Overview

Shielding Effect		
Played on	13 Nov 2019	
Hosted by	JenKrug	
Played with	30 players	
Played	10 of 10	

Overall Performance	
Total correct answers (%)	84,009
Total incorrect answers (%)	16,009
Average score (points)	9183,1

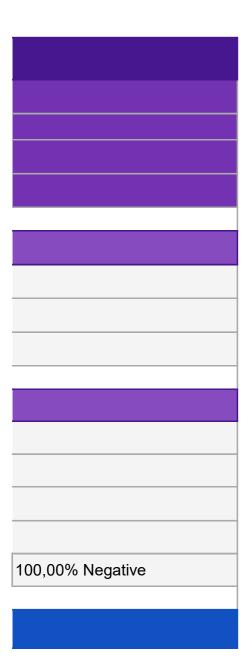
Feedback	
Number of responses	2
How fun was it? (out of 5)	0,00 o
Did you learn something?	0,00%
Do you recommend it?	0,00%
How do you feel?	•

Switch tabs/pages to view other result breakdown

Overview

%			
%			
I3 points			
ut of 5			
Yes	0,0	0% No	
Yes	0,0	0% No	
0,00% Positive	•	0,00% Neutral	•
	·		

Overview



Shielding Effect Final Scores Players Rank 1 maggie 2 ashley 3 N8 Baker 4 sophia 5 Wale (lil duub) 6 (((Jajuan))) 7 Gayatri 8 Will 9 pravleen 10 Reba 11 Sriveena 12 alex demchenko 13 julia 14 duncan 15 zoe 16 conner parker 17 riya 18 Dhanshree 19 Luke 20 Shrey 21 nya 22 ?? ronojoy ??

23	matthew
24	hunter
25	Andrew F
26	Will Paasch
27	Sam S
28	Timothy
29	Arnav(>^<)
30	Mo The Pro

Total Score (points)	Correct Answers	Incorrect Answers
13411	10	0
13406	10	0
12426	10	0
12408	10	0
12147	10	0
12113	10	0
11897	10	0
11878	10	0
11843	10	0
11437	10	0
11302	10	0
10148	9	1
9660	9	1
9443	9	1
9246	9	1
9140	9	1
9042	9	1
8728	8	2
8545	8	2
8056	8	2
8036	8	2
7302	8	2

3	2
7	3
7	3
7	3
6	4
7	3
6	4
(10

Shielding Effect

Jinelanig	
Kahoot! Sur	mmary
Rank	Players
1	maggie
2	ashley
3	N8 Baker
4	sophia
5	Wale (lil duub)
6	(((Jajuan)))
7	Gayatri
8	Will
9	pravleen
10	Reba
11	Sriveena
12	alex demchenko
13	julia
14	duncan
15	zoe

16	conner parker
	common parker
17	riya
18	Dhanshree
19	Luke
20	Shrey
21	nya
22	?? ronojoy ??
23	matthew
24	hunter
25	Andrew F
26	Will Paasch
27	Sam S
28	Timothy
29	Arnav(>^<)
30	Mo The Pro

	Ranoot. Sammary	
Total Score (points)	Q1	
13411	983	
13406	987	
12426	922	
12408	953	
12147	932	
12113	917	
11897	765	
11878	885	
11843	673	
11437	732	
11302	777	
10148	868	
9660	785	
9443	803	
9246	670	

9140	655
9042	682
8728	947
8545	675
8056	683
8036	633
7302	855
7256	0
6889	930
6502	850
6236	537
5891	0
5801	555
5305	877
0	0

What is shielding ?	Q2
Interference from inner core electrons	1087
Interference from inner core electrons	1082
Interference from inner core electrons	990
Interference from inner core electrons	898
Interference from inner core electrons	840
Interference from inner core electrons	853
Interference from inner core electrons	963
Interference from inner core electrons	882
Interference from inner core electrons	1007
Interference from inner core electrons	945
Interference from inner core electrons	790
Interference from inner core electrons	1027
Interference from inner core electrons	998
Interference from inner core electrons	740
Interference from inner core electrons	718

Interference from inner core electrons	648
Interference from inner core electrons	670
Interference from inner core electrons	0
Interference from inner core electrons	0
Interference from inner core electrons	0
Interference from inner core electrons	762
Interference from inner core electrons	808
how the negative charge affects the protons	738
Interference from inner core electrons	932
Interference from inner core electrons	905
Interference from inner core electrons	0
the ratio of electrons to neutrons	0
Interference from inner core electrons	695
Interference from inner core electrons	875
the ratio of electrons to neutrons	0

What happens to the shielding trend moving from left to right across a period on the table	Q3
number of inner orbitals stays the same	1190
number of inner orbitals stays the same	1187
number of inner orbitals stays the same	1025
number of inner orbitals stays the same	948
number of inner orbitals stays the same	917
number of inner orbitals stays the same	900
number of inner orbitals stays the same	927
number of inner orbitals stays the same	992
number of inner orbitals stays the same	1078
number of inner orbitals stays the same	1018
number of inner orbitals stays the same	1027
number of inner orbitals stays the same	0
number of inner orbitals stays the same	0
number of inner orbitals stays the same	0
number of inner orbitals stays the same	0

number of inner orbitals stays the same	0
number of inner orbitals stays the same	0
number of inner orbital increase	0
number of valence orbitals decreases	0
number of valence orbitals decreases	0
number of inner orbitals stays the same	1100
number of inner orbitals stays the same	0
number of inner orbitals stays the same	1007
number of inner orbitals stays the same	1142
number of inner orbitals stays the same	0
number of inner orbital increase	0
number of inner orbital increase	0
number of inner orbitals stays the same	0
number of inner orbitals stays the same	918
	0

What happens to the shielding trend moving down a group?	Q4
shielding increases because the atom has more inner core orbitals	1290
shielding increases because the atom has more inner core orbitals	1300
shielding increases because the atom has more inner core orbitals	1260
shielding increases because the atom has more inner core orbitals	1130
shielding increases because the atom has more inner core orbitals	1105
shielding increases because the atom has more inner core orbitals	1098
shielding increases because the atom has more inner core orbitals	1090
shielding increases because the atom has more inner core orbitals	1088
shielding increases because the atom has more inner core orbitals	1153
shielding increases because the atom has more inner core orbitals	1052
shielding increases because the atom has more inner core orbitals	1012
The number of orbitals increases	730
The number of orbitals increases	722
The number of orbitals increases	825
The number of orbitals increases	795

The number of orbitals increases	623
The number of orbitals increases	780
The number of orbitals increases	797
The number of orbitals increases	647
The number of orbitals increases	660
shielding increases because the atom has more inner core orbitals	1160
The number of orbitals increases	787
shielding increases because the atom has more inner core orbitals	988
shielding increases because the atom has more inner core orbitals	1150
The number of orbitals increases	828
The amount of neutrons decreases	585
The number of orbitals increases	692
The number of orbitals increases	517
shielding increases because the atom has more inner core orbitals	0
	0

How does shielding effect valence electrons?	Q5
inner core electrons block attractive force	1400
inner core electrons block attractive force	1390
inner core electrons block attractive force	1323
inner core electrons block attractive force	1352
inner core electrons block attractive force	1333
inner core electrons block attractive force	1335
inner core electrons block attractive force	1260
inner core electrons block attractive force	1205
inner core electrons block attractive force	1193
inner core electrons block attractive force	1182
inner core electrons block attractive force	1058
inner core electrons block attractive force	948
inner core electrons block attractive force	947
inner core electrons block attractive force	908
inner core electrons block attractive force	865

inner core electrons block attractive force	933
inner core electrons block attractive force	835
inner core electrons block attractive force	910
inner core electrons block attractive force	1062
inner core electrons block attractive force	885
inner core electrons block attractive force	1190
inner core electrons block attractive force	0
inner core electrons block attractive force	0
inner core electrons block attractive force	0
inner core electrons block attractive force	887
inner core electrons block attractive force	0
inner core electrons block attractive force	873
inner core electrons block attractive force	877
	0
	0

