	0
up and right	1365
up and right	1275
up and right	1143
down and right	0
up and right	1268
up and right	1158
up and right	1053
up and right	865

up and right	●
✓	
19	
8,88	

oints)	Current
	3789
	6707
	7679
	3948
	4336
	2596
	5374
	3066
	1778
	4057

	4024
	2700
	3804
	1683
	1728
	7604
	6576
	5926
	5199
	7552
	5172
	5982
	5522

down and right	■
✗	
1	
3,40	

Total Score (points)	Answer ti
	19,7
	9,5
	2,1
	10
	6,1
	10,4
	15,2
	7,3
	4,6
	13,9

	18,7
	5,4
	3,7
	7,1
	20
	1,4
	5
	10,3
	3,4
	5,3
	9,7
	1,9
	13,4

up and left
✗
1
4,60
me (seconds)

Electrone
9 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

Why don't noble gasses have electronegativity

s	they hav
(%)	65,22%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✗
	✗
	☑
	☑
	☑
	☑
	☑
	✗
	☑
	✗

	☑
	☑
	☑
	☑
	✗
	☑
	☑
	☑
	✗
	☑
	✗
	✗
	☑

e full valence orbitals, they are stable
nds

they have full valence orbitals	♦
√	
7	
9,04	

	Score (p
their electronegativity cannot be measured	0
	0
they are stable	1440
they are stable	823
they are stable	838
they have full valence orbitals	653
they have full valence orbitals	1195
their electronegativity cannot be measured	0
they have full valence orbitals	833
their electronegativity cannot be measured	0

they are stable	865
they are stable	760
they are stable	983
they are stable	1013
	0
they have full valence orbitals	1423
they are stable	1053
they have full valence orbitals	1015
their electronegativity cannot be measured	0
they have full valence orbitals	1405
	0
their electronegativity cannot be measured	0
they have full valence orbitals	1095

they can't have negative charges	●
✗	
0	
0,00	

oints)	Current
	3789
	6707
	9119
	4771
	5174
	3249
	6569
	3066
	2611
	4057

	4889
	3460
	4787
	2696
	1728
	9027
	7629
	6941
	5199
	8957
	5172
	5982
	6617

they are stable	■
✓	
8	
12,14	

Total Score (points)	Answer ti
	17,8
	20
	2,4
	19,1
	18,5
	17,9
	4,2
	13
	6,7
	13,6

	13,4
	13,6
	4,7
	7,5
	20
	3,1
	17,9
	19,4
	8
	3,8
	20
	6,9
	8,2

their electronegativity cannot be measured
✗
5
11,86
me (seconds)

Electrone
10 Quiz
Correct answers
Players correct (
Question duratic
Answer Sun
Answer options
Is answer correc
Number of answ
Average time tal
Answer Det
Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah

Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

gativity

Why do Cesium and Francium have high electronegativity?

s	They hav
(%)	65,22%
on	20 secor

mmary

	▲
ct?	
vers received	
ken to answer (seconds)	

ails

	Answer
	✗
	☑
	☑
	☑
	☑
	☑
	☑
	✗
	✗
	✗

	✗
	✗
	☑
	☑
	✗
	☑
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	✗
	☑
	☑
	☑
	☑

ve many orbitals but few valence electrons
nds

They have many orbitals and a nearly full valence orbital	♦
✗	
3	
15,43	

	Score (p
	0
They have many orbitals but few valence electrons	513
They have many orbitals but few valence electrons	1105
They have many orbitals but few valence electrons	955
They have many orbitals but few valence electrons	943
They have many orbitals but few valence electrons	915
They have many orbitals but few valence electrons	948
They have many orbitals and a nearly full valence orbital	0
They have the most orbitals	0
	0

They have many orbitals and a nearly full valence orbital	0
	0
They have many orbitals but few valence electrons	908
They have many orbitals but few valence electrons	1130
	0
They have many orbitals but few valence electrons	1200
They have many orbitals but few valence electrons	1080
They have many orbitals but few valence electrons	1015
They have many orbitals and a nearly full valence orbital	0
They have many orbitals but few valence electrons	1275
They have many orbitals but few valence electrons	533
They have many orbitals but few valence electrons	670
They have many orbitals but few valence electrons	975

