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Block_____

Unit 3 Notes: Atomic Structure

1. Which subatomic particles are located in the nucleus of the atom?
2. Which subatomic particle orbits outside the nucleus?
3. Fill in the table with information about the subatomic particles.

Particle	Symbol	Charge	Relative Mass	Location

4. Where is all the mass of an atom located?
5. In order for an atom to be neutral, the amount of which two subatomic particles must be equal?
6. Are all atoms of for a given element the same?
7. Define the term isotope and give examples.
8. What are the three parts of an isotopic symbol?
9. Are all subatomic particles the same size?
10. Which two subatomic particles are about the same size?
11. Which subatomic particle is 2000 times smaller?
12. Which elemental isotope was chosen in the Compromise of 1961?
13. What fraction represents 1 amu? Why?

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14. What does the MASS NUMBER represent?

15. What does the ATOMIC MASS represent?

16. There are two ways to calculate average atomic mass. Show both equations below:

17. What is the difference between ATOMIC MASS and MASS NUMBER?