How does Coulomb's Law affect the shielding trend	Q6
causes valence electrons to fall of and get lost	1485
causes valence electrons to fall of and get lost	1500
causes valence electrons to fall of and get lost	1433
causes valence electrons to fall of and get lost	1435
causes valence electrons to fall of and get lost	1400
causes valence electrons to fall of and get lost	1408
causes valence electrons to fall of and get lost	1443
causes valence electrons to fall of and get lost	1435
causes valence electrons to fall of and get lost	1423
causes valence electrons to fall of and get lost	1338
causes valence electrons to fall of and get lost	1385
causes valence electrons to fall of and get lost	1083
causes valence electrons to fall of and get lost	1120
causes valence electrons to fall of and get lost	1140
causes valence electrons to fall of and get lost	1073

causes valence electrons to fall of and get lost	1088
causes valence electrons to fall of and get lost	1125
causes valence electrons to fall of and get lost	1105
causes valence electrons to fall of and get lost	1068
causes valence electrons to fall of and get lost	1053
causes valence electrons to fall of and get lost	1415
the protons increase	908
the protons increase	808
the protons increase	910
causes valence electrons to fall of and get lost	1128
causes electrons to disintegrate	945
causes valence electrons to fall of and get lost	983
causes valence electrons to fall of and get lost	1113
	893
	0

Which period has the largest value for the shielding trend?	Q7
period 7 bc they have the most orbitals	1488
period 7 bc they have the most orbitals	1483
period 7 bc they have the most orbitals	1325
period 7 bc they have the most orbitals	1412
period 7 bc they have the most orbitals	1400
period 7 bc they have the most orbitals	1397
period 7 bc they have the most orbitals	1398
period 7 bc they have the most orbitals	1360
period 7 bc they have the most orbitals	1393
period 7 bc they have the most orbitals	1160
period 7 bc they have the most orbitals	1363
period 7 bc they have the most orbitals	1260
period 7 bc they have the most orbitals	1157
period 7 bc they have the most orbitals	1063
period 7 bc they have the most orbitals	1165

period 7 bc they have the most orbitals	1113
period 7 bc they have the most orbitals	1162
period 7 bc they have the most orbitals	962
period 7 bc they have the most orbitals	1185
period 7 bc they have the most orbitals	1108
period 7 bc they have the most orbitals	0
period 7 bc they have the most orbitals	775
period 7 bc they have the most orbitals	640
period 7 bc they have the most orbitals	0
period 7 bc they have the most orbitals	992
period 7 bc they have the most orbitals	882
period 7 bc they have the most orbitals	978
period 7 bc they have the most orbitals	1092
period 7 bc they have the most orbitals	962
	0

What is "effective nuclear charge"?	Q8
net positive charge experienced by an electron	1500
net positive charge experienced by an electron	1492
net positive charge experienced by an electron	1398
net positive charge experienced by an electron	1362
net positive charge experienced by an electron	1343
net positive charge experienced by an electron	1335
net positive charge experienced by an electron	1253
net positive charge experienced by an electron	1340
net positive charge experienced by an electron	1292
net positive charge experienced by an electron	1120
net positive charge experienced by an electron	1110
net positive charge experienced by an electron	1307
net positive charge experienced by an electron	1218
net positive charge experienced by an electron	1162
net positive charge experienced by an electron	1137

net positive charge experienced by an electron	1248
net positive charge experienced by an electron	1175
net positive charge experienced by an electron	1087
net positive charge experienced by an electron	1258
net positive charge experienced by an electron	1130
net negative charge of the electron	0
net positive charge experienced by an electron	902
net positive charge experienced by an electron	793
net negative charge of the electron	783
net positive charge experienced by an electron	0
net positive charge experienced by an electron	920
net positive charge experienced by an electron	917
net positive charge experienced by an electron	0
net positive charge experienced by an electron	780
	0

What is the net positive charge?	Q9
how many availbe protons are acting on valence electrons	1500
how many availbe protons are acting on valence electrons	1500
how many availbe protons are acting on valence electrons	1365
how many availbe protons are acting on valence electrons	1470
how many availbe protons are acting on valence electrons	1452
how many availbe protons are acting on valence electrons	1450
how many availbe protons are acting on valence electrons	1468
how many availbe protons are acting on valence electrons	1393
how many availbe protons are acting on valence electrons	1463
how many availbe protons are acting on valence electrons	1447
how many availbe protons are acting on valence electrons	1435
how many availbe protons are acting on valence electrons	1465
how many availbe protons are acting on valence electrons	1420
how many availbe protons are acting on valence electrons	1462
how many availbe protons are acting on valence electrons	1455

141/	how many availbe protons are acting on valence electrons
1463	how many availbe protons are acting on valence electrons
	how many availbe protons are acting on valence electrons
1777	how many availbe protons are acting on valence electrons
11/1/	how many availbe protons are acting on valence electrons
938	how many electrons and in orbitals
1002	how many availbe protons are acting on valence electrons
1147	how many availbe protons are acting on valence electrons
111147	how many availbe protons are acting on valence electrons
912	the average number of protons
1 14 1	how many availbe protons are acting on valence electrons
1448	how many availbe protons are acting on valence electrons
952	how many periods are in the periodic table
	how many availbe protons are acting on valence electrons
0	

Which group has the smallest value for the shielding trend?	Q10
Period 1, it has the least amount of orbitals	1488
Period 1, it has the least amount of orbitals	1485
Period 1, it has the least amount of orbitals	1385
Period 1, it has the least amount of orbitals	1448
Period 1, it has the least amount of orbitals	1425
Period 1, it has the least amount of orbitals	1420
Period 1, it has the least amount of orbitals	1330
Period 1, it has the least amount of orbitals	1298
Period 1, it has the least amount of orbitals	1168
Period 1, it has the least amount of orbitals	1443
Period 1, it has the least amount of orbitals	1345
Period 1, it has the least amount of orbitals	1460
Period 1, it has the least amount of orbitals	1293
Period 1, it has the least amount of orbitals	1340
Period 1, it has the least amount of orbitals	1368

Period 1, it has the least amount of orbitals	1415
Period 1, it has the least amount of orbitals	1150
Period 1, it has the least amount of orbitals	1463
Period 1, it has the least amount of orbitals	1428
Period 1, it has the least amount of orbitals	1395
Period 1, it has the least amount of orbitals	838
Period 1, it has the least amount of orbitals	1265
Period 1, it has the least amount of orbitals	1140
Period 1, it has the least amount of orbitals	0
Period 1, it has the least amount of orbitals	0
Period 1, it has the least amount of orbitals	1125
Period 1, it has the least amount of orbitals	0
Period 1, it has the least amount of orbitals	0
	0
	0

Shielding effect affects the neutrons and protor	าร
	False

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Fals
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Shielding

1 Quiz

Correct answers

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Players

(((Jajuan)))

Andrew F

Arnav(>^<)

Dhanshree

Gayatri

Luke

Mo The Pro

N8 Baker

Reba

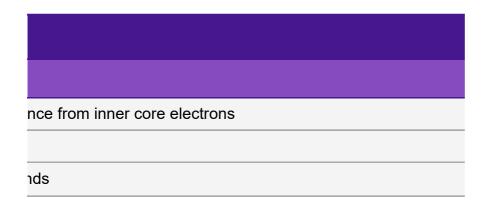
Sam S

Shrey
Sriveena
Timothy
Wale (lil duub)
Will
Will Paasch
alex demchenk
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect	
What is shielding ?	
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1 Quiz

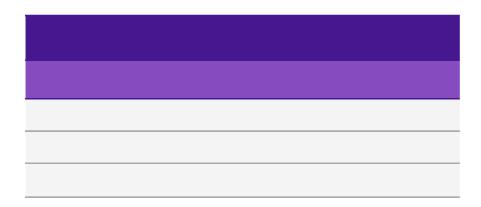
	√ □
	√ □
)	√ □
	Х
	√ □
	√ Ω
	√ Ω
	√ □
	√ □
	√ □



Interference from inner core electrons	•
√ □	
	27
12,	,16

	Score (p
Interference from inner core electrons	917
Interference from inner core electrons	850
Interference from inner core electrons	877
Interference from inner core electrons	947
Interference from inner core electrons	765
Interference from inner core electrons	675
the ratio of electrons to neutrons	0
Interference from inner core electrons	922
Interference from inner core electrons	732
the ratio of electrons to neutrons	0

