

Shielding Effect

| | |
|-------------|-------------|
| Played on | 13 Nov 2019 |
| Hosted by | JenKrug |
| Played with | 30 players |
| Played | 10 of 10 |

Overall Performance

| | |
|-----------------------------|--------|
| Total correct answers (%) | 84,00% |
| Total incorrect answers (%) | 16,00% |
| Average score (points) | 9183,1 |

Feedback



| | |
|----------------------------|---|
| Number of responses | 2 |
| How fun was it? (out of 5) | 0,00 o |
| Did you learn something? | 0,00% |
| Do you recommend it? | 0,00% |
| How do you feel? |  |

Switch tabs/pages to view other result breakdown

Overview

| |
|--|
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| |
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| |
|-----------|
| |
| % |
| % |
| 13 points |

| | | | |
|----------------|---|---------------|---|
| | | | |
| | | | |
| ut of 5 | | | |
| Yes | 0,00% No | | |
| Yes | 0,00% No | | |
| 0,00% Positive |  | 0,00% Neutral |  |

| |
|--|
| |
|--|

Overview

100,00% Negative

Final Scores

| Shielding Effect | |
|------------------|-----------------|
| Final Scores | |
| Rank | Players |
| 1 | maggie |
| 2 | ashley |
| 3 | N8 Baker |
| 4 | sophia |
| 5 | Wale (lil duub) |
| 6 | (((Jajuan))) |
| 7 | Gayatri |
| 8 | Will |
| 9 | pravleen |
| 10 | Reba |
| 11 | Sriveena |
| 12 | alex demchenko |
| 13 | julia |
| 14 | duncan |
| 15 | zoe |
| 16 | conner parker |
| 17 | riya |
| 18 | Dhanshree |
| 19 | Luke |
| 20 | Shrey |
| 21 | nya |
| 22 | ?? ronojoy ?? |

Final Scores

| | |
|----|-------------|
| 23 | matthew |
| 24 | hunter |
| 25 | Andrew F |
| 26 | Will Paasch |
| 27 | Sam S |
| 28 | Timothy |
| 29 | Arnav(>^<) |
| 30 | Mo The Pro |

Final Scores

| Total Score (points) | Correct Answers | Incorrect Answers |
|----------------------|-----------------|-------------------|
| 13411 | 10 | 0 |
| 13406 | 10 | 0 |
| 12426 | 10 | 0 |
| 12408 | 10 | 0 |
| 12147 | 10 | 0 |
| 12113 | 10 | 0 |
| 11897 | 10 | 0 |
| 11878 | 10 | 0 |
| 11843 | 10 | 0 |
| 11437 | 10 | 0 |
| 11302 | 10 | 0 |
| 10148 | 9 | 1 |
| 9660 | 9 | 1 |
| 9443 | 9 | 1 |
| 9246 | 9 | 1 |
| 9140 | 9 | 1 |
| 9042 | 9 | 1 |
| 8728 | 8 | 2 |
| 8545 | 8 | 2 |
| 8056 | 8 | 2 |
| 8036 | 8 | 2 |
| 7302 | 8 | 2 |

Final Scores

| | | |
|------|---|----|
| 7256 | 8 | 2 |
| 6889 | 7 | 3 |
| 6502 | 7 | 3 |
| 6236 | 7 | 3 |
| 5891 | 6 | 4 |
| 5801 | 7 | 3 |
| 5305 | 6 | 4 |
| 0 | 0 | 10 |

Shielding Effect

Kahoot! Summary

| Rank | Players |
|------|-----------------|
| 1 | maggie |
| 2 | ashley |
| 3 | N8 Baker |
| 4 | sophia |
| 5 | Wale (lil duub) |
| 6 | ((Jajuan))) |
| 7 | Gayatri |
| 8 | Will |
| 9 | pravleen |
| 10 | Reba |
| 11 | Sriveena |
| 12 | alex demchenko |
| 13 | julia |
| 14 | duncan |
| 15 | zoe |

Kahoot! Summary

| | |
|----|---------------|
| 16 | conner parker |
| 17 | riya |
| 18 | Dhanshree |
| 19 | Luke |
| 20 | Shrey |
| 21 | nya |
| 22 | ?? ronojoy ?? |
| 23 | matthew |
| 24 | hunter |
| 25 | Andrew F |
| 26 | Will Paasch |
| 27 | Sam S |
| 28 | Timothy |
| 29 | Arnav(>^<) |
| 30 | Mo The Pro |

Kahoot! Summary

| Total Score (points) | Q1 |
|----------------------|-----|
| 13411 | 983 |
| 13406 | 987 |
| 12426 | 922 |
| 12408 | 953 |
| 12147 | 932 |
| 12113 | 917 |
| 11897 | 765 |
| 11878 | 885 |
| 11843 | 673 |
| 11437 | 732 |
| 11302 | 777 |
| 10148 | 868 |
| 9660 | 785 |
| 9443 | 803 |
| 9246 | 670 |

Kahoot! Summary

| | |
|------|-----|
| 9140 | 655 |
| 9042 | 682 |
| 8728 | 947 |
| 8545 | 675 |
| 8056 | 683 |
| 8036 | 633 |
| 7302 | 855 |
| 7256 | 0 |
| 6889 | 930 |
| 6502 | 850 |
| 6236 | 537 |
| 5891 | 0 |
| 5801 | 555 |
| 5305 | 877 |
| 0 | 0 |

Kahoot! Summary

| What is shielding ? | Q2 |
|--|------|
| Interference from inner core electrons | 1087 |
| Interference from inner core electrons | 1082 |
| Interference from inner core electrons | 990 |
| Interference from inner core electrons | 898 |
| Interference from inner core electrons | 840 |
| Interference from inner core electrons | 853 |
| Interference from inner core electrons | 963 |
| Interference from inner core electrons | 882 |
| Interference from inner core electrons | 1007 |
| Interference from inner core electrons | 945 |
| Interference from inner core electrons | 790 |
| Interference from inner core electrons | 1027 |
| Interference from inner core electrons | 998 |
| Interference from inner core electrons | 740 |
| Interference from inner core electrons | 718 |

Kahoot! Summary

| | |
|---|-----|
| Interference from inner core electrons | 648 |
| Interference from inner core electrons | 670 |
| Interference from inner core electrons | 0 |
| Interference from inner core electrons | 0 |
| Interference from inner core electrons | 0 |
| Interference from inner core electrons | 762 |
| Interference from inner core electrons | 808 |
| how the negative charge affects the protons | 738 |
| Interference from inner core electrons | 932 |
| Interference from inner core electrons | 905 |
| Interference from inner core electrons | 0 |
| the ratio of electrons to neutrons | 0 |
| Interference from inner core electrons | 695 |
| Interference from inner core electrons | 875 |
| the ratio of electrons to neutrons | 0 |

Kahoot! Summary

| What happens to the shielding trend moving from left to right across a period on the table | Q3 |
|--|------|
| number of inner orbitals stays the same | 1190 |
| number of inner orbitals stays the same | 1187 |
| number of inner orbitals stays the same | 1025 |
| number of inner orbitals stays the same | 948 |
| number of inner orbitals stays the same | 917 |
| number of inner orbitals stays the same | 900 |
| number of inner orbitals stays the same | 927 |
| number of inner orbitals stays the same | 992 |
| number of inner orbitals stays the same | 1078 |
| number of inner orbitals stays the same | 1018 |
| number of inner orbitals stays the same | 1027 |
| number of inner orbitals stays the same | 0 |
| number of inner orbitals stays the same | 0 |
| number of inner orbitals stays the same | 0 |
| number of inner orbitals stays the same | 0 |

Kahoot! Summary

| | |
|---|------|
| number of inner orbitals stays the same | 0 |
| number of inner orbitals stays the same | 0 |
| number of inner orbital increase | 0 |
| number of valence orbitals decreases | 0 |
| number of valence orbitals decreases | 0 |
| number of inner orbitals stays the same | 1100 |
| number of inner orbitals stays the same | 0 |
| number of inner orbitals stays the same | 1007 |
| number of inner orbitals stays the same | 1142 |
| number of inner orbitals stays the same | 0 |
| number of inner orbital increase | 0 |
| number of inner orbital increase | 0 |
| number of inner orbitals stays the same | 0 |
| number of inner orbitals stays the same | 918 |
| | 0 |

Kahoot! Summary

| What happens to the shielding trend moving down a group? | Q4 |
|---|------|
| shielding increases because the atom has more inner core orbitals | 1290 |
| shielding increases because the atom has more inner core orbitals | 1300 |
| shielding increases because the atom has more inner core orbitals | 1260 |
| shielding increases because the atom has more inner core orbitals | 1130 |
| shielding increases because the atom has more inner core orbitals | 1105 |
| shielding increases because the atom has more inner core orbitals | 1098 |
| shielding increases because the atom has more inner core orbitals | 1090 |
| shielding increases because the atom has more inner core orbitals | 1088 |
| shielding increases because the atom has more inner core orbitals | 1153 |
| shielding increases because the atom has more inner core orbitals | 1052 |
| shielding increases because the atom has more inner core orbitals | 1012 |
| The number of orbitals increases | 730 |
| The number of orbitals increases | 722 |
| The number of orbitals increases | 825 |
| The number of orbitals increases | 795 |

Kahoot! Summary

| | |
|---|------|
| The number of orbitals increases | 623 |
| The number of orbitals increases | 780 |
| The number of orbitals increases | 797 |
| The number of orbitals increases | 647 |
| The number of orbitals increases | 660 |
| shielding increases because the atom has more inner core orbitals | 1160 |
| The number of orbitals increases | 787 |
| shielding increases because the atom has more inner core orbitals | 988 |
| shielding increases because the atom has more inner core orbitals | 1150 |
| The number of orbitals increases | 828 |
| The amount of neutrons decreases | 585 |
| The number of orbitals increases | 692 |
| The number of orbitals increases | 517 |
| shielding increases because the atom has more inner core orbitals | 0 |
| | 0 |

Kahoot! Summary

| How does shielding effect valence electrons? | Q5 |
|--|------|
| inner core electrons block attractive force | 1400 |
| inner core electrons block attractive force | 1390 |
| inner core electrons block attractive force | 1323 |
| inner core electrons block attractive force | 1352 |
| inner core electrons block attractive force | 1333 |
| inner core electrons block attractive force | 1335 |
| inner core electrons block attractive force | 1260 |
| inner core electrons block attractive force | 1205 |
| inner core electrons block attractive force | 1193 |
| inner core electrons block attractive force | 1182 |
| inner core electrons block attractive force | 1058 |
| inner core electrons block attractive force | 948 |
| inner core electrons block attractive force | 947 |
| inner core electrons block attractive force | 908 |
| inner core electrons block attractive force | 865 |

Kahoot! Summary

| | |
|---|------|
| inner core electrons block attractive force | 933 |
| inner core electrons block attractive force | 835 |
| inner core electrons block attractive force | 910 |
| inner core electrons block attractive force | 1062 |
| inner core electrons block attractive force | 885 |
| inner core electrons block attractive force | 1190 |
| inner core electrons block attractive force | 0 |
| inner core electrons block attractive force | 0 |
| inner core electrons block attractive force | 0 |
| inner core electrons block attractive force | 887 |
| inner core electrons block attractive force | 0 |
| inner core electrons block attractive force | 873 |
| inner core electrons block attractive force | 877 |
| | 0 |
| | 0 |

Kahoot! Summary

| How does Coulomb's Law affect the shielding trend | Q6 |
|---|------|
| causes valence electrons to fall of and get lost | 1485 |
| causes valence electrons to fall of and get lost | 1500 |
| causes valence electrons to fall of and get lost | 1433 |
| causes valence electrons to fall of and get lost | 1435 |
| causes valence electrons to fall of and get lost | 1400 |
| causes valence electrons to fall of and get lost | 1408 |
| causes valence electrons to fall of and get lost | 1443 |
| causes valence electrons to fall of and get lost | 1435 |
| causes valence electrons to fall of and get lost | 1423 |
| causes valence electrons to fall of and get lost | 1338 |
| causes valence electrons to fall of and get lost | 1385 |
| causes valence electrons to fall of and get lost | 1083 |
| causes valence electrons to fall of and get lost | 1120 |
| causes valence electrons to fall of and get lost | 1140 |
| causes valence electrons to fall of and get lost | 1073 |

Kahoot! Summary

| | |
|--|------|
| causes valence electrons to fall of and get lost | 1088 |
| causes valence electrons to fall of and get lost | 1125 |
| causes valence electrons to fall of and get lost | 1105 |
| causes valence electrons to fall of and get lost | 1068 |
| causes valence electrons to fall of and get lost | 1053 |
| causes valence electrons to fall of and get lost | 1415 |
| the protons increase | 908 |
| the protons increase | 808 |
| the protons increase | 910 |
| causes valence electrons to fall of and get lost | 1128 |
| causes electrons to disintegrate | 945 |
| causes valence electrons to fall of and get lost | 983 |
| causes valence electrons to fall of and get lost | 1113 |
| | 893 |
| | 0 |