They have the most orbitals	●
✗	
1	
2,90	

oints)	Current
	3789
	7220
	10224
	5726
	6117
	4164
	7517
	3066
	2611
	4057

	4889
	3460
	5695
	3826
	1728
	10227
	8709
	7956
	5199
	10232
	5705
	6652
	7592

They have many orbitals but few valence electrons	■
✓	
15	
15,03	

Total Score (points)	Answer ti
	20
	19,5
	15,8
	17,8
	18,3
	11,4
	18,1
	19,1
	2,9
	20

	17,1
	20
	11,7
	6,8
	20
	12
	16,8
	19,4
	10,1
	9
	18,7
	13,2
	17

They have the most electrons
✗
0
0,00
me (seconds)

Question Number
1 Quiz
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Question
Define the word "electronegativity"
Define the word "electronegativity"
Define the word "electronegativity"
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Define the word "electronegativity"
Define the word "electronegativity"
Define the word "electronegativity"
Define the word "electronegativity"
Define the word "electronegativity"
How is electronegativity measured?
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How is electronegativity measured?
How is electronegativity measured?
Which element has the most electricity?

Which element has the most electricity?
Which element has the most electricity?
Which element has the most electricity?
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Which element has the most electricity?
Which element has the most electricity?
Which element has the most electricity?
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Which element has the most electricity?
Which element has the most electricity?
Fluorine has the least electronegativity
Fluorine has the least electronegativity
Fluorine has the least electronegativity
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Fluorine has the least electronegativity
Fluorine has the least electronegativity
What is the range of the Pauling Scale?
What is the range of the Pauling Scale?
What is the range of the Pauling Scale?

What is the range of the Pauling Scale?
What is the range of the Pauling Scale?
What is the range of the Pauling Scale?
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What is the range of the Pauling Scale?

What is the range of the Pauling Scale?
What is the range of the Pauling Scale?
What is the range of the Pauling Scale?
What is the range of the Pauling Scale?
The ______ of two bonded atoms' electronegativity determines which type of bond they will form
The ______ of two bonded atoms' electronegativity determines which type of bond they will form
The ______ of two bonded atoms' electronegativity determines which type of bond they will form
The ______ of two bonded atoms' electronegativity determines which type of bond they will form
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The ______ of two bonded atoms' electronegativity determines which type of bond they will form
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The ______ of two bonded atoms' electronegativity determines which type of bond they will form
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Ionization energy is directly proportional to electronegativity
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Ionization energy is directly proportional to electronegativity
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Ionization energy is directly proportional to electronegativity
Ionization energy is directly proportional to electronegativity
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
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How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
How does electronegativity increase on the periodic table?
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity

Why don't noble gasses have electronegativity
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Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity
Why don't noble gasses have electronegativity

Why do Cesium and Francium have high electronegativity?
Why do Cesium and Francium have high electronegativity?
Why do Cesium and Francium have high electronegativity?
Why do Cesium and Francium have high electronegativity?
Why do Cesium and Francium have high electronegativity?
Why do Cesium and Francium have high electronegativity?
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Why do Cesium and Francium have high electronegativity?
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Why do Cesium and Francium have high electronegativity?
Why do Cesium and Francium have high electronegativity?
Why do Cesium and Francium have high electronegativity?

Answer 1	Answer 2
A measure of the size of an atom	The electric charge of an atom
A measure of the size of an atom	The electric charge of an atom
A measure of the size of an atom	The electric charge of an atom
A measure of the size of an atom	The electric charge of an atom
A measure of the size of an atom	The electric charge of an atom
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A measure of the size of an atom	The electric charge of an atom

A measure of the size of an atom	The electric charge of an atom
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A measure of the size of an atom	The electric charge of an atom
A measure of the size of an atom	The electric charge of an atom
A measure of the size of an atom	The electric charge of an atom
A measure of the size of an atom	The electric charge of an atom
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)

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by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
by the number of shells an atom has	in Joules (J)
Helium	Fluorine

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Helium	Fluorine
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Helium	Fluorine
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Helium	Fluorine
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0 to 4	-4 to 4
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0 to 4	-4 to 4
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False	True
False	True
down and left	up and right
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they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges

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they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges
they have full valence orbitals	they can't have negative charges

They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
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They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals

They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals
They have many orbitals and a nearly full valence orbital	They have the most orbitals