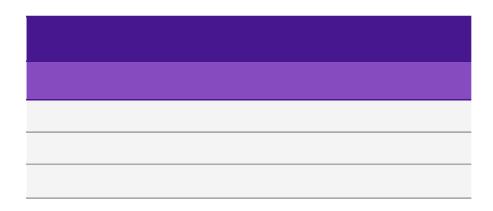
Interference from inner core electrons	683
Interference from inner core electrons	777
Interference from inner core electrons	555
Interference from inner core electrons	932
Interference from inner core electrons	885
Interference from inner core electrons	537
Interference from inner core electrons	868
Interference from inner core electrons	987
Interference from inner core electrons	655
Interference from inner core electrons	803
Interference from inner core electrons	930
Interference from inner core electrons	785
Interference from inner core electrons	983
how the negative charge affects the protons	0
Interference from inner core electrons	633
Interference from inner core electrons	673
Interference from inner core electrons	682
Interference from inner core electrons	953
Interference from inner core electrons	670
Interference from inner core electrons	855



when electrons dont fit on the orbital		•
X		
	0	
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oints)	Current
	917
	850
	877
	947
	765
	675
	0
	922
	732
	0

683
777
555
932
885
537
868
987
655
803
930
785
983
0
633
673
682
953
670
855



how the negative charge affects the protons		•
X		
	1	
5,6	60	

Total Score (points)	Answer ti
	5
	9
	7,4
	3,2
	14,1
	19,5
	27,5
	4,7
	16,1
	26,8

19
13,4
26,7
4,1
6,9
27,8
7,9
0,8
20,7
11,8
4,2
12,9
1
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22
19,6
19,1
2,8
19,8
8,7
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the ratio of electrons to neutrons	
X	
	2
	27,15
me (seconds)	

Shielding

2 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

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duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect	
What happens to the shielding trend m	noving from left to right acr
;	number
(%)	80,00%
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ken to answer (seconds)	
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	Answer
	√ □
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oss a period on the table

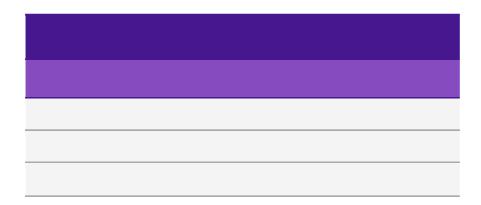
of inner orbitals stays the same

nds

number of inner orbital increase	*
X	
3	
10,57	

	Score (p
number of inner orbitals stays the same	853
number of inner orbitals stays the same	905
number of inner orbitals stays the same	875
number of inner orbital increase	0
number of inner orbitals stays the same	963
number of valence orbitals decreases	0
	0
number of inner orbitals stays the same	990
number of inner orbitals stays the same	945
number of inner orbital increase	0

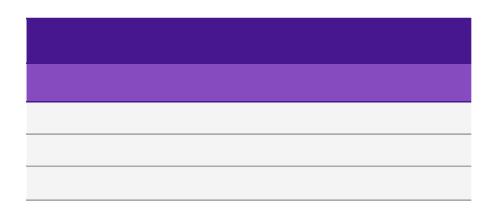
number of valence orbitals decreases	0
number of inner orbitals stays the same	790
number of inner orbitals stays the same	695
number of inner orbitals stays the same	840
number of inner orbitals stays the same	882
number of inner orbital increase	0
number of inner orbitals stays the same	1027
number of inner orbitals stays the same	1082
number of inner orbitals stays the same	648
number of inner orbitals stays the same	740
number of inner orbitals stays the same	932
number of inner orbitals stays the same	998
number of inner orbitals stays the same	1087
number of inner orbitals stays the same	738
number of inner orbitals stays the same	762
number of inner orbitals stays the same	1007
number of inner orbitals stays the same	670
number of inner orbitals stays the same	898
number of inner orbitals stays the same	718
number of inner orbitals stays the same	808



number of valence orbitals increases	•
X	
0	
0,00	

oints)	Current
	1770
	1755
	1752
	947
	1728
	675
	0
	1912
	1677
	0

683
1567
1250
1772
1767
537
1895
2069
1303
1543
1862
1783
2070
738
1395
1680
1352
1851
1388
1663



number of inner orbitals stays the same		-
√ □		
	24	
	13,62	

Total Score (points)	Answer ti
	14,8
	11,7
	13,5
	5,5
	8,2
	28,6
	30
	6,6
	9,3
	13,4

22,1
18,6
24,3
15,6
13,1
12,8
4,4
1,1
27,1
21,6
10,1
6,1
0,8
15,7
20,3
5,6
25,8
12,1
22,9
17,5

number of valence orbitals decreases	
X	
	2
	25,35
ima (accanda)	
me (seconds)	

Shielding

3 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

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N8 Baker

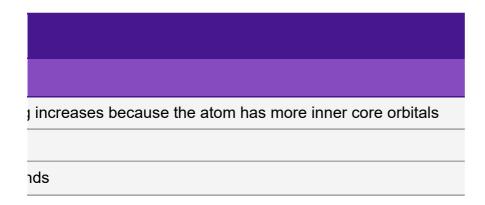
Reba

Sam S

Shrey
Sriveena
Timothy
Wale (lil duub)
Will
Will Paasch
alex demchenko
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect What happens to the shielding trend moving down a group? shielding (%) 50,00% 30 secor on nmary ct? ers received ken to answer (seconds) ails Answer **√**□ X **√**□ X **√**□ X X **√**□ **√**□ X

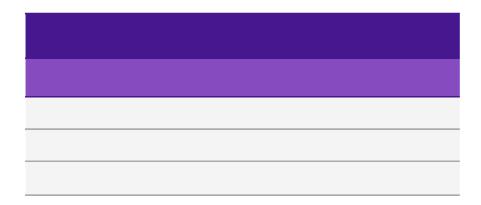
	Х
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	Х
	Х



The amount of neutrons decreases	•
X	
1	
26,40	

	Score (p
shielding increases because the atom has more inner core orbitals	900
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	918
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	927
The number of orbitals increases	0
	0
shielding increases because the atom has more inner core orbitals	1025
shielding increases because the atom has more inner core orbitals	1018
The number of orbitals increases	0

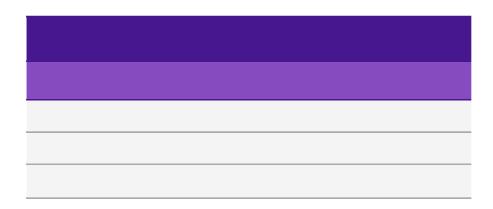
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	1027
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	917
shielding increases because the atom has more inner core orbitals	992
The amount of neutrons decreases	0
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	1187
The number of orbitals increases	0
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	1142
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	1190
shielding increases because the atom has more inner core orbitals	1007
shielding increases because the atom has more inner core orbitals	1100
shielding increases because the atom has more inner core orbitals	1078
The number of orbitals increases	0
shielding increases because the atom has more inner core orbitals	948
The number of orbitals increases	0
The number of orbitals increases	0



The mass increases	•
X	
0	
0,00	

oints)	Current
	2670
	1755
	2670
	947
	2655
	675
	0
	2937
	2695
	0

683
2594
1250
2689
2759
537
1895
3256
1303
1543
3004
1783
3260
1745
2495
2758
1352
2799
1388
1663



The number of orbitals increases	
X	
13	
12,38	

Total Score (points)	Answer ti
	18
	8,6
	16,9
	5,7
	16,4
	22
	30
	10,5
	10,9
	28,8

8,4
10,4
26,9
17
12,5
26,4
6,2
0,8
23,6
5,5
3,5
7
0,6
5,6
6
7,3
5,1
15,1
6,1
7,1

shielding increases because the atom has more inner core orbitals	
√I	
15	5
	\dashv
10,10	, -
me (seconds)	
	-
	_
	_
	_
	_
	_
	1

Shielding

4 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

Dhanshree

Gayatri

Luke

Mo The Pro

N8 Baker

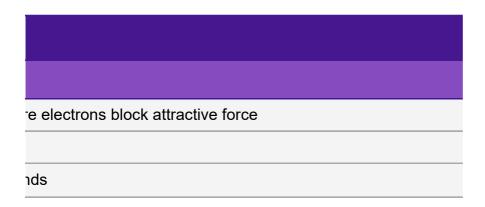
Reba

Sam S

Shrey
Sriveena
Timothy
Wale (lil duub)
Will
Will Paasch
alex demchenk
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