Kahoot! Summary

| Which period has the largest value for the shielding trend? | Q7 |
|---|------|
| period 7 bc they have the most orbitals | 1488 |
| period 7 bc they have the most orbitals | 1483 |
| period 7 bc they have the most orbitals | 1325 |
| period 7 bc they have the most orbitals | 1412 |
| period 7 bc they have the most orbitals | 1400 |
| period 7 bc they have the most orbitals | 1397 |
| period 7 bc they have the most orbitals | 1398 |
| period 7 bc they have the most orbitals | 1360 |
| period 7 bc they have the most orbitals | 1393 |
| period 7 bc they have the most orbitals | 1160 |
| period 7 bc they have the most orbitals | 1363 |
| period 7 bc they have the most orbitals | 1260 |
| period 7 bc they have the most orbitals | 1157 |
| period 7 bc they have the most orbitals | 1063 |
| period 7 bc they have the most orbitals | 1165 |

Kahoot! Summary

| | |
|---|------|
| period 7 bc they have the most orbitals | 1113 |
| period 7 bc they have the most orbitals | 1162 |
| period 7 bc they have the most orbitals | 962 |
| period 7 bc they have the most orbitals | 1185 |
| period 7 bc they have the most orbitals | 1108 |
| period 7 bc they have the most orbitals | 0 |
| period 7 bc they have the most orbitals | 775 |
| period 7 bc they have the most orbitals | 640 |
| period 7 bc they have the most orbitals | 0 |
| period 7 bc they have the most orbitals | 992 |
| period 7 bc they have the most orbitals | 882 |
| period 7 bc they have the most orbitals | 978 |
| period 7 bc they have the most orbitals | 1092 |
| period 7 bc they have the most orbitals | 962 |
| | 0 |

Kahoot! Summary

| What is “effective nuclear charge”? | Q8 |
|--|------|
| net positive charge experienced by an electron | 1500 |
| net positive charge experienced by an electron | 1492 |
| net positive charge experienced by an electron | 1398 |
| net positive charge experienced by an electron | 1362 |
| net positive charge experienced by an electron | 1343 |
| net positive charge experienced by an electron | 1335 |
| net positive charge experienced by an electron | 1253 |
| net positive charge experienced by an electron | 1340 |
| net positive charge experienced by an electron | 1292 |
| net positive charge experienced by an electron | 1120 |
| net positive charge experienced by an electron | 1110 |
| net positive charge experienced by an electron | 1307 |
| net positive charge experienced by an electron | 1218 |
| net positive charge experienced by an electron | 1162 |
| net positive charge experienced by an electron | 1137 |

Kahoot! Summary

| | |
|--|------|
| net positive charge experienced by an electron | 1248 |
| net positive charge experienced by an electron | 1175 |
| net positive charge experienced by an electron | 1087 |
| net positive charge experienced by an electron | 1258 |
| net positive charge experienced by an electron | 1130 |
| net negative charge of the electron | 0 |
| net positive charge experienced by an electron | 902 |
| net positive charge experienced by an electron | 793 |
| net negative charge of the electron | 783 |
| net positive charge experienced by an electron | 0 |
| net positive charge experienced by an electron | 920 |
| net positive charge experienced by an electron | 917 |
| net positive charge experienced by an electron | 0 |
| net positive charge experienced by an electron | 780 |
| | 0 |

Kahoot! Summary

| What is the net positive charge? | Q9 |
|--|------|
| how many availbe protons are acting on valence electrons | 1500 |
| how many availbe protons are acting on valence electrons | 1500 |
| how many availbe protons are acting on valence electrons | 1365 |
| how many availbe protons are acting on valence electrons | 1470 |
| how many availbe protons are acting on valence electrons | 1452 |
| how many availbe protons are acting on valence electrons | 1450 |
| how many availbe protons are acting on valence electrons | 1468 |
| how many availbe protons are acting on valence electrons | 1393 |
| how many availbe protons are acting on valence electrons | 1463 |
| how many availbe protons are acting on valence electrons | 1447 |
| how many availbe protons are acting on valence electrons | 1435 |
| how many availbe protons are acting on valence electrons | 1465 |
| how many availbe protons are acting on valence electrons | 1420 |
| how many availbe protons are acting on valence electrons | 1462 |
| how many availbe protons are acting on valence electrons | 1455 |

Kahoot! Summary

| | |
|--|------|
| how many availbe protons are acting on valence electrons | 1417 |
| how many availbe protons are acting on valence electrons | 1463 |
| how many availbe protons are acting on valence electrons | 1457 |
| how many availbe protons are acting on valence electrons | 1222 |
| how many availbe protons are acting on valence electrons | 1142 |
| how many electrons and in orbitals | 938 |
| how many availbe protons are acting on valence electrons | 1002 |
| how many availbe protons are acting on valence electrons | 1142 |
| how many availbe protons are acting on valence electrons | 1042 |
| the average number of protons | 912 |
| how many availbe protons are acting on valence electrons | 1242 |
| how many availbe protons are acting on valence electrons | 1448 |
| how many periods are in the periodic table | 952 |
| how many availbe protons are acting on valence electrons | 0 |
| | 0 |

Kahoot! Summary

| Which group has the smallest value for the shielding trend? | Q10 |
|---|------|
| Period 1, it has the least amount of orbitals | 1488 |
| Period 1, it has the least amount of orbitals | 1485 |
| Period 1, it has the least amount of orbitals | 1385 |
| Period 1, it has the least amount of orbitals | 1448 |
| Period 1, it has the least amount of orbitals | 1425 |
| Period 1, it has the least amount of orbitals | 1420 |
| Period 1, it has the least amount of orbitals | 1330 |
| Period 1, it has the least amount of orbitals | 1298 |
| Period 1, it has the least amount of orbitals | 1168 |
| Period 1, it has the least amount of orbitals | 1443 |
| Period 1, it has the least amount of orbitals | 1345 |
| Period 1, it has the least amount of orbitals | 1460 |
| Period 1, it has the least amount of orbitals | 1293 |
| Period 1, it has the least amount of orbitals | 1340 |
| Period 1, it has the least amount of orbitals | 1368 |

Kahoot! Summary

| | |
|---|------|
| Period 1, it has the least amount of orbitals | 1415 |
| Period 1, it has the least amount of orbitals | 1150 |
| Period 1, it has the least amount of orbitals | 1463 |
| Period 1, it has the least amount of orbitals | 1428 |
| Period 1, it has the least amount of orbitals | 1395 |
| Period 1, it has the least amount of orbitals | 838 |
| Period 1, it has the least amount of orbitals | 1265 |
| Period 1, it has the least amount of orbitals | 1140 |
| Period 1, it has the least amount of orbitals | 0 |
| Period 1, it has the least amount of orbitals | 0 |
| Period 1, it has the least amount of orbitals | 1125 |
| Period 1, it has the least amount of orbitals | 0 |
| Period 1, it has the least amount of orbitals | 0 |
| | 0 |
| | 0 |

Kahoot! Summary

| |
|---|
| |
| |
| Shielding effect affects the neutrons and protons |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |

Kahoot! Summary

| |
|-------|
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| False |
| True |
| True |
| False |
| True |
| True |
| |
| |

| Shielding |
|-------------------|
| 1 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

1 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|---------------------|-----------|
| What is shielding ? | |
| s | Interfere |
| (%) | 90,00% |
| on | 30 secur |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |

| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✗ |

1 Quiz

| | |
|---|---|
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
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| | ✓ |
|) | ✓ |
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| |
|-------------------------------|
| |
| |
| nce from inner core electrons |
| |
| nds |

| | |
|--|---|
| | |
| Interference from inner core electrons | ◆ |
| ✓ | |
| 27 | |
| 12,16 | |

| | |
|--|----------|
| | Score (p |
| Interference from inner core electrons | 917 |
| Interference from inner core electrons | 850 |
| Interference from inner core electrons | 877 |
| Interference from inner core electrons | 947 |
| Interference from inner core electrons | 765 |
| Interference from inner core electrons | 675 |
| the ratio of electrons to neutrons | 0 |
| Interference from inner core electrons | 922 |
| Interference from inner core electrons | 732 |
| the ratio of electrons to neutrons | 0 |

1 Quiz

| | |
|---|-----|
| Interference from inner core electrons | 683 |
| Interference from inner core electrons | 777 |
| Interference from inner core electrons | 555 |
| Interference from inner core electrons | 932 |
| Interference from inner core electrons | 885 |
| Interference from inner core electrons | 537 |
| Interference from inner core electrons | 868 |
| Interference from inner core electrons | 987 |
| Interference from inner core electrons | 655 |
| Interference from inner core electrons | 803 |
| Interference from inner core electrons | 930 |
| Interference from inner core electrons | 785 |
| Interference from inner core electrons | 983 |
| how the negative charge affects the protons | 0 |
| Interference from inner core electrons | 633 |
| Interference from inner core electrons | 673 |
| Interference from inner core electrons | 682 |
| Interference from inner core electrons | 953 |
| Interference from inner core electrons | 670 |
| Interference from inner core electrons | 855 |

1 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--|----------------------------------|
| | |
| when electrons dont fit on the orbital | <input checked="" type="radio"/> |
| X | |
| 0 | |
| 0,00 | |

| | |
|--------|---------|
| | |
| oints) | Current |
| | 917 |
| | 850 |
| | 877 |
| | 947 |
| | 765 |
| | 675 |
| | 0 |
| | 922 |
| | 732 |
| | 0 |

1 Quiz

| | |
|--|-----|
| | 683 |
| | 777 |
| | 555 |
| | 932 |
| | 885 |
| | 537 |
| | 868 |
| | 987 |
| | 655 |
| | 803 |
| | 930 |
| | 785 |
| | 983 |
| | 0 |
| | 633 |
| | 673 |
| | 682 |
| | 953 |
| | 670 |
| | 855 |

1 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|---|--------------------------|
| | |
| how the negative charge affects the protons | <input type="checkbox"/> |
| X | |
| 1 | |
| 5,60 | |

| | |
|----------------------|-----------|
| | |
| Total Score (points) | Answer ti |
| | 5 |
| | 9 |
| | 7,4 |
| | 3,2 |
| | 14,1 |
| | 19,5 |
| | 27,5 |
| | 4,7 |
| | 16,1 |
| | 26,8 |

1 Quiz

| | |
|--|------|
| | 19 |
| | 13,4 |
| | 26,7 |
| | 4,1 |
| | 6,9 |
| | 27,8 |
| | 7,9 |
| | 0,8 |
| | 20,7 |
| | 11,8 |
| | 4,2 |
| | 12,9 |
| | 1 |
| | 5,6 |
| | 22 |
| | 19,6 |
| | 19,1 |
| | 2,8 |
| | 19,8 |
| | 8,7 |

The chart displays the ratio of electrons to neutrons over time. The y-axis represents the ratio, and the x-axis represents time in seconds. A significant event is marked at 27.15 seconds, where the ratio increases sharply, indicated by a red bar with a white 'X'.