Answer 3	Answer 4
the least amount of energy needed to remove the most loosely bound electron	a measure of the tendency to attract a bonding pair of electrons
the least amount of energy needed to remove the most loosely bound electron	a measure of the tendency to attract a bonding pair of electrons
the least amount of energy needed to remove the most loosely bound electron	a measure of the tendency to attract a bonding pair of electrons
the least amount of energy needed to remove the most loosely bound electron	a measure of the tendency to attract a bonding pair of electrons
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the least amount of energy needed to remove the most loosely bound electron	a measure of the tendency to attract a bonding pair of electrons
the least amount of energy needed to remove the most loosely bound electron	a measure of the tendency to attract a bonding pair of electrons
the least amount of energy needed to remove the most loosely bound electron	a measure of the tendency to attract a bonding pair of electrons
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale

in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
in positive (+) or negative (-) charge	in the Pauling Scale
Radium	Cesium

Radium	Cesium
Radium	Cesium
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Radium	Cesium
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Radium	Cesium
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Radium	Cesium
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Radium	Cesium
Radium	Cesium

1 to 4	0 to 9
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average	product
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down and right	up and left
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they are stable	their electronegativity cannot be measured
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They have many orbitals but few valence electrons	They have the most electrons
They have many orbitals but few valence electrons	They have the most electrons
They have many orbitals but few valence electrons	They have the most electrons
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They have many orbitals but few valence electrons	They have the most electrons
They have many orbitals but few valence electrons	They have the most electrons

Correct Answers	Time Allotted to Answer (seconds)
a measure of the tendency to attract a bonding pair of electrons	20
a measure of the tendency to attract a bonding pair of electrons	20
a measure of the tendency to attract a bonding pair of electrons	20
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in the Pauling Scale	20
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they have full valence orbitals, they are stable	20
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They have many orbitals but few valence electrons	20
They have many orbitals but few valence electrons	20
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They have many orbitals but few valence electrons	20

Players
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah
Parker
Sam Sweetser
Shepard
Tyler the Great
adelle

aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan
Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)

Mithil
Noah
Parker
Sam Sweetser
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Chris

Cole
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Chris
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Hudson
Karlee
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Kyle Daniels
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Sam Sweetser
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Tyler the Great
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caroline
kaitlyn

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Chris
Cole
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Chris
Cole
Hudson
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Kyle
Kyle Daniels
Maddie

Mikayla :-)
Mithil
Noah
Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan
aj
braxton
bri
caroline
kaitlyn
netta
reagan

Chris
Cole
Hudson
Karlee
Kyle
Kyle Daniels
Maddie
Mikayla :-)
Mithil
Noah
Parker
Sam Sweetser
Shepard
Tyler the Great
adelle
aidan

aj
braxton
bri
caroline
kaitlyn
netta
reagan

Answer	Correct / Incorrect	Correct
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
The electric charge of an atom	Incorrect	0
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
The electric charge of an atom	Incorrect	0
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
the least amount of energy needed to remove the most loosely bound electron	Incorrect	0
a measure of the tendency to attract a bonding pair of electrons	Correct	1
The electric charge of an atom	Incorrect	0
a measure of the tendency to attract a bonding pair of electrons	Correct	1

a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
a measure of the tendency to attract a bonding pair of electrons	Correct	1
in the Pauling Scale	Correct	1
in the Pauling Scale	Correct	1
in the Pauling Scale	Correct	1
by the number of shells an atom has	Incorrect	0
in Joules (J)	Incorrect	0
by the number of shells an atom has	Incorrect	0
in the Pauling Scale	Correct	1
in the Pauling Scale	Correct	1

in Joules (J)	Incorrect	0
in the Pauling Scale	Correct	1
in positive (+) or negative (-) charge	Incorrect	0
in the Pauling Scale	Correct	1
in the Pauling Scale	Correct	1
in positive (+) or negative (-) charge	Incorrect	0
in positive (+) or negative (-) charge	Incorrect	0
in the Pauling Scale	Correct	1
in the Pauling Scale	Correct	1
in the Pauling Scale	Correct	1
by the number of shells an atom has	Incorrect	0
in the Pauling Scale	Correct	1
by the number of shells an atom has	Incorrect	0
in the Pauling Scale	Correct	1
in the Pauling Scale	Correct	1
	Incorrect	0

Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Helium	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Radium	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0

Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
Fluorine	Incorrect	0
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1

False	Correct	1
False	Correct	1
False	Correct	1
True	Incorrect	0
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
1 to 4	Incorrect	0
0 to 4	Correct	1
0 to 4	Correct	1

1 to 4	Incorrect	0
	Incorrect	0
-4 to 4	Incorrect	0
1 to 4	Incorrect	0
1 to 4	Incorrect	0
1 to 4	Incorrect	0
1 to 4	Incorrect	0
0 to 4	Correct	1
1 to 4	Incorrect	0
1 to 4	Incorrect	0
0 to 9	Incorrect	0
	Incorrect	0
0 to 4	Correct	1
0 to 4	Correct	1
0 to 4	Correct	1
0 to 4	Correct	1

0 to 4	Correct	1
0 to 4	Correct	1
0 to 4	Correct	1
1 to 4	Incorrect	0
average	Incorrect	0
difference	Correct	1
difference	Correct	1
difference	Correct	1
difference	Correct	1
difference	Correct	1
difference	Correct	1
	Incorrect	0
product	Incorrect	0
average	Incorrect	0
	Incorrect	0
difference	Correct	1

difference	Correct	1
product	Incorrect	0
	Incorrect	0
difference	Correct	1
difference	Correct	1
difference	Correct	1
difference	Correct	1
difference	Correct	1
difference	Correct	1
	Incorrect	0
difference	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1

False	Incorrect	0
True	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1
False	Incorrect	0
False	Incorrect	0
True	Correct	1
	Incorrect	0
True	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1
True	Correct	1

True	Correct	1
True	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
down and left	Incorrect	0
up and left	Incorrect	0
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1

	Incorrect	0
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
down and right	Incorrect	0
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
up and right	Correct	1
their electronegativity cannot be measured	Incorrect	0
	Incorrect	0
they are stable	Correct	1
they are stable	Correct	1
they are stable	Correct	1
they have full valence orbitals	Correct	1
they have full valence orbitals	Correct	1

their electronegativity cannot be measured	Incorrect	0
they have full valence orbitals	Correct	1
their electronegativity cannot be measured	Incorrect	0
they are stable	Correct	1
they are stable	Correct	1
they are stable	Correct	1
they are stable	Correct	1
	Incorrect	0
they have full valence orbitals	Correct	1
they are stable	Correct	1
they have full valence orbitals	Correct	1
their electronegativity cannot be measured	Incorrect	0
they have full valence orbitals	Correct	1
	Incorrect	0
their electronegativity cannot be measured	Incorrect	0
they have full valence orbitals	Correct	1

	Incorrect	0
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals and a nearly full valence orbital	Incorrect	0
They have the most orbitals	Incorrect	0
	Incorrect	0
They have many orbitals and a nearly full valence orbital	Incorrect	0
	Incorrect	0
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
	Incorrect	0
They have many orbitals but few valence electrons	Correct	1

They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals and a nearly full valence orbital	Incorrect	0
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1
They have many orbitals but few valence electrons	Correct	1

Incorrect	Score (points)	Score without Answer Streak Bonus (points)
0	730	730
0	788	788
0	893	893
0	705	705
0	845	845
1	0	0
0	803	803
0	763	763
1	0	0
0	808	808
0	803	803
1	0	0
0	735	735
1	0	0
0	815	815

0	850	850
0	905	905
0	603	603
0	828	828
0	823	823
0	520	520
0	908	908
0	823	823
0	840	740
0	783	683
0	1015	915
1	0	0
1	0	0
1	0	0
0	950	850
0	835	735

1	0	0
0	863	763
1	0	0
0	560	560
0	688	588
1	0	0
1	0	0
0	1048	948
0	805	705
0	615	515
1	0	0
0	1045	945
1	0	0
0	1005	905
0	1003	903
1	0	0

1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0

1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
0	853	853
0	938	938
0	940	940
0	828	828
0	953	953
0	933	933
0	948	948
0	553	553
0	938	938
0	770	770

0	700	700
0	765	765
0	858	858
1	0	0
0	913	913
0	860	860
0	973	973
0	885	885
0	940	940
0	938	938
0	813	813
0	968	968
0	920	920
1	0	0
0	1055	955
0	1063	963

1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
0	1055	955
1	0	0
1	0	0
1	0	0
1	0	0
0	1055	955
0	660	560
0	880	780
0	1073	973