How does shielding effect valence electrons? inner (%) 93,33	
(%)	
	r cor
	3%
on 30 se	ecor
nmary	
	•
xt?	
/ers received	
ken to answer (seconds)	
ails	
Answ	Wor
✓□	
X	
✓□	
√ □	
√ □	
X	
<u>√</u> □	
√ □	

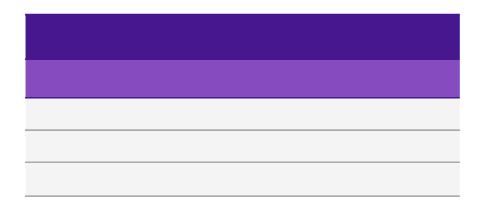
	√ □
	√ □
)	√ □
	√ □



smaller electrons fall off the orbital	•
X	
0	
0,00	

inner core electrons block attractive force 1	Score (p 1098 328
inner core electrons block attractive force 8	
	328
0	
)
inner core electrons block attractive force 7	' 97
inner core electrons block attractive force	090
inner core electrons block attractive force 6	647
0)
inner core electrons block attractive force 1	260
inner core electrons block attractive force 1	052
inner core electrons block attractive force 6	692

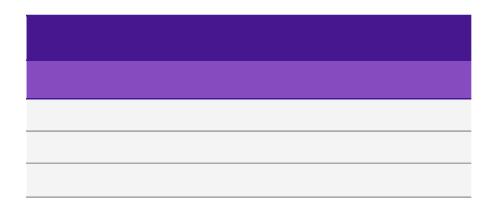
inner core electrons block attractive force	660
inner core electrons block attractive force	1012
inner core electrons block attractive force	517
inner core electrons block attractive force	1105
inner core electrons block attractive force	1088
inner core electrons block attractive force	585
inner core electrons block attractive force	730
inner core electrons block attractive force	1300
inner core electrons block attractive force	623
inner core electrons block attractive force	825
inner core electrons block attractive force	1150
inner core electrons block attractive force	722
inner core electrons block attractive force	1290
inner core electrons block attractive force	988
inner core electrons block attractive force	1160
inner core electrons block attractive force	1153
inner core electrons block attractive force	780
inner core electrons block attractive force	1130
inner core electrons block attractive force	795
inner core electrons block attractive force	787
	



inner core electrons block attractive force		•
√ □		
	28	
	13,38	

oints)	Current
	3768
	2583
	2670
	1744
	3745
	1322
	0
	4197
	3747
	692

1343
3606
1767
3794
3847
1122
2625
4556
1926
2368
4154
2505
4550
2733
3655
3911
2132
3929
2183
2450



protons are not related to valence electrons	-
X	
	ו
0,00	ס

Total Score (points)	Answer ti
	12,1
	10,3
	30
	12,2
	12,6
	21,2
	30
	2,4
	14,9
	18,5

20,4
17,3
29
11,7
12,7
24,9
16,2
0,4
22,6
10,5
9
16,7
0,6
12,7
8,4
8,8
13,2
10,2
12,3
12,8

protons are weaker than valance electrons	
X	
	0
	0,00
me (seconds)	

Shielding

5 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

Dhanshree

Gayatri

Luke

Mo The Pro

N8 Baker

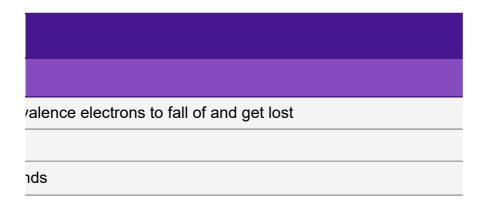
Reba

Sam S

Shrey
Sriveena
Timothy
Wale (lil duub)
Will
Will Paasch
alex demchenk
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect	
How does Coulomb's Law affect	t the shielding trend
;	causes
(%)	80,00%
on	30 seco
nmary	
xt?	
ers received	
ken to answer (seconds)	
ails	
	Answer
	√ □
	√ □
	X
	√ □
	√ □
	√□
	X
	√ □
	√ □
	√ 1

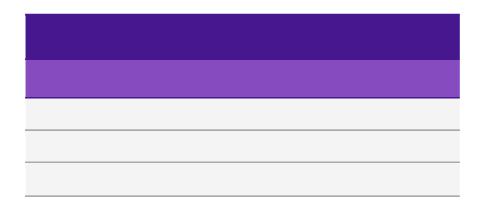
	√ □
	√ □
	Х
)	√ □
	Х
	√ □
	√ □
	Х
	√ □
	Х



the element changes shape and color	•
X	
()
0,00)

	Score (p
causes valence electrons to fall of and get lost	1335
causes valence electrons to fall of and get lost	887
	0
causes valence electrons to fall of and get lost	910
causes valence electrons to fall of and get lost	1260
causes valence electrons to fall of and get lost	1062
	0
causes valence electrons to fall of and get lost	1323
causes valence electrons to fall of and get lost	1182
causes valence electrons to fall of and get lost	873

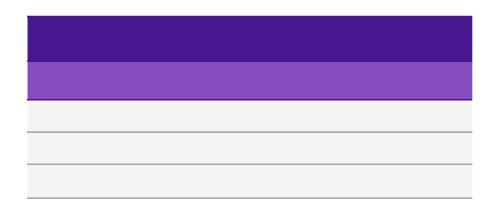
causes valence electrons to fall of and get lost	885
causes valence electrons to fall of and get lost	1058
causes valence electrons to fall of and get lost	877
causes valence electrons to fall of and get lost	1333
causes valence electrons to fall of and get lost	1205
causes electrons to disintegrate	0
causes valence electrons to fall of and get lost	948
causes valence electrons to fall of and get lost	1390
causes valence electrons to fall of and get lost	933
causes valence electrons to fall of and get lost	908
the protons increase	0
causes valence electrons to fall of and get lost	947
causes valence electrons to fall of and get lost	1400
the protons increase	0
causes valence electrons to fall of and get lost	1190
causes valence electrons to fall of and get lost	1193
causes valence electrons to fall of and get lost	835
causes valence electrons to fall of and get lost	1352
causes valence electrons to fall of and get lost	865
the protons increase	0



causes valence electrons to fall of and get lost	•
√ □	
24	
9,64	

oints)	Current
	5103
	3470
	2670
	2654
	5005
	2384
	0
	5520
	4929
	1565

2228
4664
2644
5127
5052
1122
3573
5946
2859
3276
4154
3452
5950
2733
4845
5104
2967
5281
3048
2450



the protons increase	_
X	
3	
15,90	

Total Score (points)	Answer ti
	3,9
	12,8
	30
	11,4
	8,4
	2,3
	30
	4,6
	13,1
	13,6

12,9
20,5
13,4
4
11,7
26,5
9,1
0,6
10
11,5
8,8
9,2
0,4
9,9
12,6
12,4
15,9
2,9
14,1
29

causes electrons to disintegrate	
Х	1
	26,50
me (seconds)	

Shielding

6 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

Dhanshree

Gayatri

Luke

Mo The Pro

N8 Baker

Reba

Sam S

Shrey
Sriveena
Timothy
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ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect Which period has the largest value for the shielding trend? period 7 (%) 96,67% 20 secor on nmary ct? ers received ken to answer (seconds) ails Answer **√**□ **√**□ **√**□ **√**□ **√**□ **√**□ X **√**□ **√**□ **√**□

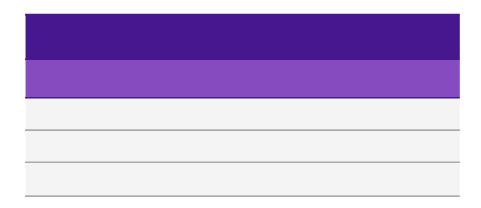
	√ □
	√ □
)	√ □
	√ □

bc they have the most orbitals
nds

period 7 bc they have the most orbitals	•
√ □	
29	
3,83	3

	Score (p
period 7 bc they have the most orbitals	1408
period 7 bc they have the most orbitals	1128
period 7 bc they have the most orbitals	893
period 7 bc they have the most orbitals	1105
period 7 bc they have the most orbitals	1443
period 7 bc they have the most orbitals	1068
	0
period 7 bc they have the most orbitals	1433
period 7 bc they have the most orbitals	1338
period 7 bc they have the most orbitals	983