| Time (seconds) | Ratio of electrons to neutrons |
|----------------|--------------------------------|
| 0 | ~0.1 |
| ~27.15 | ~0.2 |
| ~27.15 | ~0.3 |
| ~27.15 | ~0.4 |
| ~27.15 | ~0.5 |
| ~27.15 | ~0.6 |
| ~27.15 | ~0.7 |
| ~27.15 | ~0.8 |
| ~27.15 | ~0.9 |
| ~27.15 | ~1.0 |
| ~27.15 | ~1.1 |
| ~27.15 | ~1.2 |
| ~27.15 | ~1.3 |
| ~27.15 | ~1.4 |
| ~27.15 | ~1.5 |
| ~27.15 | ~1.6 |
| ~27.15 | ~1.7 |
| ~27.15 | ~1.8 |
| ~27.15 | ~1.9 |
| ~27.15 | ~2.0 |
| ~27.15 | ~2.1 |
| ~27.15 | ~2.2 |
| ~27.15 | ~2.3 |
| ~27.15 | ~2.4 |
| ~27.15 | ~2.5 |
| ~27.15 | ~2.6 |
| ~27.15 | ~2.7 |
| ~27.15 | ~2.8 |
| ~27.15 | ~2.9 |
| ~27.15 | ~3.0 |
| ~27.15 | ~3.1 |
| ~27.15 | ~3.2 |
| ~27.15 | ~3.3 |
| ~27.15 | ~3.4 |
| ~27.15 | ~3.5 |
| ~27.15 | ~3.6 |
| ~27.15 | ~3.7 |
| ~27.15 | ~3.8 |
| ~27.15 | ~3.9 |
| ~27.15 | ~4.0 |
| ~27.15 | ~4.1 |
| ~27.15 | ~4.2 |
| ~27.15 | ~4.3 |
| ~27.15 | ~4.4 |
| ~27.15 | ~4.5 |
| ~27.15 | ~4.6 |
| ~27.15 | ~4.7 |
| ~27.15 | ~4.8 |
| ~27.15 | ~4.9 |
| ~27.15 | ~5.0 |
| ~27.15 | ~5.1 |
| ~27.15 | ~5.2 |
| ~27.15 | ~5.3 |
| ~27.15 | ~5.4 |
| ~27.15 | ~5.5 |
| ~27.15 | ~5.6 |
| ~27.15 | ~5.7 |
| ~27.15 | ~5.8 |
| ~27.15 | ~5.9 |
| ~27.15 | ~6.0 |
| ~27.15 | ~6.1 |
| ~27.15 | ~6.2 |
| ~27.15 | ~6.3 |
| ~27.15 | ~6.4 |
| ~27.15 | ~6.5 |
| ~27.15 | ~6.6 |
| ~27.15 | ~6.7 |
| ~27.15 | ~6.8 |
| ~27.15 | ~6.9 |
| ~27.15 | ~7.0 |
| ~27.15 | ~7.1 |
| ~27.15 | ~7.2 |
| ~27.15 | ~7.3 |
| ~27.15 | ~7.4 |
| ~27.15 | ~7.5 |
| ~27.15 | ~7.6 |
| ~27.15 | ~7.7 |
| ~27.15 | ~7.8 |
| ~27.15 | ~7.9 |
| ~27.15 | ~8.0 |
| ~27.15 | ~8.1 |
| ~27.15 | ~8.2 |
| ~27.15 | ~8.3 |
| ~27.15 | ~8.4 |
| ~27.15 | ~8.5 |
| ~27.15 | ~8.6 |
| ~27.15 | ~8.7 |
| ~27.15 | ~8.8 |
| ~27.15 | ~8.9 |
| ~27.15 | ~9.0 |
| ~27.15 | ~9.1 |
| ~27.15 | ~9.2 |
| ~27.15 | ~9.3 |
| ~27.15 | ~9.4 |
| ~27.15 | ~9.5 |
| ~27.15 | ~9.6 |
| ~27.15 | ~9.7 |
| ~27.15 | ~9.8 |
| ~27.15 | ~9.9 |
| ~27.15 | ~10.0 |

1 Quiz

[illegible]

| Shielding |
|-------------------|
| 2 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

2 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|--|-------------------|
| What happens to the shielding trend moving from left to right across the periodic table? | |
| Shielding effect | number of protons |
| Shielding effect (%) | 80,00% |
| Time taken to answer (seconds) | 30 seconds |

| Summary | |
|--------------------------------|---|
| Score | ▲ |
| Correct? | |
| Questions received | |
| Time taken to answer (seconds) | |

| Details | |
|--|--------|
| Question | Answer |
| What happens to the shielding trend moving from left to right across the periodic table? | ✓ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✓ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✓ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✗ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✓ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✗ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✗ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✓ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✓ |
| What happens to the shielding trend moving from left to right across the periodic table? | ✗ |

2 Quiz

| | |
|---|---|
| | X |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | X |
|) | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |

| |
|----------------------------------|
| |
| ross a period on the table |
| of inner orbitals stays the same |
| |
| nds |

| | |
|----------------------------------|---|
| | |
| number of inner orbital increase | ◆ |
| X | |
| 3 | |
| 10,57 | |

| | |
|---|----------|
| | Score (p |
| number of inner orbitals stays the same | 853 |
| number of inner orbitals stays the same | 905 |
| number of inner orbitals stays the same | 875 |
| number of inner orbital increase | 0 |
| number of inner orbitals stays the same | 963 |
| number of valence orbitals decreases | 0 |
| | 0 |
| number of inner orbitals stays the same | 990 |
| number of inner orbitals stays the same | 945 |
| number of inner orbital increase | 0 |

2 Quiz

| | |
|---|------|
| number of valence orbitals decreases | 0 |
| number of inner orbitals stays the same | 790 |
| number of inner orbitals stays the same | 695 |
| number of inner orbitals stays the same | 840 |
| number of inner orbitals stays the same | 882 |
| number of inner orbital increase | 0 |
| number of inner orbitals stays the same | 1027 |
| number of inner orbitals stays the same | 1082 |
| number of inner orbitals stays the same | 648 |
| number of inner orbitals stays the same | 740 |
| number of inner orbitals stays the same | 932 |
| number of inner orbitals stays the same | 998 |
| number of inner orbitals stays the same | 1087 |
| number of inner orbitals stays the same | 738 |
| number of inner orbitals stays the same | 762 |
| number of inner orbitals stays the same | 1007 |
| number of inner orbitals stays the same | 670 |
| number of inner orbitals stays the same | 898 |
| number of inner orbitals stays the same | 718 |
| number of inner orbitals stays the same | 808 |

2 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--------------------------------------|----------------------------------|
| | |
| number of valence orbitals increases | <input checked="" type="radio"/> |
| X | <input type="radio"/> |
| 0 | |
| 0,00 | |

| (points) | Current |
|----------|---------|
| | 1770 |
| | 1755 |
| | 1752 |
| | 947 |
| | 1728 |
| | 675 |
| | 0 |
| | 1912 |
| | 1677 |
| | 0 |

2 Quiz

| | |
|--|------|
| | 683 |
| | 1567 |
| | 1250 |
| | 1772 |
| | 1767 |
| | 537 |
| | 1895 |
| | 2069 |
| | 1303 |
| | 1543 |
| | 1862 |
| | 1783 |
| | 2070 |
| | 738 |
| | 1395 |
| | 1680 |
| | 1352 |
| | 1851 |
| | 1388 |
| | 1663 |

2 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|---|------------------------|
| | |
| number of inner orbitals stays the same | <div><div></div></div> |
| <div>✓</div> | |
| 24 | |
| 13,62 | |

| | |
|----------------------|-----------|
| | |
| Total Score (points) | Answer ti |
| | 14,8 |
| | 11,7 |
| | 13,5 |
| | 5,5 |
| | 8,2 |
| | 28,6 |
| | 30 |
| | 6,6 |
| | 9,3 |
| | 13,4 |

2 Quiz

| | |
|--|------|
| | 22,1 |
| | 18,6 |
| | 24,3 |
| | 15,6 |
| | 13,1 |
| | 12,8 |
| | 4,4 |
| | 1,1 |
| | 27,1 |
| | 21,6 |
| | 10,1 |
| | 6,1 |
| | 0,8 |
| | 15,7 |
| | 20,3 |
| | 5,6 |
| | 25,8 |
| | 12,1 |
| | 22,9 |
| | 17,5 |

| Shielding |
|-------------------|
| 3 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

3 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|--|-----------|
| What happens to the shielding trend moving down a group? | |
| s | shielding |
| (%) | 50,00% |
| on | 30 secur |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |

| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | X |
| | ✓ |
| | X |
| | ✓ |
| | X |
| | X |
| | ✓ |
| | ✓ |
| | X |

3 Quiz

| | |
|---|---|
| | X |
| | ✓ |
| | X |
| | ✓ |
| | ✓ |
| | X |
|) | X |
| | ✓ |
| | X |
| | X |
| | ✓ |
| | X |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | X |
| | ✓ |
| | X |
| | X |

3 Quiz

| |
|---|
| |
| |
| g increases because the atom has more inner core orbitals |
| |
| nds |
| |

| | |
|----------------------------------|---|
| | |
| The amount of neutrons decreases | ◆ |
| X | |
| 1 | |
| 26,40 | |
| | |

| | |
|---|----------|
| | Score (p |
| shielding increases because the atom has more inner core orbitals | 900 |
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 918 |
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 927 |
| The number of orbitals increases | 0 |
| | 0 |
| shielding increases because the atom has more inner core orbitals | 1025 |
| shielding increases because the atom has more inner core orbitals | 1018 |
| The number of orbitals increases | 0 |

3 Quiz

| | |
|---|------|
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 1027 |
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 917 |
| shielding increases because the atom has more inner core orbitals | 992 |
| The amount of neutrons decreases | 0 |
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 1187 |
| The number of orbitals increases | 0 |
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 1142 |
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 1190 |
| shielding increases because the atom has more inner core orbitals | 1007 |
| shielding increases because the atom has more inner core orbitals | 1100 |
| shielding increases because the atom has more inner core orbitals | 1078 |
| The number of orbitals increases | 0 |
| shielding increases because the atom has more inner core orbitals | 948 |
| The number of orbitals increases | 0 |
| The number of orbitals increases | 0 |

3 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--------------------|----------------------------------|
| | |
| The mass increases | <input checked="" type="radio"/> |
| X | |
| 0 | |
| 0,00 | |

| (points) | Current |
|----------|---------|
| | 2670 |
| | 1755 |
| | 2670 |
| | 947 |
| | 2655 |
| | 675 |
| | 0 |
| | 2937 |
| | 2695 |
| | 0 |

3 Quiz

| | |
|--|------|
| | 683 |
| | 2594 |
| | 1250 |
| | 2689 |
| | 2759 |
| | 537 |
| | 1895 |
| | 3256 |
| | 1303 |
| | 1543 |
| | 3004 |
| | 1783 |
| | 3260 |
| | 1745 |
| | 2495 |
| | 2758 |
| | 1352 |
| | 2799 |
| | 1388 |
| | 1663 |

3 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|----------------------------------|--------------------------|
| | |
| The number of orbitals increases | <input type="checkbox"/> |
| X | |
| 13 | |
| 12,38 | |

| | |
|----------------------|-----------|
| | |
| Total Score (points) | Answer ti |
| | 18 |
| | 8,6 |
| | 16,9 |
| | 5,7 |
| | 16,4 |
| | 22 |
| | 30 |
| | 10,5 |
| | 10,9 |
| | 28,8 |

3 Quiz

| | |
|--|------|
| | 8,4 |
| | 10,4 |
| | 26,9 |
| | 17 |
| | 12,5 |
| | 26,4 |
| | 6,2 |
| | 0,8 |
| | 23,6 |
| | 5,5 |
| | 3,5 |
| | 7 |
| | 0,6 |
| | 5,6 |
| | 6 |
| | 7,3 |
| | 5,1 |
| | 15,1 |
| | 6,1 |
| | 7,1 |

3 Quiz

shielding increases because the atom has more inner core orbitals

☒

15

10,10

time (seconds)

3 Quiz

[illegible]

| Shielding |
|-------------------|
| 4 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

4 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|--|-----------|
| How does shielding effect valence electrons? | |
| s | inner cor |
| (%) | 93,33% |
| on | 30 secur |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |



| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |

4 Quiz

| | |
|---|---|
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
|) | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |

4 Quiz

| |
|--------------------------------------|
| |
| |
| the electrons block attractive force |
| |
| nds |

| | |
|--|--|
| | |
| smaller electrons fall off the orbital |  |
| X |  |
| 0 | |
| 0,00 | |

| | |
|---|----------|
| | Score (p |
| inner core electrons block attractive force | 1098 |
| inner core electrons block attractive force | 828 |
| | 0 |
| inner core electrons block attractive force | 797 |
| inner core electrons block attractive force | 1090 |
| inner core electrons block attractive force | 647 |
| | 0 |
| inner core electrons block attractive force | 1260 |
| inner core electrons block attractive force | 1052 |
| inner core electrons block attractive force | 692 |

4 Quiz

| | |
|---|------|
| inner core electrons block attractive force | 660 |
| inner core electrons block attractive force | 1012 |
| inner core electrons block attractive force | 517 |
| inner core electrons block attractive force | 1105 |
| inner core electrons block attractive force | 1088 |
| inner core electrons block attractive force | 585 |
| inner core electrons block attractive force | 730 |
| inner core electrons block attractive force | 1300 |
| inner core electrons block attractive force | 623 |
| inner core electrons block attractive force | 825 |
| inner core electrons block attractive force | 1150 |
| inner core electrons block attractive force | 722 |
| inner core electrons block attractive force | 1290 |
| inner core electrons block attractive force | 988 |
| inner core electrons block attractive force | 1160 |
| inner core electrons block attractive force | 1153 |
| inner core electrons block attractive force | 780 |
| inner core electrons block attractive force | 1130 |
| inner core electrons block attractive force | 795 |
| inner core electrons block attractive force | 787 |

