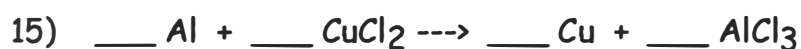
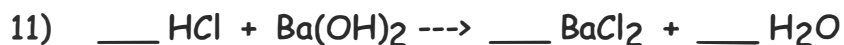
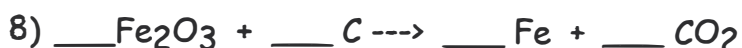
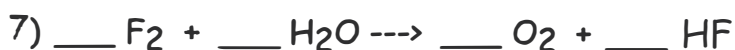
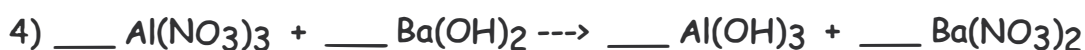
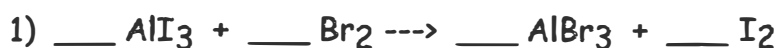


Part A: Balancing Chemical Equations



Part B: Identifying the Type of Reaction

For each reaction below, identify the type of reaction.

_____ 1) sodium + bromine $\rightarrow$ sodium bromide

_____ 2) calcium chloride $\rightarrow$ calcium + chlorine

_____ 3) ethane (C_2H_6) + oxygen $\rightarrow$ carbon dioxide + water

_____ 4) aluminum + sulfur $\rightarrow$ aluminum sulfide

_____ 5) barium chloride + copper (II) sulfate $\rightarrow$ barium sulfate + copper (II) chloride

_____ 6) copper (II) carbonate $\rightarrow$ copper (II) oxide + carbon dioxide

_____ 7) nickel + lead (IV) nitrate $\rightarrow$ lead + nickel (II) nitrate

_____ 8) propane (C_3H_8) + oxygen $\rightarrow$ carbon dioxide + water

_____ 9) diphosphorous pentoxide + water $\rightarrow$ phosphoric acid

_____ 10) mercury (II) oxide $\rightarrow$ mercury + oxygen

_____ 11) gold + magnesium bromide $\rightarrow$ gold (III) bromide + magnesium

_____ 12) hydrochloric acid + sodium hydroxide $\rightarrow$ sodium chloride + water

_____ 13) potassium oxide + water $\rightarrow$ potassium hydroxide

_____ 14) magnesium chlorate $\rightarrow$ magnesium chloride + oxygen

_____ 15) butane (C_4H_{10}) + oxygen $\rightarrow$ carbon dioxide + water

Part C: Predicting the Products

Predict the products for each reaction below. Then write the balanced chemical equation.

Hint: It will help to identify the type of reaction first.

1) iron (II) nitrate + potassium hydroxide ---->

2) zinc + copper (I) bromide ---->

3) hydrogen + oxygen ---->

4) water + sulfur trioxide ---->

5) sulfuric acid + potassium hydroxide ---->

6) barium carbonate ---->

7) phosphorous pentachloride ---->

8) silver nitrate + calcium bromide ---->