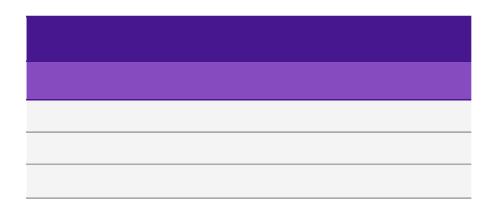
period 7 bc they have the most orbitals	1053
period 7 bc they have the most orbitals	1385
period 7 bc they have the most orbitals	1113
period 7 bc they have the most orbitals	1400
period 7 bc they have the most orbitals	1435
period 7 bc they have the most orbitals	945
period 7 bc they have the most orbitals	1083
period 7 bc they have the most orbitals	1500
period 7 bc they have the most orbitals	1088
period 7 bc they have the most orbitals	1140
period 7 bc they have the most orbitals	910
period 7 bc they have the most orbitals	1120
period 7 bc they have the most orbitals	1485
period 7 bc they have the most orbitals	808
period 7 bc they have the most orbitals	1415
period 7 bc they have the most orbitals	1423
period 7 bc they have the most orbitals	1125
period 7 bc they have the most orbitals	1435
period 7 bc they have the most orbitals	1073
period 7 bc they have the most orbitals	908



period 3 bc they have the most protons		•
X		
	0	
	0,00	

·	Current 6511
6	5511
4	1598
З	3563
З	3759
6	6448
З	3452
С)
6	6953
6	6267
2	2548

3281
6049
3757
6527
6487
2067
4656
7446
3947
4416
5064
4572
7435
3541
6260
6527
4092
6716
4121
3358



period 5 bc they have the most electrons		•
X		
	0	
	0,00	

Total Score (points)	Answer ti
	3,7
	2,9
	4,3
	3,8
	2,3
	5,3
	20
	2,7
	6,5
	8,7

5,9
4,6
3,5
4
2,6
2,2
4,7
0,4
4,5
2,4
3,6
3,2
0,6
7,7
3,4
3,1
3
2,6
5,1
3,7

period 6 bc the have the most neutrons	
X	
	0
	0,00
me (seconds)	

Shielding

7 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

Dhanshree

Gayatri

Luke

Mo The Pro

N8 Baker

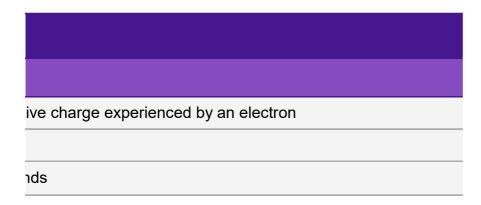
Reba

Sam S

Shrey
Sriveena
Timothy
Wale (lil duub)
Will
Will Paasch
alex demchenk
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect	
What is "effective nuclear charge"?	
;	net posit
(%)	90,00%
n	30 secor
nmary	
	A
ot?	
/ers received	
ken to answer (seconds)	
ails	
	Answer
	√ □
	Х
	√ □
	√ □
	√ □

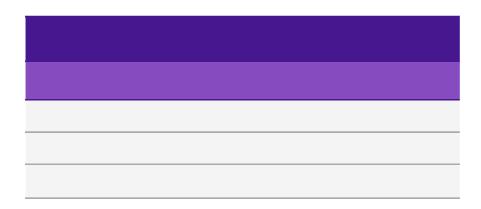
	√ □
	√ □
	√ 0
	√ □
	√ □
	√ □
)	√ □
	Х
	X ✓
	√ □
	Δ.
	\D \D
	χ .α .α
	A A A A
	л л л х



the average amount of electrons	
X	
0	
0,00	

	Score (p
net positive charge experienced by an electron	1397
net positive charge experienced by an electron	992
net positive charge experienced by an electron	962
net positive charge experienced by an electron	962
net positive charge experienced by an electron	1398
net positive charge experienced by an electron	1185
	0
net positive charge experienced by an electron	1325
net positive charge experienced by an electron	1160
net positive charge experienced by an electron	978

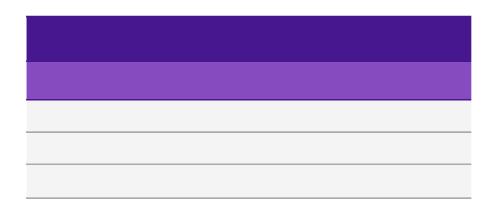
net positive charge experienced by an electron	1108
net positive charge experienced by an electron	1363
net positive charge experienced by an electron	1092
net positive charge experienced by an electron	1400
net positive charge experienced by an electron	1360
net positive charge experienced by an electron	882
net positive charge experienced by an electron	1260
net positive charge experienced by an electron	1483
net positive charge experienced by an electron	1113
net positive charge experienced by an electron	1063
net negative charge of the electron	0
net positive charge experienced by an electron	1157
net positive charge experienced by an electron	1488
net positive charge experienced by an electron	640
net negative charge of the electron	0
net positive charge experienced by an electron	1393
	1
net positive charge experienced by an electron	1162
net positive charge experienced by an electron net positive charge experienced by an electron	1162
, , ,	
net positive charge experienced by an electron	1412



net negative charge of the electron		•
X		
	2	
	13,95	

oints) Cu 790 550 453 473 784 463 0 823	
559 452 472 784 463 0 823	ırrent
453 473 784 463 0 823	80
473 784 463 0 82	90
784 463 0 821	25
463 0 823	21
0 82	46
82	37
742	78
	27
352	26

4389
7412
4849
7927
7847
2949
5916
8929
5060
5479
5064
5729
8923
4181
6260
7920
5254
8128
5286
4133



net positive charge experienced by an electron	
√ □	
27	
10,72	

Total Score (points)	Answer ti
	6,2
	18,5
	8,3
	20,3
	6,1
	6,9
	30
	10,5
	20,4
	19,3

11,5
8,2
12,5
6
8,4
13,1
2,4
1
11,2
14,2
15,5
8,6
0,7
27,6
12,4
6,4
8,3
5,3
8,1
19,5
1.0,0

average number of protons	
X	
^	0
	0
	0,00
me (seconds)	

Shielding

8 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

Dhanshree

Gayatri

Luke

Mo The Pro

N8 Baker

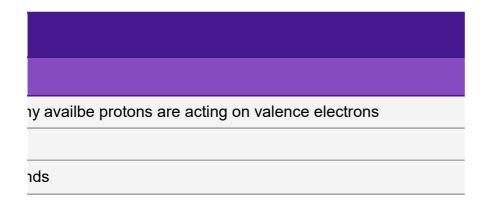
Reba

Sam S

Shrey
Sriveena
Timothy
Wale (lil duub)
Will
Will Paasch
alex demchenko
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect	
What is the net positive charge?	
;	how mai
(%)	86,67%
n	30 secon
nmary	
	<u> </u>
xt?	
ers received	
ken to answer (seconds)	
ails	
	Answer
	<u>√</u> 1
	X
	√ □
	√□
	√□
	√ □
	X
	√ □
	√ □
	√ 1

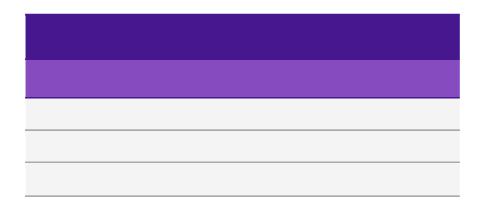
	√ □
	√ □
	Х
	√ □
	√ □
	√ □
)	√ □
	Х
	√ □



how many availbe protons are acting on valence electrons	•
√∆	
26	
13,72	

	Score (p
how many availbe protons are acting on valence electrons	1335
the average number of protons	0
how many availbe protons are acting on valence electrons	780
how many availbe protons are acting on valence electrons	1087
how many availbe protons are acting on valence electrons	1253
how many availbe protons are acting on valence electrons	1258
	0
how many availbe protons are acting on valence electrons	1398
how many availbe protons are acting on valence electrons	1120
how many availbe protons are acting on valence electrons	917