4 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|---|------------------------|
| | |
| inner core electrons block attractive force | <div><div></div></div> |
| <div>✔</div> | |
| 28 | |
| 13,38 | |

| | |
|--------|---------|
| | |
| oints) | Current |
| | 3768 |
| | 2583 |
| | 2670 |
| | 1744 |
| | 3745 |
| | 1322 |
| | 0 |
| | 4197 |
| | 3747 |
| | 692 |

4 Quiz

| | |
|--|------|
| | 1343 |
| | 3606 |
| | 1767 |
| | 3794 |
| | 3847 |
| | 1122 |
| | 2625 |
| | 4556 |
| | 1926 |
| | 2368 |
| | 4154 |
| | 2505 |
| | 4550 |
| | 2733 |
| | 3655 |
| | 3911 |
| | 2132 |
| | 3929 |
| | 2183 |
| | 2450 |

4 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--|--------------------------|
| | |
| protons are not related to valence electrons | <input type="checkbox"/> |
| X | |
| 0 | |
| 0,00 | |

| | |
|----------------------|-----------|
| | |
| Total Score (points) | Answer ti |
| | 12,1 |
| | 10,3 |
| | 30 |
| | 12,2 |
| | 12,6 |
| | 21,2 |
| | 30 |
| | 2,4 |
| | 14,9 |
| | 18,5 |

4 Quiz

| | |
|--|------|
| | 20,4 |
| | 17,3 |
| | 29 |
| | 11,7 |
| | 12,7 |
| | 24,9 |
| | 16,2 |
| | 0,4 |
| | 22,6 |
| | 10,5 |
| | 9 |
| | 16,7 |
| | 0,6 |
| | 12,7 |
| | 8,4 |
| | 8,8 |
| | 13,2 |
| | 10,2 |
| | 12,3 |
| | 12,8 |

4 Quiz

Time (seconds)

0

0,00

Gravity

Electromagnetic force

Strong nuclear force

Weak nuclear force

Protons are weaker than valence electrons

X

| Shielding |
|-------------------|
| 5 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

5 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|---|----------|
| How does Coulomb's Law affect the shielding trend | |
| s | causes v |
| (%) | 80,00% |
| on | 30 secur |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |

| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |

5 Quiz

| | |
|---|---|
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
|) | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |

5 Quiz

| |
|---|
| |
| |
| valence electrons to fall of and get lost |
| |
| nds |

| | |
|-------------------------------------|---|
| | |
| the element changes shape and color | ◆ |
| X | |
| 0 | |
| 0,00 | |

| | |
|--|----------|
| | Score (p |
| causes valence electrons to fall of and get lost | 1335 |
| causes valence electrons to fall of and get lost | 887 |
| | 0 |
| causes valence electrons to fall of and get lost | 910 |
| causes valence electrons to fall of and get lost | 1260 |
| causes valence electrons to fall of and get lost | 1062 |
| | 0 |
| causes valence electrons to fall of and get lost | 1323 |
| causes valence electrons to fall of and get lost | 1182 |
| causes valence electrons to fall of and get lost | 873 |

5 Quiz

| | |
|--|------|
| causes valence electrons to fall of and get lost | 885 |
| causes valence electrons to fall of and get lost | 1058 |
| causes valence electrons to fall of and get lost | 877 |
| causes valence electrons to fall of and get lost | 1333 |
| causes valence electrons to fall of and get lost | 1205 |
| causes electrons to disintegrate | 0 |
| causes valence electrons to fall of and get lost | 948 |
| causes valence electrons to fall of and get lost | 1390 |
| causes valence electrons to fall of and get lost | 933 |
| causes valence electrons to fall of and get lost | 908 |
| the protons increase | 0 |
| causes valence electrons to fall of and get lost | 947 |
| causes valence electrons to fall of and get lost | 1400 |
| the protons increase | 0 |
| causes valence electrons to fall of and get lost | 1190 |
| causes valence electrons to fall of and get lost | 1193 |
| causes valence electrons to fall of and get lost | 835 |
| causes valence electrons to fall of and get lost | 1352 |
| causes valence electrons to fall of and get lost | 865 |
| the protons increase | 0 |

5 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--|-------------|
| | |
| causes valence electrons to fall of and get lost | <div></div> |
| <div>✔</div> | |
| 24 | |
| 9,64 | |

| | |
|--------|---------|
| | |
| oints) | Current |
| | 5103 |
| | 3470 |
| | 2670 |
| | 2654 |
| | 5005 |
| | 2384 |
| | 0 |
| | 5520 |
| | 4929 |
| | 1565 |

5 Quiz

| | |
|--|------|
| | 2228 |
| | 4664 |
| | 2644 |
| | 5127 |
| | 5052 |
| | 1122 |
| | 3573 |
| | 5946 |
| | 2859 |
| | 3276 |
| | 4154 |
| | 3452 |
| | 5950 |
| | 2733 |
| | 4845 |
| | 5104 |
| | 2967 |
| | 5281 |
| | 3048 |
| | 2450 |

5 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|----------------------|--------------------------|
| | |
| the protons increase | <input type="checkbox"/> |
| X | |
| 3 | |
| 15,90 | |

| | |
|----------------------|-----------|
| | |
| Total Score (points) | Answer ti |
| | 3,9 |
| | 12,8 |
| | 30 |
| | 11,4 |
| | 8,4 |
| | 2,3 |
| | 30 |
| | 4,6 |
| | 13,1 |
| | 13,6 |

5 Quiz

| | |
|--|------|
| | 12,9 |
| | 20,5 |
| | 13,4 |
| | 4 |
| | 11,7 |
| | 26,5 |
| | 9,1 |
| | 0,6 |
| | 10 |
| | 11,5 |
| | 8,8 |
| | 9,2 |
| | 0,4 |
| | 9,9 |
| | 12,6 |
| | 12,4 |
| | 15,9 |
| | 2,9 |
| | 14,1 |
| | 29 |

5 Quiz

[illegible]

5 Quiz

[illegible]

| Shielding |
|-------------------|
| 6 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

6 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|---|----------|
| Which period has the largest value for the shielding trend? | |
| s | period 7 |
| (%) | 96,67% |
| on | 20 secur |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |

| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |

6 Quiz

| | |
|---|---|
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
|) | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
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| | ✓ |
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| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |

6 Quiz

| |
|--------------------------------|
| |
| |
| bc they have the most orbitals |
| |
| nds |

| | |
|---|---|
| | |
| period 7 bc they have the most orbitals | ◆ |
| ✓ | |
| 29 | |
| 3,83 | |

| | |
|---|----------|
| | Score (p |
| period 7 bc they have the most orbitals | 1408 |
| period 7 bc they have the most orbitals | 1128 |
| period 7 bc they have the most orbitals | 893 |
| period 7 bc they have the most orbitals | 1105 |
| period 7 bc they have the most orbitals | 1443 |
| period 7 bc they have the most orbitals | 1068 |
| | 0 |
| period 7 bc they have the most orbitals | 1433 |
| period 7 bc they have the most orbitals | 1338 |
| period 7 bc they have the most orbitals | 983 |

6 Quiz

| | |
|---|------|
| period 7 bc they have the most orbitals | 1053 |
| period 7 bc they have the most orbitals | 1385 |
| period 7 bc they have the most orbitals | 1113 |
| period 7 bc they have the most orbitals | 1400 |
| period 7 bc they have the most orbitals | 1435 |
| period 7 bc they have the most orbitals | 945 |
| period 7 bc they have the most orbitals | 1083 |
| period 7 bc they have the most orbitals | 1500 |
| period 7 bc they have the most orbitals | 1088 |
| period 7 bc they have the most orbitals | 1140 |
| period 7 bc they have the most orbitals | 910 |
| period 7 bc they have the most orbitals | 1120 |
| period 7 bc they have the most orbitals | 1485 |
| period 7 bc they have the most orbitals | 808 |
| period 7 bc they have the most orbitals | 1415 |
| period 7 bc they have the most orbitals | 1423 |
| period 7 bc they have the most orbitals | 1125 |
| period 7 bc they have the most orbitals | 1435 |
| period 7 bc they have the most orbitals | 1073 |
| period 7 bc they have the most orbitals | 908 |

6 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--|------------------------|
| | |
| period 3 bc they have the most protons | <div><div></div></div> |
| X | |
| 0 | |
| 0,00 | |

| | |
|--------|---------|
| | |
| oints) | Current |
| | 6511 |
| | 4598 |
| | 3563 |
| | 3759 |
| | 6448 |
| | 3452 |
| | 0 |
| | 6953 |
| | 6267 |
| | 2548 |

6 Quiz

| | |
|--|------|
| | 3281 |
| | 6049 |
| | 3757 |
| | 6527 |
| | 6487 |
| | 2067 |
| | 4656 |
| | 7446 |
| | 3947 |
| | 4416 |
| | 5064 |
| | 4572 |
| | 7435 |
| | 3541 |
| | 6260 |
| | 6527 |
| | 4092 |
| | 6716 |
| | 4121 |
| | 3358 |

6 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--|-------------------------------------|
| | |
| period 5 bc they have the most electrons | <input checked="" type="checkbox"/> |
| X | |
| 0 | |
| 0,00 | |

| | |
|----------------------|----------|
| | |
| Total Score (points) | Answer t |
| | 3,7 |
| | 2,9 |
| | 4,3 |
| | 3,8 |
| | 2,3 |
| | 5,3 |
| | 20 |
| | 2,7 |
| | 6,5 |
| | 8,7 |

6 Quiz

| | |
|--|-----|
| | 5,9 |
| | 4,6 |
| | 3,5 |
| | 4 |
| | 2,6 |
| | 2,2 |
| | 4,7 |
| | 0,4 |
| | 4,5 |
| | 2,4 |
| | 3,6 |
| | 3,2 |
| | 0,6 |
| | 7,7 |
| | 3,4 |
| | 3,1 |
| | 3 |
| | 2,6 |
| | 5,1 |
| | 3,7 |

6 Quiz

period 6 bc the have the most neutrons

X

0

0,00

time (seconds)

| Shielding |
|-------------------|
| 7 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

7 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|-------------------------------------|-----------|
| What is “effective nuclear charge”? | |
| is | net posit |
| (%) | 90,00% |
| on | 30 secur |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |

| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |

7 Quiz

| | |
|---|---|
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
|) | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |

7 Quiz

| |
|---------------------------------------|
| |
| |
| ive charge experienced by an electron |
| |
| nds |

| | |
|---------------------------------|---|
| | |
| the average amount of electrons | ◆ |
| X | |
| 0 | |
| 0,00 | |

| | |
|--|----------|
| | Score (p |
| net positive charge experienced by an electron | 1397 |
| net positive charge experienced by an electron | 992 |
| net positive charge experienced by an electron | 962 |
| net positive charge experienced by an electron | 962 |
| net positive charge experienced by an electron | 1398 |
| net positive charge experienced by an electron | 1185 |
| | 0 |
| net positive charge experienced by an electron | 1325 |
| net positive charge experienced by an electron | 1160 |
| net positive charge experienced by an electron | 978 |

7 Quiz

| | |
|--|------|
| net positive charge experienced by an electron | 1108 |
| net positive charge experienced by an electron | 1363 |
| net positive charge experienced by an electron | 1092 |
| net positive charge experienced by an electron | 1400 |
| net positive charge experienced by an electron | 1360 |
| net positive charge experienced by an electron | 882 |
| net positive charge experienced by an electron | 1260 |
| net positive charge experienced by an electron | 1483 |
| net positive charge experienced by an electron | 1113 |
| net positive charge experienced by an electron | 1063 |
| net negative charge of the electron | 0 |
| net positive charge experienced by an electron | 1157 |
| net positive charge experienced by an electron | 1488 |
| net positive charge experienced by an electron | 640 |
| net negative charge of the electron | 0 |
| net positive charge experienced by an electron | 1393 |
| net positive charge experienced by an electron | 1162 |
| net positive charge experienced by an electron | 1412 |
| net positive charge experienced by an electron | 1165 |
| net positive charge experienced by an electron | 775 |

7 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|-------------------------------------|-------------|
| | |
| net negative charge of the electron | <div></div> |
| X | |
| 2 | |
| 13,95 | |

| | |
|--------|---------|
| | |
| oints) | Current |
| | 7908 |
| | 5590 |
| | 4525 |
| | 4721 |
| | 7846 |
| | 4637 |
| | 0 |
| | 8278 |
| | 7427 |
| | 3526 |

7 Quiz

| | |
|--|------|
| | 4389 |
| | 7412 |
| | 4849 |
| | 7927 |
| | 7847 |
| | 2949 |
| | 5916 |
| | 8929 |
| | 5060 |
| | 5479 |
| | 5064 |
| | 5729 |
| | 8923 |
| | 4181 |
| | 6260 |
| | 7920 |
| | 5254 |
| | 8128 |
| | 5286 |
| | 4133 |