0	1050	950
0	765	665
0	1068	968
1	0	0
1	0	0
0	750	550
0	1135	935
0	555	555
0	530	530
0	923	923
0	863	863
1	0	0
1	0	0
1	0	0
1	0	0
0	510	510

0	615	615
1	0	0
1	0	0
0	1153	953
0	763	563
0	705	505
0	1095	895
0	1160	960
0	758	558
1	0	0
0	888	888
0	758	758
0	1230	930
0	1285	985
0	910	810
0	960	860

1	0	0
0	990	890
0	915	915
0	840	840
0	863	863
0	833	833
1	0	0
1	0	0
0	760	760
1	0	0
0	1273	973
0	1195	895
0	1095	795
0	1263	963
0	1268	968
0	1158	858

0	980	980
0	1023	923
0	608	508
0	1163	763
0	1348	948
0	950	750
0	1048	848
0	740	740
0	820	620
1	0	0
1	0	0
0	753	653
0	633	533
0	865	865
0	908	908
0	923	823

1	0	0
0	1365	965
0	1275	875
0	1143	743
1	0	0
0	1268	868
0	1158	758
0	1053	953
0	865	665
1	0	0
1	0	0
0	1440	940
0	823	523
0	838	538
0	653	553
0	1195	895

1	0	0
0	833	833
1	0	0
0	865	665
0	760	660
0	983	883
0	1013	813
1	0	0
0	1423	923
0	1053	553
0	1015	515
1	0	0
0	1405	905
1	0	0
1	0	0
0	1095	795

1	0	0
0	513	513
0	1105	605
0	955	555
0	943	543
0	915	715
0	948	548
1	0	0
1	0	0
1	0	0
1	0	0
1	0	0
0	908	708
0	1130	830
1	0	0
0	1200	700

0	1080	580
0	1015	515
1	0	0
0	1275	775
0	533	533
0	670	670
0	975	575

Current Total Score (points)	Answer Time (%)
730	54.00%
788	42.50%
893	21.50%
705	59.00%
845	31.00%
0	97.00%
803	39.50%
763	47.50%
0	29.50%
808	38.50%
803	39.50%
0	52.50%
735	53.00%
0	27.50%
815	37.00%

850	30.00%
905	19.00%
603	79.50%
828	34.50%
823	35.50%
520	96.00%
908	18.50%
823	35.50%
1570	52.00%
1571	63.50%
1908	17.00%
705	95.50%
845	77.00%
0	78.50%
1753	30.00%
1598	53.00%

0	29.00%
1671	47.50%
803	94.50%
560	88.00%
1423	82.50%
0	56.50%
815	53.50%
1898	10.50%
1710	59.00%
1218	97.00%
828	83.00%
1868	11.00%
520	98.50%
1913	19.00%
1826	19.50%
1570	100.00%

1571	45.00%
1908	6.50%
705	30.00%
845	60.50%
0	8.00%
1753	27.00%
1598	32.00%
0	78.50%
1671	16.50%
803	36.50%
560	54.00%
1423	10.00%
0	37.50%
815	14.50%
1898	24.50%
1710	56.50%

1218	61.50%
828	18.50%
1868	10.00%
520	65.50%
1913	9.50%
1826	19.50%
2423	29.50%
2509	12.50%
2848	12.00%
1533	34.50%
1798	9.50%
933	13.50%
2701	10.50%
2151	89.50%
938	12.50%
2441	46.00%

1503	60.00%
1325	47.00%
2281	28.50%
0	4.00%
1728	17.50%
2758	28.00%
2683	5.50%
2103	23.00%
1768	12.00%
2806	12.50%
1333	37.50%
2881	6.50%
2746	16.00%
2423	35.00%
3564	9.00%
3911	7.50%

1533	29.00%
1798	100.00%
933	87.00%
2701	13.50%
2151	25.00%
938	22.00%
2441	15.00%
2558	9.00%
1325	34.50%
2281	15.50%
0	31.00%
1728	100.00%
3813	9.00%
3343	88.00%
2983	44.00%
2841	5.50%

3856	10.00%
2098	67.00%
3949	6.50%
2746	9.00%
2423	57.50%
4314	90.00%
5046	13.00%
2088	89.00%
2328	94.00%
1856	15.50%
3564	27.50%
2151	100.00%
938	51.00%
2441	38.50%
2558	100.00%
1835	98.00%