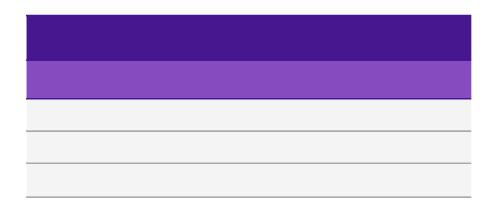
how many availbe protons are acting on valence electrons	1130
how many availbe protons are acting on valence electrons	1110
how many periods are in the periodic table	0
how many availbe protons are acting on valence electrons	1343
how many availbe protons are acting on valence electrons	1340
how many availbe protons are acting on valence electrons	920
how many availbe protons are acting on valence electrons	1307
how many availbe protons are acting on valence electrons	1492
how many availbe protons are acting on valence electrons	1248
how many availbe protons are acting on valence electrons	1162
how many availbe protons are acting on valence electrons	783
how many availbe protons are acting on valence electrons	1218
how many availbe protons are acting on valence electrons	1500
how many availbe protons are acting on valence electrons	793
how many electrons and in orbitals	0
how many availbe protons are acting on valence electrons	1292
how many availbe protons are acting on valence electrons	1175
how many availbe protons are acting on valence electrons	1362
how many availbe protons are acting on valence electrons	1137
how many availbe protons are acting on valence electrons	902
2,22,310	



how many periods are in the periodic table		•
X		
	1	
29	9,10	

oints)	Current
	9243
	5590
	5305
	5808
	9099
	5895
	0
	9676
	8547
	4443

5519
8522
4849
9270
9187
3869
7223
10421
6308
6641
5847
6947
10423
4974
6260
9212
6429
9490
6423
5035



how many electrons and in orbitals	
X	
1	
26,20	

Total Score (points)	Answer ti
	9,9
	10
	25,2
	18,8
	14,8
	8,5
	30
	6,1
	22,8
	29

16,2
23,4
29,1
9,4
9,6
16,8
5,6
0,5
9,1
14,3
13
10,9
0,4
24,4
26,2
12,5
13,5
8,3
15,8
17,9

the average number of protons	
X	
	1
10),00
ime (seconds)	

Shielding

9 Quiz

Correct answers

Players correct (

Question duration

Answer Sun

Answer options

Is answer correct

Number of answ

Average time tal

Answer Deta

Players

(((Jajuan)))

Andrew F

Arnav(>^<)

Dhanshree

Gayatri

Luke

Mo The Pro

N8 Baker

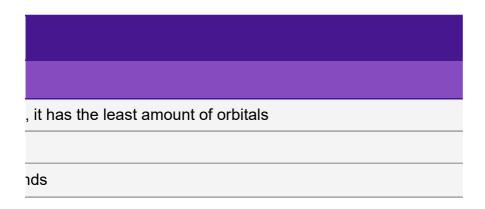
Reba

Sam S

Shrey
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Will Paasch
alex demchenk
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Effect	
Which group has the smallest value for the	shielding trend?
;	Period 1
(%)	93,33%
n	30 secor
	·
nmary	
	A
pt?	
vers received	
ken to answer (seconds)	
ails	
	Answer
	√ □
	√ □
	Х
	√ □
	√ □
	√ □
	Х
	√ □
	√ □
	√□

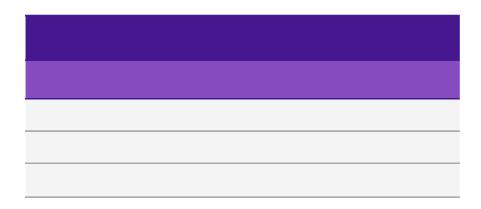
	√ □
	√ □
)	√ □
	√ □



Period 4, because they have the most protons	*
X	
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	Score (p
Period 1, it has the least amount of orbitals	1450
Period 1, it has the least amount of orbitals	912
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Period 1, it has the least amount of orbitals	1457
Period 1, it has the least amount of orbitals	1468
Period 1, it has the least amount of orbitals	1222
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Period 1, it has the least amount of orbitals	1365
Period 1, it has the least amount of orbitals	1447
Period 1, it has the least amount of orbitals	1448

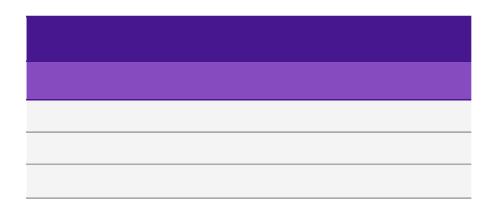
Period 1, it has the least amount of orbitals	1142
Period 1, it has the least amount of orbitals	1435
Period 1, it has the least amount of orbitals	952
Period 1, it has the least amount of orbitals	1452
Period 1, it has the least amount of orbitals	1393
Period 1, it has the least amount of orbitals	1242
Period 1, it has the least amount of orbitals	1465
Period 1, it has the least amount of orbitals	1500
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Period 1, it has the least amount of orbitals	1420
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Period 1, it has the least amount of orbitals	1142
Period 1, it has the least amount of orbitals	938
Period 1, it has the least amount of orbitals	1463
Period 1, it has the least amount of orbitals	1463
Period 1, it has the least amount of orbitals	1470
Period 1, it has the least amount of orbitals	1455
Period 1, it has the least amount of orbitals	1002



Period 1, it has the least amount of orbitals		•
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10580
5111
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7725
8103
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8367
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10675
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Period 6, has the most mass	•
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21,5
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2,9
2,9
6,4
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17,9

Period 3, it has the most mass	
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Shielding

10 Quiz

Correct answers

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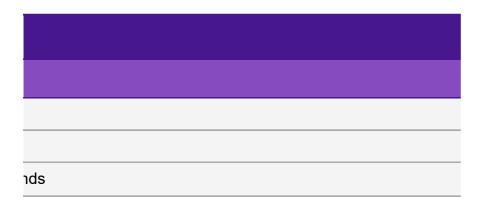
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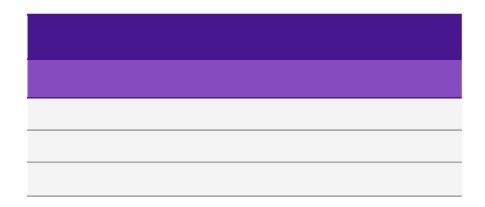
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10 Quiz

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5801
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	Which group has the smallest value for the shielding trend?
Shielding effect affects the neutrons and protons	Which group has the smallest value for the shielding trend?
	Shielding effect affects the neutrons and protons

Shielding effect affects the neutrons and protons
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Shielding effect affects the neutrons and protons

Answer 1	Answer 2
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
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Interference from inner core electrons	when electrons dont fit on the orbital

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Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
Interference from inner core electrons	when electrons dont fit on the orbital
number of inner orbital increase	number of valence orbitals increases

number of inner orbital increase	number of valence orbitals increases
number of inner orbital increase	number of valence orbitals increases
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number of inner orbital increase	number of valence orbitals increases
number of inner orbital increase	number of valence orbitals increases
number of inner orbital increase	number of valence orbitals increases
The amount of neutrons decreases	The mass increases
The amount of neutrons decreases	The mass increases
The amount of neutrons decreases	The mass increases

The amount of neutrons decreases	The mass increases
The amount of neutrons decreases	The mass increases
The amount of neutrons decreases	The mass increases
The amount of neutrons decreases	The mass increases
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The amount of neutrons decreases	The mass increases
The amount of neutrons decreases	The mass increases
smaller electrons fall off the orbital	inner core electrons block attractive force
smaller electrons fall off the orbital	inner core electrons block attractive force
smaller electrons fall off the orbital	inner core electrons block attractive force
smaller electrons fall off the orbital	inner core electrons block attractive force
smaller electrons fall off the orbital	inner core electrons block attractive force

smaller electrons fall off the orbital	inner core electrons block attractive force
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smaller electrons fall off the orbital	inner core electrons block attractive force
smaller electrons fall off the orbital	inner core electrons block attractive force
smaller electrons fall off the orbital	inner core electrons block attractive force
the element changes shape and color	causes valence electrons to fall of and get lost
the element changes shape and color	causes valence electrons to fall of and get lost
the element changes shape and color	causes valence electrons to fall of and get lost
the element changes shape and color	causes valence electrons to fall of and get lost
the element changes shape and color	causes valence electrons to fall of and get lost
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the element changes shape and color	causes valence electrons to fall of and get lost
the element changes shape and color	causes valence electrons to fall of and get lost

causes valence electrons to fall of and get lost
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causes valence electrons to fall of and get lost
period 3 bc they have the most protons
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period 3 bc they have the most protons

period 7 bc they have the most orbitals	period 3 bc they have the most protons
period 7 bc they have the most orbitals	period 3 bc they have the most protons
period 7 bc they have the most orbitals	period 3 bc they have the most protons
period 7 bc they have the most orbitals	period 3 bc they have the most protons
period 7 bc they have the most orbitals	period 3 bc they have the most protons
the average amount of electrons	net negative charge of the electron
the average amount of electrons	net negative charge of the electron
the average amount of electrons	net negative charge of the electron
the average amount of electrons	net negative charge of the electron
the average amount of electrons	net negative charge of the electron
the average amount of electrons	net negative charge of the electron
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the average amount of electrons	net negative charge of the electron

the average amount of electrons	net negative charge of the electron
the average amount of electrons	net negative charge of the electron
the average amount of electrons	net negative charge of the electron
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
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how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table

how many availbe protons are acting on valence electrons	how many periods are in the periodic table
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how many availbe protons are acting on valence electrons	how many periods are in the periodic table
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how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table
how many availbe protons are acting on valence electrons	how many periods are in the periodic table

how many availbe protons are acting on valence electrons	how many periods are in the periodic table
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals

Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
Period 4, because they have the most protons	Period 1, it has the least amount of orbitals
False	True

False	True
False	True
False	True
False	True
False	True
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False	True
False	True
False	True
False	True
False	True
False	True
False	True
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False	True

False	True
False	True
False	True
False	True
False	True
False	True
False	True
False	True
False	True
False	True
False	True
False	True
False	True

Answer 3	Answer 4
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons

how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
how the negative charge affects the protons	the ratio of electrons to neutrons
number of inner orbitals stays the same	number of valence orbitals decreases

number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
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number of inner orbitals stays the same	number of valence orbitals decreases
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number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases

number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
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number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
number of inner orbitals stays the same	number of valence orbitals decreases
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals

The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
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The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals

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The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
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The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
The number of orbitals increases	shielding increases because the atom has more inner core orbitals
protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons

protons are not related to valence electrons	protons are weaker than valance electrons
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protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons
protons are not related to valence electrons	protons are weaker than valance electrons
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate

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the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
the protons increase	causes electrons to disintegrate
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons

period 5 bc they have the most electrons	period 6 bc the have the most neutrons
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period 5 bc they have the most electrons	period 6 bc the have the most neutrons
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period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
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period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons

period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
period 5 bc they have the most electrons	period 6 bc the have the most neutrons
net positive charge experienced by an electron	average number of protons
net positive charge experienced by an electron	average number of protons
net positive charge experienced by an electron	average number of protons
net positive charge experienced by an electron	average number of protons
net positive charge experienced by an electron	average number of protons
net positive charge experienced by an electron	average number of protons
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net positive charge experienced by an electron	average number of protons

average number of protons
average number of protons
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the average number of protons

how many electrons and in orbitals	the average number of protons
how many electrons and in orbitals	the average number of protons
how many electrons and in orbitals	the average number of protons
how many electrons and in orbitals	the average number of protons
how many electrons and in orbitals	the average number of protons
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how many electrons and in orbitals	the average number of protons
how many electrons and in orbitals	the average number of protons
how many electrons and in orbitals	the average number of protons
how many electrons and in orbitals	the average number of protons

number of protons as the most mass
as the most mass
as the most mass

Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass
Period 6, has the most mass	Period 3, it has the most mass

Correct Answers	Time Allotted to Answer (seconds)
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
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Interference from inner core electrons	30

Interference from inner core electrons	30
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Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
Interference from inner core electrons	30
number of inner orbitals stays the same	30

number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
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number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
number of inner orbitals stays the same	30
shielding increases because the atom has more inner core orbitals	30
shielding increases because the atom has more inner core orbitals	30
shielding increases because the atom has more inner core orbitals	30

shielding increases because the atom has more inner core orbitals	30
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shielding increases because the atom has more inner core orbitals	30
shielding increases because the atom has more inner core orbitals	30
inner core electrons block attractive force	30
inner core electrons block attractive force	30
inner core electrons block attractive force	30
inner core electrons block attractive force	30
inner core electrons block attractive force	30

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inner core electrons block attractive force	30
inner core electrons block attractive force	30
inner core electrons block attractive force	30
causes valence electrons to fall of and get lost	30
causes valence electrons to fall of and get lost	30
causes valence electrons to fall of and get lost	30
causes valence electrons to fall of and get lost	30
causes valence electrons to fall of and get lost	30
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causes valence electrons to fall of and get lost	30
period 7 bc they have the most orbitals	20
period 7 bc they have the most orbitals	20
period 7 bc they have the most orbitals	20
period 7 bc they have the most orbitals	20
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period 7 bc they have the most orbitals	20
net positive charge experienced by an electron	30
net positive charge experienced by an electron	30
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net positive charge experienced by an electron	30
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net positive charge experienced by an electron	30
net positive charge experienced by an electron	30
net positive charge experienced by an electron	30
how many availbe protons are acting on valence electrons	30
how many availbe protons are acting on valence electrons	30
how many availbe protons are acting on valence electrons	30
how many availbe protons are acting on valence electrons	30
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how many availbe protons are acting on valence electrons	30
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how many availbe protons are acting on valence electrons	30
Period 1, it has the least amount of orbitals	30
Period 1, it has the least amount of orbitals	30
Period 1, it has the least amount of orbitals	30
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Period 1, it has the least amount of orbitals	30
False	20

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False	20
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False	20
False	20
False	20
False	20
False	20
False	20
False	20

Players
((((Jajuan)))
Andrew F
Arnav(>^<)
Dhanshree
Gayatri
Luke
Mo The Pro
N8 Baker
Reba
Sam S
Shrey
Sriveena
Timothy
Wale (lil duub)
Will

Will Paasch
alex demchenko
ashley
conner parker
duncan
hunter
julia
maggie
matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??
((((Jajuan)))

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matthew
nya
pravleen
riya
sophia
zoe
?? ronojoy ??

Answer	Correct / Incorrect	Correct
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
the ratio of electrons to neutrons	Incorrect	0
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
the ratio of electrons to neutrons	Incorrect	0
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1

Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
how the negative charge affects the protons	Incorrect	0
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
Interference from inner core electrons	Correct	1
number of inner orbitals stays the same	Correct	1

number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbital increase	Incorrect	0
number of inner orbitals stays the same	Correct	1
number of valence orbitals decreases	Incorrect	0
	Incorrect	0
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbital increase	Incorrect	0
number of valence orbitals decreases	Incorrect	0
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbital increase	Incorrect	0
number of inner orbitals stays the same	Correct	1

number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
number of inner orbitals stays the same	Correct	1
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1

The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0
	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0
The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
shielding increases because the atom has more inner core orbitals	Correct	1
The amount of neutrons decreases	Incorrect	0
The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0

The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
shielding increases because the atom has more inner core orbitals	Correct	1
shielding increases because the atom has more inner core orbitals	Correct	1
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0
shielding increases because the atom has more inner core orbitals	Correct	1
The number of orbitals increases	Incorrect	0
The number of orbitals increases	Incorrect	0
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
	Incorrect	0
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
	•	

inner core electrons block attractive force	Correct	1
	Incorrect	0
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
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inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
inner core electrons block attractive force	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
	Incorrect	0
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
	Incorrect	0

causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes electrons to disintegrate	Incorrect	0
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
the protons increase	Incorrect	0
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1

the protons increase	Incorrect	0
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
causes valence electrons to fall of and get lost	Correct	1
the protons increase	Incorrect	0
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
	Incorrect	0
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1

period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1

period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
period 7 bc they have the most orbitals	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
	Incorrect	0
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1

net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net negative charge of the electron	Incorrect	0
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net negative charge of the electron	Incorrect	0
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1

net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
net positive charge experienced by an electron	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
the average number of protons	Incorrect	0
how many availbe protons are acting on valence electrons	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
	Incorrect	0
how many availbe protons are acting on valence electrons	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
how many availbe protons are acting on valence electrons	Correct	1
how many periods are in the periodic table	Incorrect	0