7 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--|------------------------|
| | |
| net positive charge experienced by an electron | <div><div></div></div> |
| <div><div>✓</div></div> | |
| 27 | |
| 10,72 | |

| | |
|----------------------|-----------|
| | |
| Total Score (points) | Answer ti |
| | 6,2 |
| | 18,5 |
| | 8,3 |
| | 20,3 |
| | 6,1 |
| | 6,9 |
| | 30 |
| | 10,5 |
| | 20,4 |
| | 19,3 |

7 Quiz

| | |
|--|------|
| | 11,5 |
| | 8,2 |
| | 12,5 |
| | 6 |
| | 8,4 |
| | 13,1 |
| | 2,4 |
| | 1 |
| | 11,2 |
| | 14,2 |
| | 15,5 |
| | 8,6 |
| | 0,7 |
| | 27,6 |
| | 12,4 |
| | 6,4 |
| | 8,3 |
| | 5,3 |
| | 8,1 |
| | 19,5 |

| Shielding |
|-------------------|
| 8 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

8 Quiz

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|----------------------------------|------------|
| What is the net positive charge? | |
| is | how many |
| (%) | 86,67% |
| on | 30 seconds |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |



| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |

8 Quiz

| | |
|---|---|
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
|) | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |

8 Quiz

| |
|--|
| |
| |
| how many available protons are acting on valence electrons |
| |
| points |

| | |
|---|---|
| | |
| how many available protons are acting on valence electrons |  |
|  | |
| 26 | |
| 13,72 | |

| | |
|--|----------------|
| | Score (points) |
| how many available protons are acting on valence electrons | 1335 |
| the average number of protons | 0 |
| how many available protons are acting on valence electrons | 780 |
| how many available protons are acting on valence electrons | 1087 |
| how many available protons are acting on valence electrons | 1253 |
| how many available protons are acting on valence electrons | 1258 |
| | 0 |
| how many available protons are acting on valence electrons | 1398 |
| how many available protons are acting on valence electrons | 1120 |
| how many available protons are acting on valence electrons | 917 |

8 Quiz

| | |
|--|------|
| how many availbe protons are acting on valence electrons | 1130 |
| how many availbe protons are acting on valence electrons | 1110 |
| how many periods are in the periodic table | 0 |
| how many availbe protons are acting on valence electrons | 1343 |
| how many availbe protons are acting on valence electrons | 1340 |
| how many availbe protons are acting on valence electrons | 920 |
| how many availbe protons are acting on valence electrons | 1307 |
| how many availbe protons are acting on valence electrons | 1492 |
| how many availbe protons are acting on valence electrons | 1248 |
| how many availbe protons are acting on valence electrons | 1162 |
| how many availbe protons are acting on valence electrons | 783 |
| how many availbe protons are acting on valence electrons | 1218 |
| how many availbe protons are acting on valence electrons | 1500 |
| how many availbe protons are acting on valence electrons | 793 |
| how many electrons and in orbitals | 0 |
| how many availbe protons are acting on valence electrons | 1292 |
| how many availbe protons are acting on valence electrons | 1175 |
| how many availbe protons are acting on valence electrons | 1362 |
| how many availbe protons are acting on valence electrons | 1137 |
| how many availbe protons are acting on valence electrons | 902 |

8 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|--|------------------------|
| | |
| how many periods are in the periodic table | <div><div></div></div> |
| X | |
| 1 | |
| 29,10 | |

| | |
|--------|---------|
| | |
| oints) | Current |
| | 9243 |
| | 5590 |
| | 5305 |
| | 5808 |
| | 9099 |
| | 5895 |
| | 0 |
| | 9676 |
| | 8547 |
| | 4443 |

8 Quiz

| | |
|--|-------|
| | 5519 |
| | 8522 |
| | 4849 |
| | 9270 |
| | 9187 |
| | 3869 |
| | 7223 |
| | 10421 |
| | 6308 |
| | 6641 |
| | 5847 |
| | 6947 |
| | 10423 |
| | 4974 |
| | 6260 |
| | 9212 |
| | 6429 |
| | 9490 |
| | 6423 |
| | 5035 |

8 Quiz

| |
|--|
| |
| |
| |
| |
| |

| | |
|------------------------------------|--------------------------|
| | |
| how many electrons and in orbitals | <input type="checkbox"/> |
| X | |
| 1 | |
| 26,20 | |

| | |
|----------------------|-----------|
| | |
| Total Score (points) | Answer ti |
| | 9,9 |
| | 10 |
| | 25,2 |
| | 18,8 |
| | 14,8 |
| | 8,5 |
| | 30 |
| | 6,1 |
| | 22,8 |
| | 29 |

8 Quiz

| | |
|--|------|
| | 16,2 |
| | 23,4 |
| | 29,1 |
| | 9,4 |
| | 9,6 |
| | 16,8 |
| | 5,6 |
| | 0,5 |
| | 9,1 |
| | 14,3 |
| | 13 |
| | 10,9 |
| | 0,4 |
| | 24,4 |
| | 26,2 |
| | 12,5 |
| | 13,5 |
| | 8,3 |
| | 15,8 |
| | 17,9 |

8 Quiz

The graph displays the average number of protons (N) on the vertical axis against time in seconds on the horizontal axis. The vertical axis ranges from 0 to 10,000 with major grid lines every 2,000 units. The horizontal axis ranges from 0 to 10 with major grid lines every 2 units. A series of red 'x' markers connected by a line shows a decreasing trend. The data points are approximately as follows:

| time (seconds) | the average number of protons (N) |
|----------------|---------------------------------------|
| 0 | 10,000 |
| 1 | 8,000 |
| 2 | 6,500 |
| 3 | 5,500 |
| 4 | 4,800 |
| 5 | 4,200 |
| 6 | 3,800 |
| 7 | 3,400 |
| 8 | 3,100 |
| 9 | 2,800 |
| 10 | 2,500 |

8 Quiz

[illegible]

| Shielding |
|-------------------|
| 9 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
| |
| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
| Average time tal |
| |
| Answer Deta |
| Players |
| ((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

| |
|-----------------|
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|---|----------|
| Which group has the smallest value for the shielding trend? | |
| s | Period 1 |
| (%) | 93,33% |
| on | 30 secor |

| Summary | |
|-------------------------|---|
| | ▲ |
| st? | |
| ers received | |
| ken to answer (seconds) | |

| ails | |
|------|--------|
| | Answer |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✗ |
| | ✓ |
| | ✓ |
| | ✓ |

9 Quiz

| | |
|---|---|
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
|) | ✓ |
| | ✓ |
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| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |
| | ✓ |

| |
|---------------------------------------|
| |
| |
| , it has the least amount of orbitals |
| |
| nds |

| | |
|--|---|
| | |
| Period 4, because they have the most protons | ◆ |
| X | |
| 0 | |
| 0,00 | |

| | |
|---|----------|
| | Score (p |
| Period 1, it has the least amount of orbitals | 1450 |
| Period 1, it has the least amount of orbitals | 912 |
| | 0 |
| Period 1, it has the least amount of orbitals | 1457 |
| Period 1, it has the least amount of orbitals | 1468 |
| Period 1, it has the least amount of orbitals | 1222 |
| | 0 |
| Period 1, it has the least amount of orbitals | 1365 |
| Period 1, it has the least amount of orbitals | 1447 |
| Period 1, it has the least amount of orbitals | 1448 |

9 Quiz

| | |
|---|------|
| Period 1, it has the least amount of orbitals | 1142 |
| Period 1, it has the least amount of orbitals | 1435 |
| Period 1, it has the least amount of orbitals | 952 |
| Period 1, it has the least amount of orbitals | 1452 |
| Period 1, it has the least amount of orbitals | 1393 |
| Period 1, it has the least amount of orbitals | 1242 |
| Period 1, it has the least amount of orbitals | 1465 |
| Period 1, it has the least amount of orbitals | 1500 |
| Period 1, it has the least amount of orbitals | 1417 |
| Period 1, it has the least amount of orbitals | 1462 |
| Period 1, it has the least amount of orbitals | 1042 |
| Period 1, it has the least amount of orbitals | 1420 |
| Period 1, it has the least amount of orbitals | 1500 |
| Period 1, it has the least amount of orbitals | 1142 |
| Period 1, it has the least amount of orbitals | 938 |
| Period 1, it has the least amount of orbitals | 1463 |
| Period 1, it has the least amount of orbitals | 1463 |
| Period 1, it has the least amount of orbitals | 1470 |
| Period 1, it has the least amount of orbitals | 1455 |
| Period 1, it has the least amount of orbitals | 1002 |

9 Quiz

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| | |
| Period 1, it has the least amount of orbitals | <div><div></div></div> |
| ✓ | |
| 28 | |
| 5,12 | |

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| | |
| oints) | Current |
| | 10693 |
| | 6502 |
| | 5305 |
| | 7265 |
| | 10567 |
| | 7117 |
| | 0 |
| | 11041 |
| | 9994 |
| | 5891 |

9 Quiz

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| | 6661 |
| | 9957 |
| | 5801 |
| | 10722 |
| | 10580 |
| | 5111 |
| | 8688 |
| | 11921 |
| | 7725 |
| | 8103 |
| | 6889 |
| | 8367 |
| | 11923 |
| | 6116 |
| | 7198 |
| | 10675 |
| | 7892 |
| | 10960 |
| | 7878 |
| | 6037 |

9 Quiz

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| | |
| Period 6, has the most mass | <input type="checkbox"/> |
| X | |
| 0 | |
| 0,00 | |

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| | |
| Total Score (points) | Answer ti |
| | 3 |
| | 5,3 |
| | 30 |
| | 2,6 |
| | 1,9 |
| | 16,7 |
| | 30 |
| | 8,1 |
| | 3,2 |
| | 3,1 |

9 Quiz

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| | 21,5 |
| | 3,9 |
| | 2,9 |
| | 2,9 |
| | 6,4 |
| | 3,5 |
| | 2,1 |
| | 0,4 |
| | 5 |
| | 2,3 |
| | 3,5 |
| | 4,8 |
| | 0,3 |
| | 9,5 |
| | 3,7 |
| | 2,2 |
| | 2,2 |
| | 1,8 |
| | 2,7 |
| | 17,9 |

Period 3, it has the most mass

X

0

0,00

Time (seconds)

| Shielding |
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| 10 Quiz |
| Correct answers |
| Players correct (|
| Question duratic |
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| Answer Sum |
| Answer options |
| Is answer correc |
| Number of answ |
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| Answer Deta |
| Players |
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| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |

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| Sriveena |
| Timothy |
| Wale (lil duub) |
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| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Effect | |
|---|----------|
| Shielding effect affects the neutrons and protons | |
| s | False |
| (%) | 80,00% |
| on | 20 secor |

| Summary | |
|-------------------------|---|
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| st? | |
| ers received | |
| ken to answer (seconds) | |




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| | Answer |
| | ✓ |
| | X |
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10 Quiz

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| False |  |
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| 24 | |
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| | Score (p |
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| False | 1420 |
| True | 0 |
| | 0 |
| False | 1463 |
| False | 1330 |
| False | 1428 |
| | 0 |
| False | 1385 |
| False | 1443 |
| True | 0 |

10 Quiz

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| False | 1395 |
| False | 1345 |
| True | 0 |
| False | 1425 |
| False | 1298 |
| False | 1125 |
| False | 1460 |
| False | 1485 |
| False | 1415 |
| False | 1340 |
| True | 0 |
| False | 1293 |
| False | 1488 |
| False | 1140 |
| False | 838 |
| False | 1168 |
| False | 1150 |
| False | 1448 |
| False | 1368 |
| False | 1265 |

10 Quiz

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| | |
| True | <input checked="" type="radio"/> |
| X | |
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| | |
| oints) | Current |
| | 12113 |
| | 6502 |
| | 5305 |
| | 8728 |
| | 11897 |
| | 8545 |
| | 0 |
| | 12426 |
| | 11437 |
| | 5891 |

10 Quiz

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| | 8056 |
| | 11302 |
| | 5801 |
| | 12147 |
| | 11878 |
| | 6236 |
| | 10148 |
| | 13406 |
| | 9140 |
| | 9443 |
| | 6889 |
| | 9660 |
| | 13411 |
| | 7256 |
| | 8036 |
| | 11843 |
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| | 7302 |

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| | 3 |
| | 20 |
| | 1,5 |
| | 6,8 |
| | 2,9 |
| | 20 |
| | 4,6 |
| | 2,3 |
| | 18,5 |