2896	77.00%
0	62.00%
1728	100.00%
4966	9.50%
4106	87.50%
3688	99.00%
3936	21.00%
5016	8.00%
2856	88.50%
3949	100.00%
3634	22.50%
3181	48.50%
5544	14.00%
6331	3.00%
2998	38.00%
3288	28.00%

1856	26.00%
4554	22.00%
3066	17.00%
1778	32.00%
3304	27.50%
3391	33.50%
1835	61.50%
2896	20.00%
760	48.00%
1728	100.00%
6239	5.50%
5301	21.00%
4783	41.00%
5199	7.50%
6284	6.50%
4014	28.50%

4929	4.00%
4657	15.50%
3789	98.50%
6707	47.50%
7679	10.50%
3948	50.00%
4336	30.50%
2596	52.00%
5374	76.00%
3066	36.50%
1778	23.00%
4057	69.50%
4024	93.50%
2700	27.00%
3804	18.50%
1683	35.50%

1728	100.00%
7604	7.00%
6576	25.00%
5926	51.50%
5199	17.00%
7552	26.50%
5172	48.50%
5982	9.50%
5522	67.00%
3789	89.00%
6707	100.00%
9119	12.00%
4771	95.50%
5174	92.50%
3249	89.50%
6569	21.00%

3066	65.00%
2611	33.50%
4057	68.00%
4889	67.00%
3460	68.00%
4787	23.50%
2696	37.50%
1728	100.00%
9027	15.50%
7629	89.50%
6941	97.00%
5199	40.00%
8957	19.00%
5172	100.00%
5982	34.50%
6617	41.00%

3789	100.00%
7220	97.50%
10224	79.00%
5726	89.00%
6117	91.50%
4164	57.00%
7517	90.50%
3066	95.50%
2611	14.50%
4057	100.00%
4889	85.50%
3460	100.00%
5695	58.50%
3826	34.00%
1728	100.00%
10227	60.00%

8709	84.00%
7956	97.00%
5199	50.50%
10232	45.00%
5705	93.50%
6652	66.00%
7592	85.00%

Answer Time (seconds)
10,8
8,5
4,3
11,8
6,2
19,4
7,9
9,5
5,9
7,7
7,9
10,5
10,6
5,5
7,4

6
3,8
15,9
6,9
7,1
19,2
3,7
7,1
10,4
12,7
3,4
19,1
15,4
15,7
6
10,6

5,8
9,5
18,9
17,6
16,5
11,3
10,7
2,1
11,8
19,4
16,6
2,2
19,7
3,8
3,9
20

9
1,3
6
12,1
1,6
5,4
6,4
15,7
3,3
7,3
10,8
2
7,5
2,9
4,9
11,3

12,3
3,7
2
13,1
1,9
3,9
5,9
2,5
2,4
6,9
1,9
2,7
2,1
17,9
2,5
9,2

12
9,4
5,7
0,8
3,5
5,6
1,1
4,6
2,4
2,5
7,5
1,3
3,2
7
1,8
1,5

5,8
20
17,4
2,7
5
4,4
3
1,8
6,9
3,1
6,2
20
1,8
17,6
8,8
1,1

2
13,4
1,3
1,8
11,5
18
2,6
17,8
18,8
3,1
5,5
20
10,2
7,7
20
19,6

15,4
12,4
20
1,9
17,5
19,8
4,2
1,6
17,7
20
4,5
9,7
2,8
0,6
7,6
5,6

5,2
4,4
3,4
6,4
5,5
6,7
12,3
4
9,6
20
1,1
4,2
8,2
1,5
1,3
5,7

0,8
3,1
19,7
9,5
2,1
10
6,1
10,4
15,2
7,3
4,6
13,9
18,7
5,4
3,7
7,1

20
1,4
5
10,3
3,4
5,3
9,7
1,9
13,4
17,8
20
2,4
19,1
18,5
17,9
4,2

13
6,7
13,6
13,4
13,6
4,7
7,5
20
3,1
17,9
19,4
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3,8
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6,9
8,2

20
19,5
15,8
17,8
18,3
11,4
18,1
19,1
2,9
20
17,1
20
11,7
6,8
20
12

16,8
19,4
10,1
9
18,7
13,2
17