Correct	1
Correct	1
Correct	1
Correct	1
Correct	1
Correct	1
Correct	1
Correct	1
Correct	1
Correct	1
Correct	1
Incorrect	0
Correct	1
	Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct Correct

how many availbe protons are acting on valence electrons	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
	Incorrect	0
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
	Incorrect	0
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1

Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
Period 1, it has the least amount of orbitals	Correct	1
False	Correct	1

True	Incorrect	0
	Incorrect	0
False	Correct	1
False	Correct	1
False	Correct	1
	Incorrect	0
False	Correct	1
False	Correct	1
True	Incorrect	0
False	Correct	1
False	Correct	1
True	Incorrect	0
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1

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False	Correct	1
False	Correct	1
False	Correct	1
True	Incorrect	0
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1
False	Correct	1

Incorrect	Score (points)	Score without Answer Streak Bonus (points)
0	917	917
0	850	850
0	877	877
0	947	947
0	765	765
0	675	675
1	0	0
0	922	922
0	732	732
1	0	0
0	683	683
0	777	777
0	555	555
0	932	932
0	885	885

0	537	537
0	868	868
0	987	987
0	655	655
0	803	803
0	930	930
0	785	785
0	983	983
1	0	0
0	633	633
0	673	673
0	682	682
0	953	953
0	670	670
0	855	855
0	853	753

0	905	805
0	875	775
1	0	0
0	963	863
1	0	0
1	0	0
0	990	890
0	945	845
1	0	0
1	0	0
0	790	690
0	695	595
0	840	740
0	882	782
1	0	0
0	1027	927

0	1082	982
0	648	548
0	740	640
0	932	832
0	998	898
0	1087	987
0	738	738
0	762	662
0	1007	907
0	670	570
0	898	798
0	718	618
0	808	708
0	900	700
1	0	0
0	918	718

1	0	0
0	927	727
1	0	0
1	0	0
0	1025	825
0	1018	818
1	0	0
1	0	0
0	1027	827
1	0	0
0	917	717
0	992	792
1	0	0
1	0	0
0	1187	987
1	0	0

0	0	1
942	1142	0
0	0	1
990	1190	0
907	1007	0
900	1100	0
878	1078	0
0	0	1
748	948	0
0	0	1
0	0	1
798	1098	0
828	828	0
0	0	1
797	797	0
790	1090	0

0	647	647
1	0	0
0	1260	960
0	1052	752
0	692	692
0	660	660
0	1012	712
0	517	517
0	1105	805
0	1088	788
0	585	585
0	730	730
0	1300	1000
0	623	623
0	825	825
0	1150	850

722	722	0
990	1290	0
788	988	0
860	1160	0
853	1153	0
780	780	0
830	1130	0
795	795	0
787	787	0
935	1335	0
787	887	0
0	0	1
810	910	0
860	1260	0
962	1062	0
0	0	1

923	1323	0
782	1182	0
773	873	0
785	885	0
658	1058	0
777	877	0
933	1333	0
805	1205	0
0	0	1
848	948	0
990	1390	0
833	933	0
808	908	0
0	0	1
847	947	0
1000	1400	0

1	0	0
0	1190	790
0	1193	793
0	835	735
0	1352	952
0	865	765
1	0	0
0	1408	908
0	1128	928
0	893	893
0	1105	905
0	1443	943
0	1068	868
1	0	0
0	1433	933
0	1338	838

0	983	783
0	1053	853
0	1385	885
0	1113	913
0	1400	900
0	1435	935
0	945	945
0	1083	883
0	1500	1000
0	1088	888
0	1140	940
0	910	910
0	1120	920
0	1485	985
0	808	808
0	1415	915

0	1423	923
0	1125	925
0	1435	935
0	1073	873
0	908	908
0	1397	897
0	992	692
0	962	862
0	962	662
0	1398	898
0	1185	885
1	0	0
0	1325	825
0	1160	660
0	978	678
0	1108	808

0	1363	863
0	1092	792
0	1400	900
0	1360	860
0	882	782
0	1260	960
0	1483	983
0	1113	813
0	1063	763
1	0	0
0	1157	857
0	1488	988
0	640	540
1	0	0
0	1393	893
0	1162	862

0	1412	912
0	1165	865
0	775	675
0	1335	835
1	0	0
0	780	580
0	1087	687
0	1253	753
0	1258	858
1	0	0
0	1398	898
0	1120	620
0	917	517
0	1130	730
0	1110	610
1	0	0

0	1343	843
0	1340	840
0	920	720
0	1307	907
0	1492	992
0	1248	848
0	1162	762
0	783	783
0	1218	818
0	1500	1000
0	793	593
1	0	0
0	1292	792
0	1175	775
0	1362	862
0	1137	737

0	902	702
0	1450	950
0	912	912
1	0	0
0	1457	957
0	1468	968
0	1222	722
1	0	0
0	1365	865
0	1447	947
0	1448	948
0	1142	642
0	1435	935
0	952	952
0	1452	952
0	1393	893

0	1242	942
0	1465	965
0	1500	1000
0	1417	917
0	1462	962
0	1042	942
0	1420	920
0	1500	1000
0	1142	842
0	938	938
0	1463	963
0	1463	963
0	1470	970
0	1455	955
0	1002	702
0	1420	920

0	0	1
0	0	1
963	1463	0
830	1330	0
928	1428	0
0	0	1
885	1385	0
943	1443	0
0	0	1
895	1395	0
845	1345	0
0	0	1
925	1425	0
798	1298	0
725	1125	0
960	1460	0

0	1485	985
0	1415	915
0	1340	840
1	0	0
0	1293	793
0	1488	988
0	1140	740
0	838	738
0	1168	668
0	1150	650
0	1448	948
0	1368	868
0	1265	865

Current Total Score (points)	Answer Time (%)
917	16.67%
850	30.00%
877	24.67%
947	10.67%
765	47.00%
675	65.00%
0	91.67%
922	15.67%
732	53.67%
0	89.33%
683	63.33%
777	44.67%
555	89.00%
932	13.67%
885	23.00%

537	92.67%
868	26.33%
987	2.67%
655	69.00%
803	39.33%
930	14.00%
785	43.00%
983	3.33%
0	18.67%
633	73.33%
673	65.33%
682	63.67%
953	9.33%
670	66.00%
855	29.00%
1770	49.33%

1755	39.00%
1752	45.00%
947	18.33%
1728	27.33%
675	95.33%
0	100.00%
1912	22.00%
1677	31.00%
0	44.67%
683	73.67%
1567	62.00%
1250	81.00%
1772	52.00%
1767	43.67%
537	42.67%
1895	14.67%

2069	3.67%
1303	90.33%
1543	72.00%
1862	33.67%
1783	20.33%
2070	2.67%
738	52.33%
1395	67.67%
1680	18.67%
1352	86.00%
1851	40.33%
1388	76.33%
1663	58.33%
2670	60.00%
1755	28.67%
2670	56.33%

947	19.00%
2655	54.67%
675	73.33%
0	100.00%
2937	35.00%
2695	36.33%
0	96.00%
683	28.00%
2594	34.67%
1250	89.67%
2689	56.67%
2759	41.67%
537	88.00%
1895	20.67%
3256	2.67%
1303	78.67%

1543	18.33%
3004	11.67%
1783	23.33%
3260	2.00%
1745	18.67%
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2758	24.33%
1352	17.00%
2799	50.33%
1388	20.33%
1663	23.67%
3768	40.33%
2583	34.33%
2670	100.00%
1744	40.67%
3745	42.00%

1322	70.67%
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4197	8.00%
3747	49.67%
692	61.67%
1343	68.00%
3606	57.67%
1767	96.67%
3794	39.00%
3847	42.33%
1122	83.00%
2625	54.00%
4556	1.33%
1926	75.33%
2368	35.00%
4154	30.00%

2505	55.67%
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2733	42.33%
3655	28.00%
3911	29.33%
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2183	41.00%
2450	42.67%
5103	13.00%
3470	42.67%
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2654	38.00%
5005	28.00%
2384	7.67%
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5064	51.67%
5729	28.67%
8923	2.33%
4181	92.00%
6260	41.33%
7920	21.33%
5254	27.67%

8128	17.67%
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4133	65.00%
9243	33.00%
5590	33.33%
5305	84.00%
5808	62.67%
9099	49.33%
5895	28.33%
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9676	20.33%
8547	76.00%
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Answer Time (seconds)	
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