10 Quiz

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| | 4,2 |
| | 6,2 |
| | 15 |
| | 3 |
| | 8,1 |
| | 11 |
| | 1,6 |
| | 0,6 |
| | 3,4 |
| | 6,4 |
| | 1,2 |
| | 8,3 |
| | 0,5 |
| | 10,4 |
| | 10,5 |
| | 13,3 |
| | 14 |
| | 2,1 |
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| What happens to the shielding trend moving from left to right across a period on the table |

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| Which period has the largest value for the shielding trend? |
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| What is the net positive charge? |
| Which group has the smallest value for the shielding trend? |
| Which group has the smallest value for the shielding trend? |
| Which group has the smallest value for the shielding trend? |
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| Which group has the smallest value for the shielding trend? |
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| Which group has the smallest value for the shielding trend? |
| Shielding effect affects the neutrons and protons |

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| Shielding effect affects the neutrons and protons |
| Shielding effect affects the neutrons and protons |
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| Shielding effect affects the neutrons and protons |
| Shielding effect affects the neutrons and protons |
| Shielding effect affects the neutrons and protons |

| Answer 1 | Answer 2 |
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| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |

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| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| Interference from inner core electrons | when electrons dont fit on the orbital |
| number of inner orbital increase | number of valence orbitals increases |

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| number of inner orbital increase | number of valence orbitals increases |
| number of inner orbital increase | number of valence orbitals increases |
| number of inner orbital increase | number of valence orbitals increases |
| number of inner orbital increase | number of valence orbitals increases |
| number of inner orbital increase | number of valence orbitals increases |
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| number of inner orbital increase | number of valence orbitals increases |
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| number of inner orbital increase | number of valence orbitals increases |
| number of inner orbital increase | number of valence orbitals increases |
| number of inner orbital increase | number of valence orbitals increases |
| number of inner orbital increase | number of valence orbitals increases |
| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |

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| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |
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| The amount of neutrons decreases | The mass increases |
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| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |
| The amount of neutrons decreases | The mass increases |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |

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| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
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| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |

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| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| smaller electrons fall off the orbital | inner core electrons block attractive force |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |

| | |
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| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
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| the element changes shape and color | causes valence electrons to fall of and get lost |
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| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |

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| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| the element changes shape and color | causes valence electrons to fall of and get lost |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |

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| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
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| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |

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| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| period 7 bc they have the most orbitals | period 3 bc they have the most protons |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |

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| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
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| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |

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| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| the average amount of electrons | net negative charge of the electron |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |

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| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |

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| how many availbe protons are acting on valence electrons | how many periods are in the periodic table |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |

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| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| Period 4, because they have the most protons | Period 1, it has the least amount of orbitals |
| False | True |

RawReportData Data

| | |
|-------|------|
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
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| False | True |
| False | True |
| False | True |

RawReportData Data

| | |
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| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |
| False | True |

| Answer 3 | Answer 4 |
|---|------------------------------------|
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |

| | |
|---|--------------------------------------|
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| how the negative charge affects the protons | the ratio of electrons to neutrons |
| number of inner orbitals stays the same | number of valence orbitals decreases |

| | |
|---|--------------------------------------|
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
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| number of inner orbitals stays the same | number of valence orbitals decreases |

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| number of inner orbitals stays the same | number of valence orbitals decreases |
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| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| number of inner orbitals stays the same | number of valence orbitals decreases |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |

| | |
|----------------------------------|---|
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |

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| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| The number of orbitals increases | shielding increases because the atom has more inner core orbitals |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |

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| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |

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| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| protons are not related to valence electrons | protons are weaker than valance electrons |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |

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| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
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| the protons increase | causes electrons to disintegrate |

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| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| the protons increase | causes electrons to disintegrate |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |

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| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |

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| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| period 5 bc they have the most electrons | period 6 bc the have the most neutrons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |

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| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |

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| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| net positive charge experienced by an electron | average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |

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| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
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| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |
| how many electrons and in orbitals | the average number of protons |

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| how many electrons and in orbitals | the average number of protons |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |

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| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
| Period 6, has the most mass | Period 3, it has the most mass |
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| Correct Answers | Time Allotted to Answer (seconds) |
|--|-----------------------------------|
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |

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|---|----|
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| Interference from inner core electrons | 30 |
| number of inner orbitals stays the same | 30 |

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|---|----|
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
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| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |

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| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| number of inner orbitals stays the same | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |

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| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
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| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
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| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |

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| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| shielding increases because the atom has more inner core orbitals | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |

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| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
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| inner core electrons block attractive force | 30 |
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| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |

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| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| inner core electrons block attractive force | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |

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| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
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| causes valence electrons to fall of and get lost | 30 |
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| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| causes valence electrons to fall of and get lost | 30 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |

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| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
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| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |

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| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| period 7 bc they have the most orbitals | 20 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |

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| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
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| net positive charge experienced by an electron | 30 |
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| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |

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| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| net positive charge experienced by an electron | 30 |
| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |
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| how many availbe protons are acting on valence electrons | 30 |
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| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |
| how many availbe protons are acting on valence electrons | 30 |

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| how many availbe protons are acting on valence electrons | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |

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| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
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| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| Period 1, it has the least amount of orbitals | 30 |
| False | 20 |

RawReportData Data

| | |
|-------|----|
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |

RawReportData Data

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| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |
| False | 20 |

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|-----------------|
| Players |
| ((Jajuan)) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |

| |
|----------------|
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |
| (((Jajuan))) |

| |
|-----------------|
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |

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| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |
| (((Jajuan))) |
| Andrew F |
| Arnav(>^<) |

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|-----------------|
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |

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| duncan |
| hunter |
| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
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| ?? ronojoy ?? |
| ((Jajuan)) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |

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| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |
| Shrey |
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| Will Paasch |
| alex demchenko |
| ashley |
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| duncan |
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| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
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| (((Jajuan))) |
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| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |

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| Sam S |
| Shrey |
| Sriveena |
| Timothy |
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| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
| julia |
| maggie |

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| matthew |
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| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |
| ((Jajuan))) |
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| N8 Baker |
| Reba |

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| Will Paasch |
| alex demchenko |
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| duncan |
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| julia |
| maggie |
| matthew |
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| sophia |
| zoe |
| ?? ronojoy ?? |
| ((Jajuan))) |
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| Dhanshree |
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| Will Paasch |
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| Sriveena |
| Timothy |

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| alex demchenko |
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| conner parker |
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| maggie |
| matthew |
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| sophia |
| zoe |

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| ?? ronojoy ?? |
| (((Jajuan))) |
| Andrew F |
| Arnav(>^<) |
| Dhanshree |
| Gayatri |
| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |

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| Will Paasch |
| alex demchenko |
| ashley |
| conner parker |
| duncan |
| hunter |
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| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |
| (((Jajuan))) |

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| Arnav(>^<) |
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| Luke |
| Mo The Pro |
| N8 Baker |
| Reba |
| Sam S |
| Shrey |
| Sriveena |
| Timothy |
| Wale (lil duub) |
| Will |
| Will Paasch |
| alex demchenko |

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| ashley |
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| julia |
| maggie |
| matthew |
| nya |
| pravleen |
| riya |
| sophia |
| zoe |
| ?? ronojoy ?? |

| Answer | Correct / Incorrect | Correct |
|--|---------------------|---------|
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| the ratio of electrons to neutrons | Incorrect | 0 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| the ratio of electrons to neutrons | Incorrect | 0 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |

| | | |
|---|-----------|---|
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| how the negative charge affects the protons | Incorrect | 0 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| Interference from inner core electrons | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |

RawReportData Data

| | | |
|---|-----------|---|
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbital increase | Incorrect | 0 |
| number of inner orbitals stays the same | Correct | 1 |
| number of valence orbitals decreases | Incorrect | 0 |
| | Incorrect | 0 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbital increase | Incorrect | 0 |
| number of valence orbitals decreases | Incorrect | 0 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbital increase | Incorrect | 0 |
| number of inner orbitals stays the same | Correct | 1 |

| | | |
|---|-----------|---|
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| number of inner orbitals stays the same | Correct | 1 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |

| | | |
|---|-----------|---|
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |
| | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The amount of neutrons decreases | Incorrect | 0 |
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |

| | | |
|---|-----------|---|
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |
| shielding increases because the atom has more inner core orbitals | Correct | 1 |
| The number of orbitals increases | Incorrect | 0 |
| The number of orbitals increases | Incorrect | 0 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| | Incorrect | 0 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |

| | | |
|---|-----------|---|
| inner core electrons block attractive force | Correct | 1 |
| | Incorrect | 0 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |

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|--|-----------|---|
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| inner core electrons block attractive force | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| | Incorrect | 0 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| | Incorrect | 0 |

RawReportData Data

| | | |
|--|-----------|---|
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes electrons to disintegrate | Incorrect | 0 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| the protons increase | Incorrect | 0 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |

| | | |
|--|-----------|---|
| the protons increase | Incorrect | 0 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| causes valence electrons to fall of and get lost | Correct | 1 |
| the protons increase | Incorrect | 0 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| | Incorrect | 0 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |

| | | |
|---|---------|---|
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |

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|--|-----------|---|
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| period 7 bc they have the most orbitals | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| | Incorrect | 0 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |

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|--|-----------|---|
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net negative charge of the electron | Incorrect | 0 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net negative charge of the electron | Incorrect | 0 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |

| | | |
|--|-----------|---|
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| net positive charge experienced by an electron | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| the average number of protons | Incorrect | 0 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| | Incorrect | 0 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many periods are in the periodic table | Incorrect | 0 |

| | | |
|--|-----------|---|
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many electrons and in orbitals | Incorrect | 0 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |
| how many availbe protons are acting on valence electrons | Correct | 1 |

| | | |
|--|-----------|---|
| how many availbe protons are acting on valence electrons | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| | Incorrect | 0 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| | Incorrect | 0 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |

| | | |
|---|---------|---|
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| Period 1, it has the least amount of orbitals | Correct | 1 |
| False | Correct | 1 |

RawReportData Data

| | | |
|-------|-----------|---|
| True | Incorrect | 0 |
| | Incorrect | 0 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| | Incorrect | 0 |
| False | Correct | 1 |
| False | Correct | 1 |
| True | Incorrect | 0 |
| False | Correct | 1 |
| False | Correct | 1 |
| True | Incorrect | 0 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |

RawReportData Data

| | | |
|-------|-----------|---|
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| True | Incorrect | 0 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |
| False | Correct | 1 |

RawReportData Data

| Incorrect | Score (points) | Score without Answer Streak Bonus (points) |
|-----------|----------------|--|
| 0 | 917 | 917 |
| 0 | 850 | 850 |
| 0 | 877 | 877 |
| 0 | 947 | 947 |
| 0 | 765 | 765 |
| 0 | 675 | 675 |
| 1 | 0 | 0 |
| 0 | 922 | 922 |
| 0 | 732 | 732 |
| 1 | 0 | 0 |
| 0 | 683 | 683 |
| 0 | 777 | 777 |
| 0 | 555 | 555 |
| 0 | 932 | 932 |
| 0 | 885 | 885 |

RawReportData Data

| | | |
|---|-----|-----|
| 0 | 537 | 537 |
| 0 | 868 | 868 |
| 0 | 987 | 987 |
| 0 | 655 | 655 |
| 0 | 803 | 803 |
| 0 | 930 | 930 |
| 0 | 785 | 785 |
| 0 | 983 | 983 |
| 1 | 0 | 0 |
| 0 | 633 | 633 |
| 0 | 673 | 673 |
| 0 | 682 | 682 |
| 0 | 953 | 953 |
| 0 | 670 | 670 |
| 0 | 855 | 855 |
| 0 | 853 | 753 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 905 | 805 |
| 0 | 875 | 775 |
| 1 | 0 | 0 |
| 0 | 963 | 863 |
| 1 | 0 | 0 |
| 1 | 0 | 0 |
| 0 | 990 | 890 |
| 0 | 945 | 845 |
| 1 | 0 | 0 |
| 1 | 0 | 0 |
| 0 | 790 | 690 |
| 0 | 695 | 595 |
| 0 | 840 | 740 |
| 0 | 882 | 782 |
| 1 | 0 | 0 |
| 0 | 1027 | 927 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 1082 | 982 |
| 0 | 648 | 548 |
| 0 | 740 | 640 |
| 0 | 932 | 832 |
| 0 | 998 | 898 |
| 0 | 1087 | 987 |
| 0 | 738 | 738 |
| 0 | 762 | 662 |
| 0 | 1007 | 907 |
| 0 | 670 | 570 |
| 0 | 898 | 798 |
| 0 | 718 | 618 |
| 0 | 808 | 708 |
| 0 | 900 | 700 |
| 1 | 0 | 0 |
| 0 | 918 | 718 |

RawReportData Data

| | | |
|---|------|-----|
| 1 | 0 | 0 |
| 0 | 927 | 727 |
| 1 | 0 | 0 |
| 1 | 0 | 0 |
| 0 | 1025 | 825 |
| 0 | 1018 | 818 |
| 1 | 0 | 0 |
| 1 | 0 | 0 |
| 0 | 1027 | 827 |
| 1 | 0 | 0 |
| 0 | 917 | 717 |
| 0 | 992 | 792 |
| 1 | 0 | 0 |
| 1 | 0 | 0 |
| 0 | 1187 | 987 |
| 1 | 0 | 0 |

RawReportData Data

| | | |
|---|------|-----|
| 1 | 0 | 0 |
| 0 | 1142 | 942 |
| 1 | 0 | 0 |
| 0 | 1190 | 990 |
| 0 | 1007 | 907 |
| 0 | 1100 | 900 |
| 0 | 1078 | 878 |
| 1 | 0 | 0 |
| 0 | 948 | 748 |
| 1 | 0 | 0 |
| 1 | 0 | 0 |
| 0 | 1098 | 798 |
| 0 | 828 | 828 |
| 1 | 0 | 0 |
| 0 | 797 | 797 |
| 0 | 1090 | 790 |

RawReportData Data

| | | |
|---|------|------|
| 0 | 647 | 647 |
| 1 | 0 | 0 |
| 0 | 1260 | 960 |
| 0 | 1052 | 752 |
| 0 | 692 | 692 |
| 0 | 660 | 660 |
| 0 | 1012 | 712 |
| 0 | 517 | 517 |
| 0 | 1105 | 805 |
| 0 | 1088 | 788 |
| 0 | 585 | 585 |
| 0 | 730 | 730 |
| 0 | 1300 | 1000 |
| 0 | 623 | 623 |
| 0 | 825 | 825 |
| 0 | 1150 | 850 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 722 | 722 |
| 0 | 1290 | 990 |
| 0 | 988 | 788 |
| 0 | 1160 | 860 |
| 0 | 1153 | 853 |
| 0 | 780 | 780 |
| 0 | 1130 | 830 |
| 0 | 795 | 795 |
| 0 | 787 | 787 |
| 0 | 1335 | 935 |
| 0 | 887 | 787 |
| 1 | 0 | 0 |
| 0 | 910 | 810 |
| 0 | 1260 | 860 |
| 0 | 1062 | 962 |
| 1 | 0 | 0 |

RawReportData Data

| | | |
|---|------|------|
| 0 | 1323 | 923 |
| 0 | 1182 | 782 |
| 0 | 873 | 773 |
| 0 | 885 | 785 |
| 0 | 1058 | 658 |
| 0 | 877 | 777 |
| 0 | 1333 | 933 |
| 0 | 1205 | 805 |
| 1 | 0 | 0 |
| 0 | 948 | 848 |
| 0 | 1390 | 990 |
| 0 | 933 | 833 |
| 0 | 908 | 808 |
| 1 | 0 | 0 |
| 0 | 947 | 847 |
| 0 | 1400 | 1000 |

RawReportData Data

| | | |
|---|------|-----|
| 1 | 0 | 0 |
| 0 | 1190 | 790 |
| 0 | 1193 | 793 |
| 0 | 835 | 735 |
| 0 | 1352 | 952 |
| 0 | 865 | 765 |
| 1 | 0 | 0 |
| 0 | 1408 | 908 |
| 0 | 1128 | 928 |
| 0 | 893 | 893 |
| 0 | 1105 | 905 |
| 0 | 1443 | 943 |
| 0 | 1068 | 868 |
| 1 | 0 | 0 |
| 0 | 1433 | 933 |
| 0 | 1338 | 838 |

RawReportData Data

| | | |
|---|------|------|
| 0 | 983 | 783 |
| 0 | 1053 | 853 |
| 0 | 1385 | 885 |
| 0 | 1113 | 913 |
| 0 | 1400 | 900 |
| 0 | 1435 | 935 |
| 0 | 945 | 945 |
| 0 | 1083 | 883 |
| 0 | 1500 | 1000 |
| 0 | 1088 | 888 |
| 0 | 1140 | 940 |
| 0 | 910 | 910 |
| 0 | 1120 | 920 |
| 0 | 1485 | 985 |
| 0 | 808 | 808 |
| 0 | 1415 | 915 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 1423 | 923 |
| 0 | 1125 | 925 |
| 0 | 1435 | 935 |
| 0 | 1073 | 873 |
| 0 | 908 | 908 |
| 0 | 1397 | 897 |
| 0 | 992 | 692 |
| 0 | 962 | 862 |
| 0 | 962 | 662 |
| 0 | 1398 | 898 |
| 0 | 1185 | 885 |
| 1 | 0 | 0 |
| 0 | 1325 | 825 |
| 0 | 1160 | 660 |
| 0 | 978 | 678 |
| 0 | 1108 | 808 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 1363 | 863 |
| 0 | 1092 | 792 |
| 0 | 1400 | 900 |
| 0 | 1360 | 860 |
| 0 | 882 | 782 |
| 0 | 1260 | 960 |
| 0 | 1483 | 983 |
| 0 | 1113 | 813 |
| 0 | 1063 | 763 |
| 1 | 0 | 0 |
| 0 | 1157 | 857 |
| 0 | 1488 | 988 |
| 0 | 640 | 540 |
| 1 | 0 | 0 |
| 0 | 1393 | 893 |
| 0 | 1162 | 862 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 1412 | 912 |
| 0 | 1165 | 865 |
| 0 | 775 | 675 |
| 0 | 1335 | 835 |
| 1 | 0 | 0 |
| 0 | 780 | 580 |
| 0 | 1087 | 687 |
| 0 | 1253 | 753 |
| 0 | 1258 | 858 |
| 1 | 0 | 0 |
| 0 | 1398 | 898 |
| 0 | 1120 | 620 |
| 0 | 917 | 517 |
| 0 | 1130 | 730 |
| 0 | 1110 | 610 |
| 1 | 0 | 0 |

RawReportData Data

| | | |
|---|------|------|
| 0 | 1343 | 843 |
| 0 | 1340 | 840 |
| 0 | 920 | 720 |
| 0 | 1307 | 907 |
| 0 | 1492 | 992 |
| 0 | 1248 | 848 |
| 0 | 1162 | 762 |
| 0 | 783 | 783 |
| 0 | 1218 | 818 |
| 0 | 1500 | 1000 |
| 0 | 793 | 593 |
| 1 | 0 | 0 |
| 0 | 1292 | 792 |
| 0 | 1175 | 775 |
| 0 | 1362 | 862 |
| 0 | 1137 | 737 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 902 | 702 |
| 0 | 1450 | 950 |
| 0 | 912 | 912 |
| 1 | 0 | 0 |
| 0 | 1457 | 957 |
| 0 | 1468 | 968 |
| 0 | 1222 | 722 |
| 1 | 0 | 0 |
| 0 | 1365 | 865 |
| 0 | 1447 | 947 |
| 0 | 1448 | 948 |
| 0 | 1142 | 642 |
| 0 | 1435 | 935 |
| 0 | 952 | 952 |
| 0 | 1452 | 952 |
| 0 | 1393 | 893 |

RawReportData Data

| | | |
|---|------|------|
| 0 | 1242 | 942 |
| 0 | 1465 | 965 |
| 0 | 1500 | 1000 |
| 0 | 1417 | 917 |
| 0 | 1462 | 962 |
| 0 | 1042 | 942 |
| 0 | 1420 | 920 |
| 0 | 1500 | 1000 |
| 0 | 1142 | 842 |
| 0 | 938 | 938 |
| 0 | 1463 | 963 |
| 0 | 1463 | 963 |
| 0 | 1470 | 970 |
| 0 | 1455 | 955 |
| 0 | 1002 | 702 |
| 0 | 1420 | 920 |

RawReportData Data

| | | |
|---|------|-----|
| 1 | 0 | 0 |
| 1 | 0 | 0 |
| 0 | 1463 | 963 |
| 0 | 1330 | 830 |
| 0 | 1428 | 928 |
| 1 | 0 | 0 |
| 0 | 1385 | 885 |
| 0 | 1443 | 943 |
| 1 | 0 | 0 |
| 0 | 1395 | 895 |
| 0 | 1345 | 845 |
| 1 | 0 | 0 |
| 0 | 1425 | 925 |
| 0 | 1298 | 798 |
| 0 | 1125 | 725 |
| 0 | 1460 | 960 |

RawReportData Data

| | | |
|---|------|-----|
| 0 | 1485 | 985 |
| 0 | 1415 | 915 |
| 0 | 1340 | 840 |
| 1 | 0 | 0 |
| 0 | 1293 | 793 |
| 0 | 1488 | 988 |
| 0 | 1140 | 740 |
| 0 | 838 | 738 |
| 0 | 1168 | 668 |
| 0 | 1150 | 650 |
| 0 | 1448 | 948 |
| 0 | 1368 | 868 |
| 0 | 1265 | 865 |

RawReportData Data

| Current Total Score (points) | Answer Time (%) |
|------------------------------|-----------------|
| 917 | 16.67% |
| 850 | 30.00% |
| 877 | 24.67% |
| 947 | 10.67% |
| 765 | 47.00% |
| 675 | 65.00% |
| 0 | 91.67% |
| 922 | 15.67% |
| 732 | 53.67% |
| 0 | 89.33% |
| 683 | 63.33% |
| 777 | 44.67% |
| 555 | 89.00% |
| 932 | 13.67% |
| 885 | 23.00% |

RawReportData Data

| | |
|------|--------|
| 537 | 92.67% |
| 868 | 26.33% |
| 987 | 2.67% |
| 655 | 69.00% |
| 803 | 39.33% |
| 930 | 14.00% |
| 785 | 43.00% |
| 983 | 3.33% |
| 0 | 18.67% |
| 633 | 73.33% |
| 673 | 65.33% |
| 682 | 63.67% |
| 953 | 9.33% |
| 670 | 66.00% |
| 855 | 29.00% |
| 1770 | 49.33% |

RawReportData Data

| | |
|------|---------|
| 1755 | 39.00% |
| 1752 | 45.00% |
| 947 | 18.33% |
| 1728 | 27.33% |
| 675 | 95.33% |
| 0 | 100.00% |
| 1912 | 22.00% |
| 1677 | 31.00% |
| 0 | 44.67% |
| 683 | 73.67% |
| 1567 | 62.00% |
| 1250 | 81.00% |
| 1772 | 52.00% |
| 1767 | 43.67% |
| 537 | 42.67% |
| 1895 | 14.67% |

RawReportData Data

| | |
|------|--------|
| 2069 | 3.67% |
| 1303 | 90.33% |
| 1543 | 72.00% |
| 1862 | 33.67% |
| 1783 | 20.33% |
| 2070 | 2.67% |
| 738 | 52.33% |
| 1395 | 67.67% |
| 1680 | 18.67% |
| 1352 | 86.00% |
| 1851 | 40.33% |
| 1388 | 76.33% |
| 1663 | 58.33% |
| 2670 | 60.00% |
| 1755 | 28.67% |
| 2670 | 56.33% |

RawReportData Data

| | |
|------|---------|
| 947 | 19.00% |
| 2655 | 54.67% |
| 675 | 73.33% |
| 0 | 100.00% |
| 2937 | 35.00% |
| 2695 | 36.33% |
| 0 | 96.00% |
| 683 | 28.00% |
| 2594 | 34.67% |
| 1250 | 89.67% |
| 2689 | 56.67% |
| 2759 | 41.67% |
| 537 | 88.00% |
| 1895 | 20.67% |
| 3256 | 2.67% |
| 1303 | 78.67% |

RawReportData Data

| | |
|------|---------|
| 1543 | 18.33% |
| 3004 | 11.67% |
| 1783 | 23.33% |
| 3260 | 2.00% |
| 1745 | 18.67% |
| 2495 | 20.00% |
| 2758 | 24.33% |
| 1352 | 17.00% |
| 2799 | 50.33% |
| 1388 | 20.33% |
| 1663 | 23.67% |
| 3768 | 40.33% |
| 2583 | 34.33% |
| 2670 | 100.00% |
| 1744 | 40.67% |
| 3745 | 42.00% |

RawReportData Data

| | |
|------|---------|
| 1322 | 70.67% |
| 0 | 100.00% |
| 4197 | 8.00% |
| 3747 | 49.67% |
| 692 | 61.67% |
| 1343 | 68.00% |
| 3606 | 57.67% |
| 1767 | 96.67% |
| 3794 | 39.00% |
| 3847 | 42.33% |
| 1122 | 83.00% |
| 2625 | 54.00% |
| 4556 | 1.33% |
| 1926 | 75.33% |
| 2368 | 35.00% |
| 4154 | 30.00% |

RawReportData Data

| | |
|------|---------|
| 2505 | 55.67% |
| 4550 | 2.00% |
| 2733 | 42.33% |
| 3655 | 28.00% |
| 3911 | 29.33% |
| 2132 | 44.00% |
| 3929 | 34.00% |
| 2183 | 41.00% |
| 2450 | 42.67% |
| 5103 | 13.00% |
| 3470 | 42.67% |
| 2670 | 100.00% |
| 2654 | 38.00% |
| 5005 | 28.00% |
| 2384 | 7.67% |
| 0 | 100.00% |

RawReportData Data

| | |
|------|--------|
| 5520 | 15.33% |
| 4929 | 43.67% |
| 1565 | 45.33% |
| 2228 | 43.00% |
| 4664 | 68.33% |
| 2644 | 44.67% |
| 5127 | 13.33% |
| 5052 | 39.00% |
| 1122 | 88.33% |
| 3573 | 30.33% |
| 5946 | 2.00% |
| 2859 | 33.33% |
| 3276 | 38.33% |
| 4154 | 29.33% |
| 3452 | 30.67% |
| 5950 | 1.33% |

RawReportData Data

| | |
|------|---------|
| 2733 | 33.00% |
| 4845 | 42.00% |
| 5104 | 41.33% |
| 2967 | 53.00% |
| 5281 | 9.67% |
| 3048 | 47.00% |
| 2450 | 96.67% |
| 6511 | 18.50% |
| 4598 | 14.50% |
| 3563 | 21.50% |
| 3759 | 19.00% |
| 6448 | 11.50% |
| 3452 | 26.50% |
| 0 | 100.00% |
| 6953 | 13.50% |
| 6267 | 32.50% |

RawReportData Data

| | |
|------|--------|
| 2548 | 43.50% |
| 3281 | 29.50% |
| 6049 | 23.00% |
| 3757 | 17.50% |
| 6527 | 20.00% |
| 6487 | 13.00% |
| 2067 | 11.00% |
| 4656 | 23.50% |
| 7446 | 2.00% |
| 3947 | 22.50% |
| 4416 | 12.00% |
| 5064 | 18.00% |
| 4572 | 16.00% |
| 7435 | 3.00% |
| 3541 | 38.50% |
| 6260 | 17.00% |

RawReportData Data

| | |
|------|---------|
| 6527 | 15.50% |
| 4092 | 15.00% |
| 6716 | 13.00% |
| 4121 | 25.50% |
| 3358 | 18.50% |
| 7908 | 20.67% |
| 5590 | 61.67% |
| 4525 | 27.67% |
| 4721 | 67.67% |
| 7846 | 20.33% |
| 4637 | 23.00% |
| 0 | 100.00% |
| 8278 | 35.00% |
| 7427 | 68.00% |
| 3526 | 64.33% |
| 4389 | 38.33% |

RawReportData Data

| | |
|------|--------|
| 7412 | 27.33% |
| 4849 | 41.67% |
| 7927 | 20.00% |
| 7847 | 28.00% |
| 2949 | 43.67% |
| 5916 | 8.00% |
| 8929 | 3.33% |
| 5060 | 37.33% |
| 5479 | 47.33% |
| 5064 | 51.67% |
| 5729 | 28.67% |
| 8923 | 2.33% |
| 4181 | 92.00% |
| 6260 | 41.33% |
| 7920 | 21.33% |
| 5254 | 27.67% |

RawReportData Data

| | |
|------|---------|
| 8128 | 17.67% |
| 5286 | 27.00% |
| 4133 | 65.00% |
| 9243 | 33.00% |
| 5590 | 33.33% |
| 5305 | 84.00% |
| 5808 | 62.67% |
| 9099 | 49.33% |
| 5895 | 28.33% |
| 0 | 100.00% |
| 9676 | 20.33% |
| 8547 | 76.00% |
| 4443 | 96.67% |
| 5519 | 54.00% |
| 8522 | 78.00% |
| 4849 | 97.00% |

RawReportData Data

| | |
|-------|--------|
| 9270 | 31.33% |
| 9187 | 32.00% |
| 3869 | 56.00% |
| 7223 | 18.67% |
| 10421 | 1.67% |
| 6308 | 30.33% |
| 6641 | 47.67% |
| 5847 | 43.33% |
| 6947 | 36.33% |
| 10423 | 1.33% |
| 4974 | 81.33% |
| 6260 | 87.33% |
| 9212 | 41.67% |
| 6429 | 45.00% |
| 9490 | 27.67% |
| 6423 | 52.67% |

RawReportData Data

| | |
|-------|---------|
| 5035 | 59.67% |
| 10693 | 10.00% |
| 6502 | 17.67% |
| 5305 | 100.00% |
| 7265 | 8.67% |
| 10567 | 6.33% |
| 7117 | 55.67% |
| 0 | 100.00% |
| 11041 | 27.00% |
| 9994 | 10.67% |
| 5891 | 10.33% |
| 6661 | 71.67% |
| 9957 | 13.00% |
| 5801 | 9.67% |
| 10722 | 9.67% |
| 10580 | 21.33% |

RawReportData Data

| | |
|-------|--------|
| 5111 | 11.67% |
| 8688 | 7.00% |
| 11921 | 1.33% |
| 7725 | 16.67% |
| 8103 | 7.67% |
| 6889 | 11.67% |
| 8367 | 16.00% |
| 11923 | 1.00% |
| 6116 | 31.67% |
| 7198 | 12.33% |
| 10675 | 7.33% |
| 7892 | 7.33% |
| 10960 | 6.00% |
| 7878 | 9.00% |
| 6037 | 59.67% |
| 12113 | 16.00% |

RawReportData Data

| | |
|-------|---------|
| 6502 | 15.00% |
| 5305 | 100.00% |
| 8728 | 7.50% |
| 11897 | 34.00% |
| 8545 | 14.50% |
| 0 | 100.00% |
| 12426 | 23.00% |
| 11437 | 11.50% |
| 5891 | 92.50% |
| 8056 | 21.00% |
| 11302 | 31.00% |
| 5801 | 75.00% |
| 12147 | 15.00% |
| 11878 | 40.50% |
| 6236 | 55.00% |
| 10148 | 8.00% |

RawReportData Data

| | |
|-------|--------|
| 13406 | 3.00% |
| 9140 | 17.00% |
| 9443 | 32.00% |
| 6889 | 6.00% |
| 9660 | 41.50% |
| 13411 | 2.50% |
| 7256 | 52.00% |
| 8036 | 52.50% |
| 11843 | 66.50% |
| 9042 | 70.00% |
| 12408 | 10.50% |
| 9246 | 26.50% |
| 7302 | 27.00% |

| Answer Time (seconds) |
|-----------------------|
| 5 |
| 9 |
| 7,4 |
| 3,2 |
| 14,1 |
| 19,5 |
| 27,5 |
| 4,7 |
| 16,1 |
| 26,8 |
| 19 |
| 13,4 |
| 26,7 |
| 4,1 |
| 6,9 |

| |
|------|
| 27,8 |
| 7,9 |
| 0,8 |
| 20,7 |
| 11,8 |
| 4,2 |
| 12,9 |
| 1 |
| 5,6 |
| 22 |
| 19,6 |
| 19,1 |
| 2,8 |
| 19,8 |
| 8,7 |
| 14,8 |

| |
|------|
| 11,7 |
| 13,5 |
| 5,5 |
| 8,2 |
| 28,6 |
| 30 |
| 6,6 |
| 9,3 |
| 13,4 |
| 22,1 |
| 18,6 |
| 24,3 |
| 15,6 |
| 13,1 |
| 12,8 |
| 4,4 |

| |
|------|
| 1,1 |
| 27,1 |
| 21,6 |
| 10,1 |
| 6,1 |
| 0,8 |
| 15,7 |
| 20,3 |
| 5,6 |
| 25,8 |
| 12,1 |
| 22,9 |
| 17,5 |
| 18 |
| 8,6 |
| 16,9 |

| |
|------|
| 5,7 |
| 16,4 |
| 22 |
| 30 |
| 10,5 |
| 10,9 |
| 28,8 |
| 8,4 |
| 10,4 |
| 26,9 |
| 17 |
| 12,5 |
| 26,4 |
| 6,2 |
| 0,8 |
| 23,6 |

| |
|------|
| 5,5 |
| 3,5 |
| 7 |
| 0,6 |
| 5,6 |
| 6 |
| 7,3 |
| 5,1 |
| 15,1 |
| 6,1 |
| 7,1 |
| 12,1 |
| 10,3 |
| 30 |
| 12,2 |
| 12,6 |

| |
|------|
| 21,2 |
| 30 |
| 2,4 |
| 14,9 |
| 18,5 |
| 20,4 |
| 17,3 |
| 29 |
| 11,7 |
| 12,7 |
| 24,9 |
| 16,2 |
| 0,4 |
| 22,6 |
| 10,5 |
| 9 |

| |
|------|
| 16,7 |
| 0,6 |
| 12,7 |
| 8,4 |
| 8,8 |
| 13,2 |
| 10,2 |
| 12,3 |
| 12,8 |
| 3,9 |
| 12,8 |
| 30 |
| 11,4 |
| 8,4 |
| 2,3 |
| 30 |

| |
|------|
| 4,6 |
| 13,1 |
| 13,6 |
| 12,9 |
| 20,5 |
| 13,4 |
| 4 |
| 11,7 |
| 26,5 |
| 9,1 |
| 0,6 |
| 10 |
| 11,5 |
| 8,8 |
| 9,2 |
| 0,4 |

| |
|------|
| 9,9 |
| 12,6 |
| 12,4 |
| 15,9 |
| 2,9 |
| 14,1 |
| 29 |
| 3,7 |
| 2,9 |
| 4,3 |
| 3,8 |
| 2,3 |
| 5,3 |
| 20 |
| 2,7 |
| 6,5 |

| |
|-----|
| 8,7 |
| 5,9 |
| 4,6 |
| 3,5 |
| 4 |
| 2,6 |
| 2,2 |
| 4,7 |
| 0,4 |
| 4,5 |
| 2,4 |
| 3,6 |
| 3,2 |
| 0,6 |
| 7,7 |
| 3,4 |

| |
|------|
| 3,1 |
| 3 |
| 2,6 |
| 5,1 |
| 3,7 |
| 6,2 |
| 18,5 |
| 8,3 |
| 20,3 |
| 6,1 |
| 6,9 |
| 30 |
| 10,5 |
| 20,4 |
| 19,3 |
| 11,5 |

| |
|------|
| 8,2 |
| 12,5 |
| 6 |
| 8,4 |
| 13,1 |
| 2,4 |
| 1 |
| 11,2 |
| 14,2 |
| 15,5 |
| 8,6 |
| 0,7 |
| 27,6 |
| 12,4 |
| 6,4 |
| 8,3 |

| |
|------|
| 5,3 |
| 8,1 |
| 19,5 |
| 9,9 |
| 10 |
| 25,2 |
| 18,8 |
| 14,8 |
| 8,5 |
| 30 |
| 6,1 |
| 22,8 |
| 29 |
| 16,2 |
| 23,4 |
| 29,1 |

| |
|------|
| 9,4 |
| 9,6 |
| 16,8 |
| 5,6 |
| 0,5 |
| 9,1 |
| 14,3 |
| 13 |
| 10,9 |
| 0,4 |
| 24,4 |
| 26,2 |
| 12,5 |
| 13,5 |
| 8,3 |
| 15,8 |

| |
|------|
| 17,9 |
| 3 |
| 5,3 |
| 30 |
| 2,6 |
| 1,9 |
| 16,7 |
| 30 |
| 8,1 |
| 3,2 |
| 3,1 |
| 21,5 |
| 3,9 |
| 2,9 |
| 2,9 |
| 6,4 |

| |
|------|
| 3,5 |
| 2,1 |
| 0,4 |
| 5 |
| 2,3 |
| 3,5 |
| 4,8 |
| 0,3 |
| 9,5 |
| 3,7 |
| 2,2 |
| 2,2 |
| 1,8 |
| 2,7 |
| 17,9 |
| 3,2 |

| |
|------|
| 3 |
| 20 |
| 1,5 |
| 6,8 |
| 2,9 |
| 20 |
| 4,6 |
| 2,3 |
| 18,5 |
| 4,2 |
| 6,2 |
| 15 |
| 3 |
| 8,1 |
| 11 |
| 1,6 |

| |
|------|
| 0,6 |
| 3,4 |
| 6,4 |
| 1,2 |
| 8,3 |
| 0,5 |
| 10,4 |
| 10,5 |
| 13,3 |
| 14 |
| 2,1 |
| 5,3 |
| 5,4 |